

Highly Reliable Superhydrophobic Protection for Organic Field-Effect Transistors by Fluoroalkylsilane-Coated TiO₂ Nanoparticles

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ABSTRACT

One of the long-standing problems in the field of organic electronics is their instability in the open environment, especially poor water-resistance. For reliable operation of organic devices, introducing an effective protection layer using organo-compatible materials and processes is highly desirable. Here, we report a facile method to deposit an organo-compatible superhydrophobic protection layer on organic semiconductors under ambient conditions. The protection layer exhibiting excellent water-repellency and self-cleaning abilities was deposited onto organic semiconductors directly by using a dip-coating process in a highly fluorinated solution with fluoroalkylsilane-coated titanium-dioxide (TiO₂) nanoparticles. The proposed protection layer did not damage the underlying organic semiconductors and had good resistance against mechanical-, thermal-, light-stress and water-based threats. The protected organic field-effect transistors (OFETs) exhibited more reliable electrical properties even exposed to strong solvents due to its superhydrophobicity. This study provides a practical solution to enhance the reliability of environmentally sensitive organic semiconductor devices in the natural environment.

KEYWORDS: Superhydrophobic surface, organic semiconductor, nanoparticles, reliability, organic electronics

Organic semiconductors have attracted significant interest for electronic device applications, low-cost, and low-temperature processabilities.¹⁻⁷ Also, they can be regarded as one of the promising candidates to be employed in additive manufacturing such as printing processes.⁸⁻¹¹ Nevertheless, it is widely reported that they suffer from environmental instability; in particular, they can be drastically degraded by water-based hindrances.¹²⁻¹⁵ Although many efforts to address this issue by encapsulating with various materials such as parylene, silicone, barrier-foil, or a superhydrophobic glass have been reported,^{13,16-19} studies on organo-compatible materials and processing to introduce an effective protection layer which can be directly deposited onto organic components are highly desirable. In this regard, the introduction of a superhydrophobic protection layer onto organic semiconductors can be a promising approach to realize more reliable organic semiconductor applications, because superhydrophobic protection layers provide the attractive water-repellency, which could eliminate water-based hindrances from surfaces of organic semiconductors. The materials and processes for the superhydrophobic protection layer formation need to be organo-compatible, so that they do not damage physically or chemically when placed on the organic layers, and thus their electrical characteristics would not be degraded. In addition, fast, facile, and low-cost approaches to implement the protection layer are desirable to bring the advantages of organic components.

In this study, we report an organo-compatible superhydrophobic

protection layer which can be directly deposited onto various organic semiconductors for highly reliable OFETs. The protection layer could be introduced by simply dipping OFETs into a highly fluorinated solvent with fluoroalkylsilane-coated TiO₂ nanoparticles. By optimizing the solvent system with a consideration of surface energy, the uniform protection layer could be deposited onto universal organic semiconductors without any physical damages. The surface roughness of TiO₂ nanoparticle-based protection layer allowed the excellent water repellency and self-cleaning abilities regardless of underlying organic semiconductors. The presented superhydrophobic protection layer provided much improved stability of OFETs against water-based threats and ambient air while exhibiting good resistance against thermal- and mechanical- stresses. Therefore, this method can foster to realize more reliable organic electronics working in the outdoor environment.

RESULTS AND DISCUSSION

A solution process using 1*H*,1*H*,2*H*,2*H*-perfluorooctyltriethoxysilane (PFOTES) attached to TiO₂ nanoparticles (average diameter < 100 nm) dispersed in 3-ethoxy-1,1,1,2,3,4,4,5,5,6,6,6-dodecafluoro-2-trifluoromethylhexane (HFE-7500) was conducted to form the protection layer (Figure 1a). Highly fluorinated HFE-7500 is immiscible to most organic materials; thus, the solution-dipping process does not damage organic layer during the protection layer formation.²⁰ Moreover, PFOTES has been widely

used to introduce superhydrophobic properties on the surface of TiO₂ nanoparticles by reacting with hydroxyl groups (Figure 1b).²¹⁻²³ Therefore, PFOTES-coated TiO₂ nanoparticle layers can deliver superhydrophobic properties to protect the underlying organic layer from water-based solutions without additional surface treatments, structures, or dedicated procedures (Figure 1c,d). Note that the PFOTES-coated TiO₂ nanoparticles highly dispersed in HFE-7500 (14.0 wt%) allow an excellent film formation universally on various organic semiconductors (pentacene, 6,13-bis(triisopropylsilylethynyl)pentacene (TIPS-pentacene), and poly(3-hexylthiophene) (P3HT) for thermally evaporated organic small-molecule, solution-deposited organic small-molecule, and solution-deposited polymer semiconductors, respectively) by simply dipping them into the prepared solution for a short time (~2 s) in ambient (see Figure S1 and Table S1 in the Supporting Information for a detailed description). It should be noted that this simple dipping process was enough to produce the uniformly deposited TiO₂ nanoparticle layers because the solution had the excellent dispersion ability and wetting property onto all of the semiconductors used in this study as well as on the Au contacts and SiO₂ gate dielectric (see Figures S2-S4 in the Supporting Information). As a result, the protection layers having similar morphologies were produced onto the all different organic semiconductors (Figure 1e-g). In addition, the surface coverages after coating with different TiO₂ nanoparticle densities were also studied (see Figure S5 in the Supporting Information). The surface roughness of the PFOTES-coated TiO₂

nanoparticle layers was measured using 3-dimensional (3D) laser profiler and atomic force microscope (AFM) (see Figures S6 and S7 in the Supporting Information). The surface roughness of TiO₂ nanoparticles and the low surface energy induced by the trifluoromethyl (-CF₃) groups²⁴ of PFOTES make the superhydrophobic surface,²⁵ which delivered a high contact angle > 150° regardless of the underlying organic semiconducting layers (Figure 2a), whereas the bare pentacene layer exhibited a water contact angle of 81° (Figure 2b). Also, the low contact angle hysteresis (< 5°) was observed for liquids having surface tension values over 44 mN m⁻¹ (see Figure S8 in the Supporting Information). It is well-known that high contact angles and low contact angle hysteresis can be observed in micro/nanoscale hierarchical superhydrophobic surfaces because of air gaps in roughened surfaces, that play a critical role for water repellency of superhydrophobic surfaces (see Figure S9 in the Supporting Information).²⁵⁻²⁸ Therefore, our surface morphology produced by the simple dipping process with PFOTES-coated TiO₂ nanoparticles created enough air gaps to allow extreme water repellency showing the ability of water-bouncing comparable to other reported superhydrophobic surfaces (Figure 2c).^{19,22,23,25-31} For further investigation, various liquid droplets having different surface tensions were also dropped with various speeds ranging from 0.3 m s⁻¹ to 1.7 m s⁻¹ onto the surface of protection layer. Even liquid droplets having surface tension value > 44 mN m⁻¹ were easily bounced off from the surface of our superhydrophobic protection layer as a whole or fragmented droplets (see

Table S2 and Figure S10 in the Supporting Information). In addition, an extremely small roll-off angle ($< 1^\circ$) was observed for the deposited PFOTES-coated TiO_2 nanoparticle layers whereas relatively large roll-off angle of 34° was observed for the bare pentacene (see Figure S11 in the Supporting Information). The small roll-off angle of the PFOTES-coated TiO_2 nanoparticle layers allows the self-cleaning ability upon exposure to water-based solvents. In other words, undesirable substances, even heavy soils, were successfully swept away by continuously dropping water droplets because the sliding water droplets picked up the soils from the surface of the self-cleaning layer (see Figure S12 in the Supporting Information). The presented water-bouncing and self-cleaning abilities from the roughened surface of the PFOTES-coated nanoparticles could offer the superhydrophobic protection for OFETs to realize more practical applications.

To examine the proposed superhydrophobic protection layer, X-ray photoelectron spectroscopy (XPS) was conducted on the TiO_2 nanoparticle films with and without the PFOTES layer coated by the aforementioned facile dipping process. The XPS spectrum showed intense photoelectron signals on the surface at binding energies of 457 eV and 463 eV related to the Ti $2p_{1/2}$ and $2p_{3/2}$ levels of the titanium atoms, respectively.^{22,32} The F 1s peak was clearly observed at the binding energy of 687 eV on the PFOTES-coated TiO_2 nanoparticle film (Figure 2d), whereas the TiO_2 nanoparticle film without PFOTES showed the absence of fluorine atoms (see Figure S13 in the Supporting Information).²² Because no clear peak related to the F 1s orbital

was observed in the TiO₂ nanoparticle film without PFOTES, the fluorine contained in HFE-7500 did not remain on the surface of the TiO₂ nanoparticles after the drying process. Note that the TiO₂ nanoparticle layers without PFOTES did not exhibit superhydrophobicity at all, which indicates that PFOTES plays a critical role in achieving low-surface energy (see Figure S14 in the Supporting Information).

To confirm that the methods and materials used in this work are organo-compatible, cross-sectional images of the organic semiconducting layers with and without the PFOTES-coated TiO₂ nanoparticle layers were obtained. Scanning electron microscopy (SEM) images revealed that the PFOTES-coated TiO₂ nanoparticle layers were well deposited on the pentacene semiconducting layer, and they did not produce any physical damages to the pentacene layer (Figure 2e,f). Moreover, the cross-sectional transmission electron microscope (TEM) and energy-dispersive X-ray spectroscopy (EDS) results also indicated that no undesirable penetration was observed after the superhydrophobic protection layer deposition because the titanium and oxygen elements were evidently detected only in the upper PFOTES-coated TiO₂ nanoparticle layers but were undetected in the underlying pentacene layer (Figure 2g). We also did not observe physical damages and penetration issues on solution-processed TIPS-pentacene and P3HT layers during the protection layer formation (see Figure S15 in the Supporting Information). Therefore, the prepared superhydrophobic protection layer could be universally applicable to various organic electronics without chemically and

physically penetrating into underlying sensitive organic semiconductors. In addition to *p*-type organic semiconductors, our superhydrophobic protection layer could be applied to *n*-type organic semiconductors without physical and chemical damages (see Figure S16 in the Supporting Information).

To use organic devices in practical applications, an excellent tolerance to realistic circumstances is necessary to achieve reliable operation maintaining their electrical characteristics when exposed to, for example, water, acid rain (pH < 5.6),³³ various beverages (pH 3.0-6.5),³⁴ and cleaning soaps (pH 9.0-12.0)^{35,36} that can be met in daily life. Therefore, an effective protection layer should have a good robustness against liquid-based substances with a wide range of pH values to prevent operation failure. To confirm the robustness of the protection film, contact angles were measured after exposure to various corrosive liquids, such as hydrochloric acid (HCl, pH 1.0), sodium hydroxide (NaOH, pH 13.0), acetic acid (CH₃COOH, pH 2.7), ammonium hydroxide (NH₃OH, pH 11.7), and deionized (DI) water for 10 s. Even though the prepared protection film was exposed to strong acid and alkaline solvents, its superhydrophobicity was well-maintained exhibiting a contact angle of over 150° (Figure 3a). Detailed information on the robustness of the prepared protection layers to corrosive liquids is presented in Figure S17. Although superhydrophobicity of the PFOTES-coated TiO₂ nanoparticle layers slightly degraded showing a contact angle of ~130° due to increased surface energy after exposure to air for one year, the PFOTES-coated TiO₂ nanoparticle layers still exhibited sufficient hydrophobicity to prevent water-based threats,

which indicates good durability in ambient (Figure 3b). With these attractive benefits, the proposed protection layer was used for environmentally sensitive OFETs (Figure 4a) and the protection layer had good resistance upon mechanical-, thermal- and light-stress.

We performed an adhesion test with a piece of commercial Scotch tape over 15 times for the superhydrophobic protection layer on different organic semiconductors. The results showed that the PFOTES-coated TiO₂ nanoparticle layers exhibited good adhesion properties without delamination issues on various organic semiconductors, maintaining their superhydrophobicity with a contact angle over 150° in these durability tests (Figure 4b). Furthermore, under shear pressure by friction between a commercial rubber roller and the prepared PFOTES-coated TiO₂ nanoparticle layers, their superhydrophobicity was well-maintained during the 50-rollings with shearing pressure of 1 kPa and speed of 1 cm s⁻¹ (see Figure S18 in the Supporting Information). Importantly, the rough surface which is a critical factor for introducing superhydrophobicity was still well-preserved without critical structure-collapsing under vertical pressure up to 50 N cm⁻², exhibiting an average water contact angle over 150° (see Figure S19 in the Supporting Information). Also, superhydrophobicity of the PFOTES-coated TiO₂ nanoparticle layers under tensile strain was well maintained after 10,000 bending cycles and even at stretching conditions (see Figures S20 and S21 in the Supporting Information). Furthermore, the superhydrophobicity did not degrade against a tap water with dynamic

pressure of approximately 100 kPa for 100 s and its extreme water repellency facilitated tap water to bounce off from the surface (see Figure S22 in the Supporting Information).

The thermal stability of the PFOTES-coated TiO₂ nanoparticle layers was also evaluated by measuring a water contact angle after thermal treatment processes. The water contact angles of the PFOTES-coated TiO₂ nanoparticle layer were well-maintained when placed onto a 80 °C hot-plate for a day (see Figure S23 in the Supporting Information), but its superhydrophobicity was drastically degraded as the annealing temperature was increased up to 250 °C, because it is close to the desorption temperature of PFOTES (see Figure S23 in the Supporting Information).³⁷ However, the stable superhydrophobic property was exhibited at 200 °C, even the PFOTES-coated TiO₂ nanoparticle layers were kept on a hot-plate for 3 h regardless of underlying organic semiconductors (see Figure S23 in the Supporting Information).

Two ultraviolet (UV) lamps with a peak wavelength of 300 nm and 360 nm were used to investigate the changes of superhydrophobicity after exposure to UV light. The following results in Figure 4c and d indicated that the contact angle decreased after exposure to UV light with power of 3500 erg s⁻¹ cm⁻² corresponding to 0.35 mW cm⁻² for 10³ s. The degradation of the contact angle was delayed up to 10⁵ s as UV lamp power was reduced (~100 erg s⁻¹ cm⁻² corresponding to 0.01 mW cm⁻²). In addition, the degradation of superhydrophobicity after exposing to visible lights was also investigated (see Figure S24 in the Supporting Information). By exploiting these

behaviors, switchable surface energy of the TiO₂ surfaces is available exhibiting reversible superhydrophobicity (>150°) and superhydrophilicity (~0°) (see Figure S24 in the Supporting Information).

We systematically performed to figure out that our superhydrophobic protection layers produced by the facile, organo-compatible and universal process could be a good alternative strategy to improve electrical reliability of OFETs for water-based threats. We characterized the electrical characteristics of various OFETs without and with PFOTES-coated TiO₂ nanoparticle protection layers before and after exposure to DI water (pH 6.7), HCl (pH 1.0), and NaOH (pH 13.0). Aforementioned, three representative organic semiconductors (evaporated pentacene, solution processed TIPS-pentacene, and P3HT) were chosen to demonstrate that the suggested superhydrophobic protection layer can be applied to a wide range of organic semiconductors, maintaining their electrical characteristics. The transfer (drain-source current *versus* gate voltage, I_{DS} - V_{GS}) characteristics were measured by sweeping V_{GS} from 40 to -60 V at a fixed drain-source voltage (V_{DS}) of -60 V (Figure 5a for pentacene OFETs and Figure S25 for TIPS-pentacene and P3HT OFETs in the Supporting Information), and the output (drain-source current *versus* drain-source voltage, I_{DS} - V_{DS}) was measured by sweeping V_{DS} from 0 to -60 V at different V_{GS} from 0 to -60 V with an increment of -10 V (Figure 5b-d for pentacene OFETs and Figures S26 and S27 for TIPS-pentacene and P3HT OFETs, respectively in the Supporting Information). To induce ordered crystals of organic semiconductors, surface

treatment using a fluorinated solution was conducted on the SiO₂ dielectric layer, which provides a sufficient hydrophobic surface with a water contact angle of 108° and thus the improved electrical performances were achieved (see Figure S28 in the Supporting Information).³⁸ For the OFETs without the superhydrophobic protection layer, the unprotected devices exhibited a high I_{DS} over 10⁻⁵ A over the entire range of gate biases when exposed to water (Figure 5a,c and Figure S29 in the Supporting Information). Because extra current paths were created between the source-drain electrodes due to ionic conduction through the liquid on the bare active channel area, the current levels increased regardless of gate voltage. However, all of the protected OFETs showed no significant degradation of the electrical characteristics including a good current ratio (I_{on}/I_{off}) and field effect mobility (μ_{FET}) even when they were exposed to various corrosive solutions (Figure 5a,d, Figure S29 and Table S3 in the Supporting Information). Even, upon exposure to a tap water for 1000 s, we could not observe significant changes of the electrical characteristics (see Table S4 and Figure S30 in the Supporting Information). Specifically, the carrier injection properties in the low V_{DS} regime showed a good linearity without current degradation and were well maintained during the solvent-robustness tests (see Figure S31 in the Supporting Information). Furthermore, we performed the stability test measuring the electrical characteristics of the protected OFETs after storing them in a chamber with high humidity (@ 75% relative humidity (RH), 20 °C). The threshold voltages (V_{TH}) of the unprotected OFETs evidently shifted by

~3.4 V on average toward the positive gate voltage direction as soon as they were exposed to high humidity, unlike the behavior of the protected devices showing no notable degradations (see Figures S32 and S33 in the Supporting Information). After 5 days, the V_{TH} of the unprotected devices shifted by ~10 V while the V_{TH} of the protected devices shifted only less than ~5 V, shown in Figure 5e,f. It is well-known that this V_{TH} shift is attributed to water-molecules adsorption as previously reported.^{39,40} More detailed information on the changes in electrical characteristics of OFETs with exposure to humid air are summarized in Table S5, the Supporting Information. Among metal oxide nanoparticle candidates, TiO_2 nanoparticle layers can also be a barrier against the adsorption with water-molecule because TiO_2 layers have shown a lower water vapor transmission rate (see Figure S34 and detailed information in the Supporting Information Section 3-7).⁴¹⁻⁴³ In addition, PFOTES attached to the TiO_2 nanoparticles allows hydrophobicity to hinder the adsorption of water molecules.^{44,45} However, there are grain boundaries between the PFOTES-coated TiO_2 nanoparticles which could act as paths for the interaction of water molecules with organic semiconductors. So, although the proposed protective layer could not prevent the permeability of moisture permanently, it could improve the environment-reliability against moisture especially in the early stage since its estimated water vapor transmission rate value was $\sim 1.61 \text{ g m}^{-2} \text{ day}^{-1}$, which were comparable to plastic packaging materials (see detailed information in the Supporting Information Section 3-8).

CONCLUSIONS

In this study, the functionalized TiO₂ nanoparticle-based superhydrophobic protection layer was utilized to realize reliable OFETs. Because of the roughened surface and low surface energy of the protection layer, excellent water repellency and self-cleaning abilities were achieved, which could preserve the electrical characteristics of environmentally sensitive organic semiconductors. The suggested superhydrophobic protection layers were applied onto various organic semiconductors directly by a facile dipping process, exhibiting good resistances against mechanical-, thermal-, and light-stress and chemical-threats. In addition, they allowed more reliable electrical characteristics of OFETs in ambient, even exposed to strong solvents due to its superhydrophobicity. This approach can be a good alternative solution to protect low-cost and flexible organic electronics working in the opened air.

METHODS

Fabrication of organic field effect transistors. *p*-type heavily doped silicon substrates with 270 nm-thick silicon dioxide (SiO₂) layers were sequentially cleaned with acetone, isopropanol, and DI water for 10 min each. On the cleaned substrate, Au/Ti (50 nm/5 nm) source and drain electrodes were deposited using an electron-beam evaporator with a deposition rate of 0.5 Å s⁻¹ at a pressure of ~10⁻⁷ torr. The surface treatment using a fluorinated solution was conducted on the SiO₂ gate dielectric for the better organic semiconducting layer deposition. After UV-ozone treatment for 10 s was done to yield better wetting properties for the SiO₂ layer. Then, a PFOTES (0.5 g) solution dispersed in HFE-7200 (7 ml) was spin-coated with a spin speed of 500 rpm for 30 s and unreacted residues of PFOTES were removed by rinsing with toluene. Samples were baked at 100 °C for 10 min on a hot plate. Then, a 120 nm-thick pentacene active layer was thermally evaporated onto the sample surfaces that had patterned Au/Ti electrodes with a deposition rate of 0.2 Å s⁻¹ at a pressure of ~10⁻⁵ torr. The completed pentacene OFETs had a channel length and width of 60 μm and 300 μm, respectively. For TIPS-pentacene OFETs, 0.5 wt% TIPS-pentacene solution in toluene was drop-casted on the patterned Au/Ti source-drain electrodes with a channel length and width of 200 μm and 300 μm, respectively, and then the active layer was dried in air for 1 h. For P3HT OFETs, 1 wt% P3HT solution in 1,2-Dichlorobenzene were spin-coated at 1500 rpm for 30 s onto the Au/Ti electrodes with a channel length and width of 100 μm and 300 μm,

respectively, and then the sample was placed on a hot-plate at 60 °C for 6 h in N₂.

Process of superhydrophobic protection layer formation. First, 8.0 g of TiO₂ nanoparticles (mixture of rutile and anatase phases, nanoparticle diameter < 100 nm) and 0.5 g of PFOTES were placed into 29.8 ml of HFE-7500. The fabricated OFETs with three organic semiconductors were dipped into the HFE-7500 solution containing PFOTES and TiO₂ nanoparticles at room temperature for a short time (~2 s). The PFOTES-coated TiO₂ nanoparticles in the solution covered the surface of the OFETs. Finally, the OFETs with the superhydrophobic protection layer were dried with N₂ gas to remove any remaining solvent or residues.

Characterization. Cross-sectional images of the devices were obtained using scanning electron microscopy (SEM) (MERLIN, ZEISS) and analytical scanning transmission electron microscopy (STEM) (JEM-2100F, JEOL). To determine the surface moieties on the PFOTES-coated TiO₂ nanoparticle layers, XPS (Axis-HSI, Kratos Inc.) was conducted with an Al monochromator anode at a power of 18 mA and 12 kV. Roughness parameters were evaluated by using 3D laser profiler (VK-250K, KENENCE) and AFM (NX10, Park Systems Corp.). The contact angle, roll-off angle and contact angle hysteresis of the OFETs with and without the superhydrophobic protection layer were measured by standard procedures (SmartDrop Lab HS, Femtofab) in ambient conditions.

The electrical characteristics of the fabricated OFETs were measured using a semiconductor parameter analyzer (Model 4200 SCS, Keithley) and a probe station (Model ST-500, JANIS).

Figure Captions

Figure 1. (a) Molecular structures of pentacene, TIPS-pentacene, P3HT, HFE-7500, and PFOTES. (b) Schematic for preparing the solution of PFOTES-coated TiO₂ nanoparticles. (c) Schematic of a water droplet contacting the PFOTES-coated TiO₂ nanoparticle layers on organic semiconducting layers. (d) Schematic for the interface between water and the fluorinated end-groups of PFOTES. (e-g) SEM images for the PFOTES-coated TiO₂ nanoparticles on (e) pentacene (f) TIPS-pentacene (g) P3HT layers

Figure 2. (a) Optical image of the water contact angle on a protected pentacene layer with the PFOTES-coated TiO₂ nanoparticles. (b) Optical image of the water contact angle on a bare pentacene layer. (c) Time-resolved images for water droplet bouncing on the PFOTES-coated TiO₂ nanoparticle layers. (d) XPS spectrum of PFOTES-coated TiO₂ nanoparticle layers. The inset image shows the rutile phase of TiO₂ and the molecular structure of PFOTES with the fluorine atoms highlighted in red. (e,f) SEM images of the pentacene layers (e) without and (f) with the PFOTES-coated TiO₂ nanoparticle layers. (g) TEM image and EDS data of the PFOTES-coated TiO₂ nanoparticle layers on a pentacene layer.

Figure 3. (a) Water contact angles of the superhydrophobic layer after exposure to HCl (pH 1.0), CH₃COOH (pH 2.5), DI water (pH 6.7), NH₄OH (pH

11.7), and NaOH (pH 13.0) droplets for 1, 5, and 10 s. (b) Water contact angles of the PFOTES-coated TiO₂ nanoparticle layers after exposure to air for 1 year.

Figure 4. (a) Fabrication process of OFETs with PFOTES-coated TiO₂ nanoparticle layers as a superhydrophobic protection layer. (b) Water contact angles of the superhydrophobic layer on three organic semiconductors in the adhesion test using a commercial scotch tape. (c, d) Degradation of water contact angles under UV light exposure with a peak-wavelength of (c) 300 and (d) 360 nm for different light-power.

Figure 5. (a) Transfer curves on the semilogarithmic scale for an unprotected pentacene OFET (*i.e.*, without PFOTES-coated TiO₂ nanoparticle layers) after exposure to DI water and for a protected pentacene OFET (*i.e.*, coated with PFOTES-coated TiO₂ nanoparticle layers) before and after exposure to aqueous solutions of different pH ranging from 1 to 13. (b) Output curves for unprotected pentacene OFETs before exposure to DI water. (c) Output curves for unprotected pentacene OFETs after exposure to DI water. (d) Output curves of the protected pentacene OFET with the PFOTES-coated TiO₂ nanoparticle layers before and after exposure to strong acid (HCl) and base (NaOH). (e, f) Transfer curves on the semilogarithmic scale for the (e) unprotected and (f) protected TIPS-pentacene OFETs after exposure to

water vapor (75% relative humidity, 20 °C).

ASSOCIATED CONTENT

Supporting Information

The Supporting Information is available free of charge on the ACS Publications website at DOI: XXX. Details on fabrication of OFETs with superhydrophobic protection layer, superhydrophobic properties of PFOTES-coated TiO₂ nanoparticle layers, XPS spectra of PFOTES-uncoated TiO₂ nanoparticle layers, TEM images and EDS profiles of TIPS-pentacene and P3HT semiconducting layer with the PFOTES-coated TiO₂ nanoparticle layers, durability of organo-compatible superhydrophobic protection layer, electrical characteristics of protected and unprotected OFETs, water vapor transmission rate of PFOTES-coated TiO₂ nanoparticle layers

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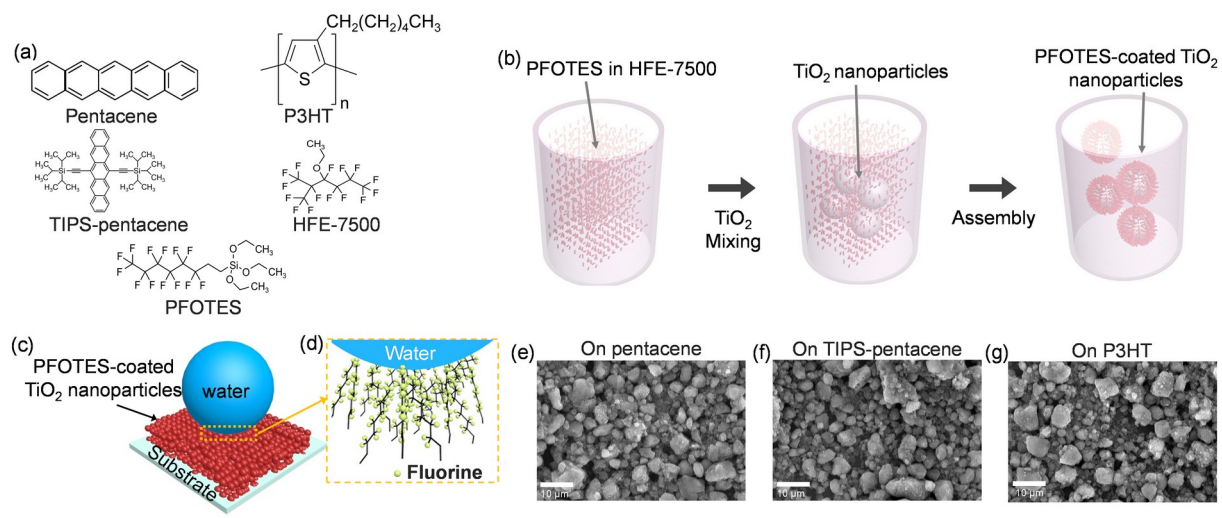


Figure 1. Yoo *et al.*

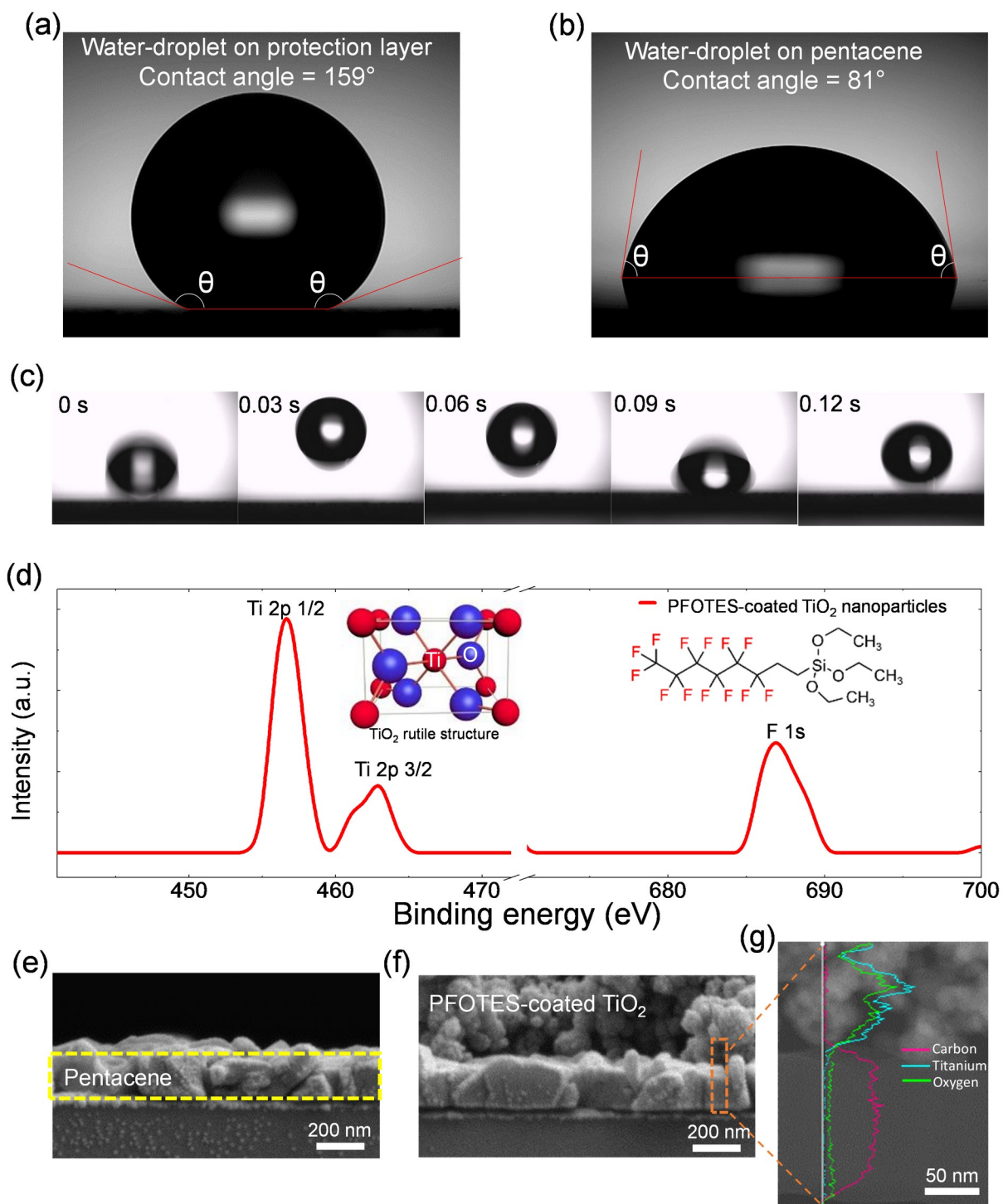


Figure 2. Yoo *et al.*

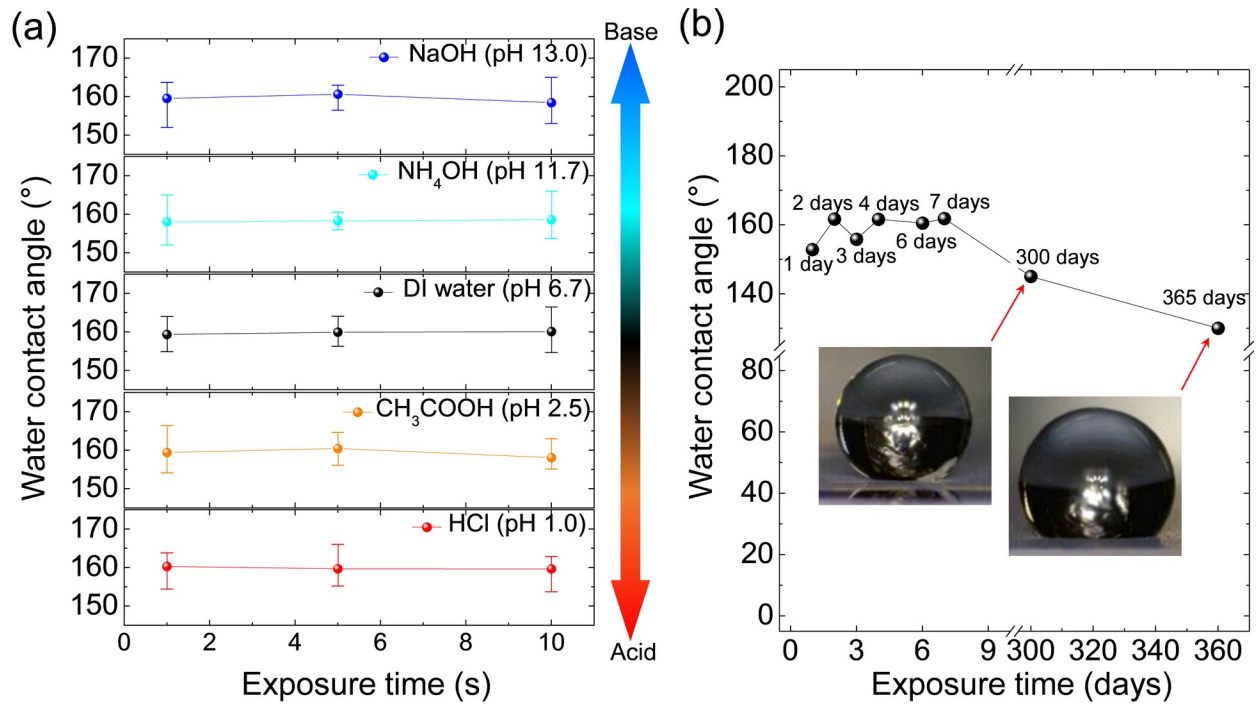


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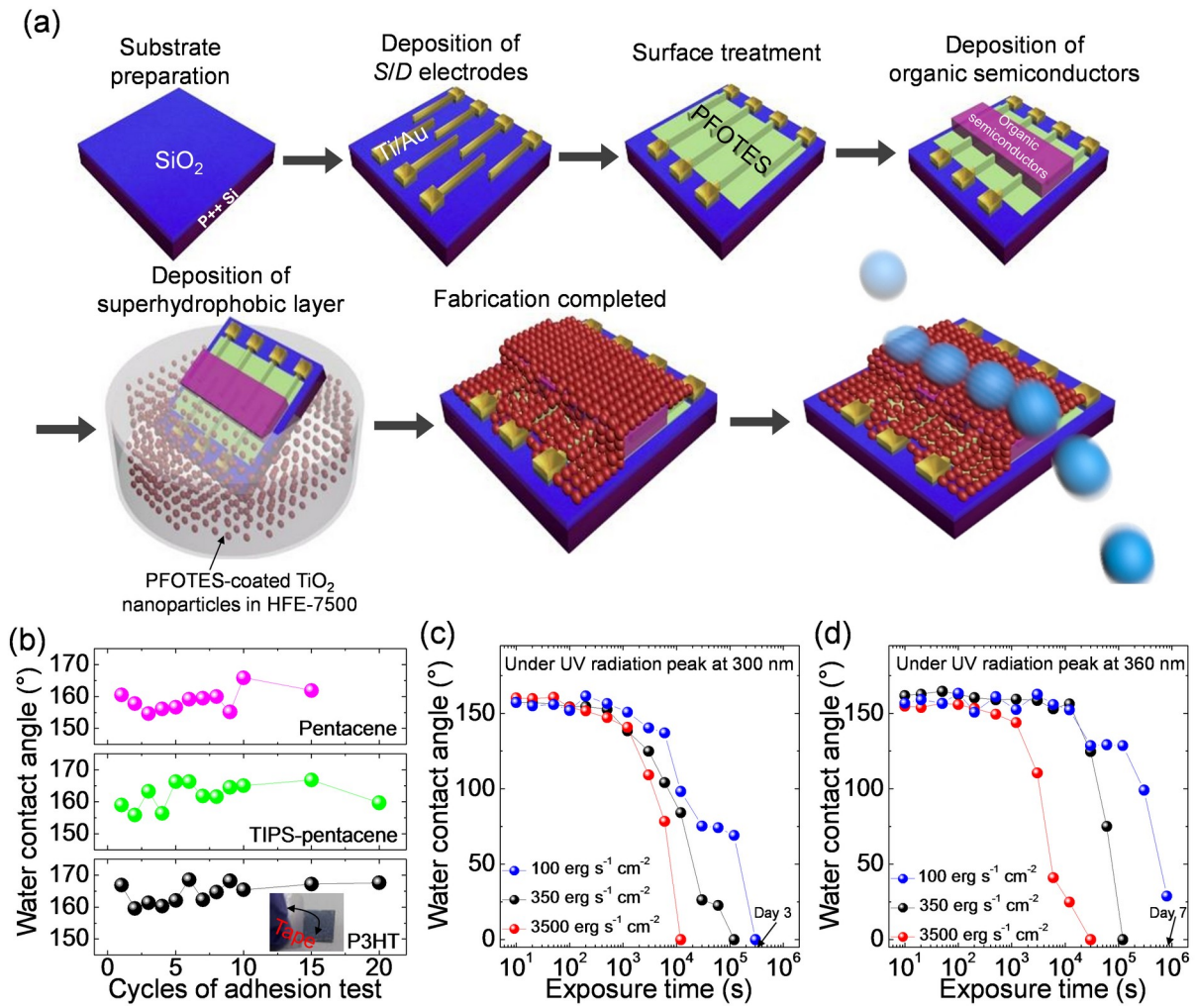


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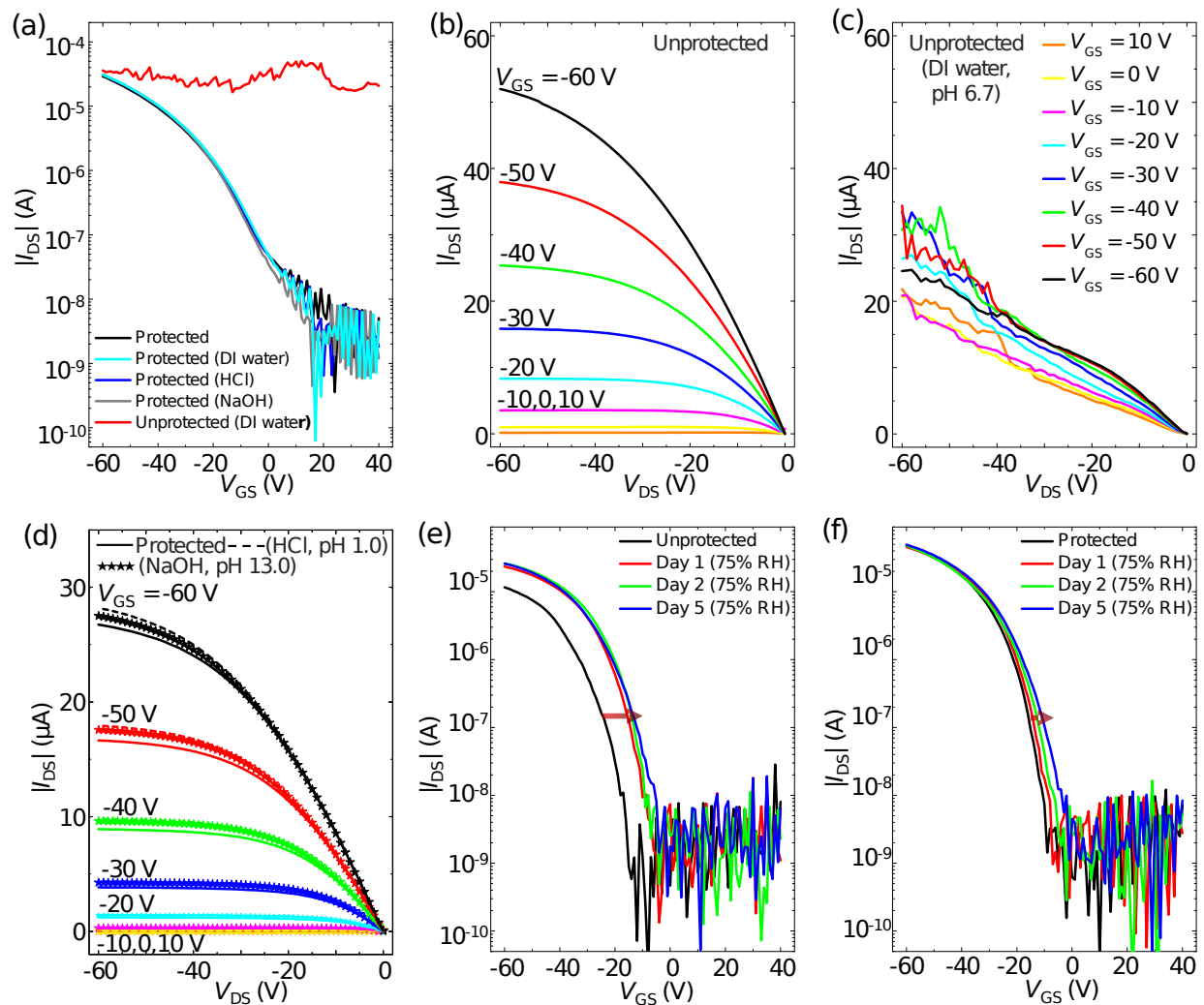


Figure 5. Yoo *et al.*

Table of Contents Graphic

Title: Highly Reliable Superhydrophobic Protection for Organic Field-Effect Transistors by Fluoroalkylsilane-Coated TiO₂ Nanoparticles

