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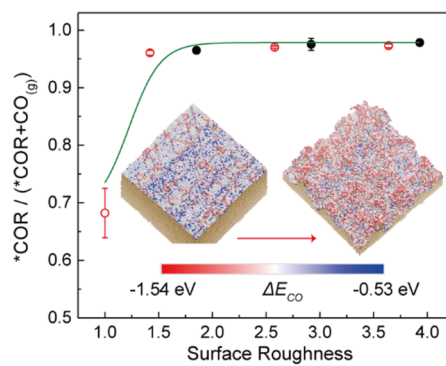
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Effects of Surface Roughness on the Electrochemical Reduction of CO₂ over Cu

Kun Jiang,[▽] Yufeng Huang,[▽] Guosong Zeng, Francesca M. Toma, William A. Goddard, III, and Alexis T. Bell*

ABSTRACT: We have investigated the role of surface roughening on the CO₂ reduction reaction (CO₂RR) over Cu. The activity and product selectivity of Cu surfaces roughened by plasma pretreatment in Ar, O₂, or N₂ were compared with that of electrochemically polished Cu samples. Differences in total and product current densities, the ratio of current densities for HER (the hydrogen evolution reaction) to CO₂RR, and the ratio of current densities for C₂₊ to C₁ products depend on the electrochemically active surface and are nearly independent of plasma composition. Theoretical analysis of an electropolished and roughened Cu surface reveals a higher fraction of undercoordinated Cu sites on the roughened surface, sites that bind CO preferentially. Roughened surfaces also contain square sites similar to those on a Cu(100) surface but with neighboring step sites, which adsorb OC-COH, a precursor to C₂₊ products. These findings explain the increases in the formation of oxygenates and hydrocarbons relative to CO and the ratio of oxygenates to hydrocarbons observed with increasing surface roughness.



Electrochemical CO₂ reduction reaction (CO₂RR) offers a promising route for the production of chemicals and fuels using renewable electricity generated from wind and solar energy.^{1–4} Among metallic electrocatalysts evaluated for this purpose, copper is the only metal that produces hydrocarbons and oxygenated products with high Faradaic efficiency (FE).^{5–10} Previous studies have shown that the activity and selectivity of Cu are strongly dependent on the surface morphology of the metal, as well as its local reaction environment (electrolyte composition and pH).^{11–16} For example, Cu(100) and Cu(211) surfaces are more active than Cu(111) surfaces and more effective in promoting C–C bond formation on both single-crystal Cu electrodes^{17,18} and Cu nanoparticles.^{19,20} Recent theoretical calculations have shown that square sites of Cu(100) bind *OCCO and *OCCHO more strongly than do sites on Cu(111) and that the step sites of Cu(211) facilitate the kinetics of CO dimerization relative to those on Cu(111), resulting in a higher C₂₊/C₁ selectivity on square and stepped facets.^{21,22} The role of other low-coordination Cu sites beyond those present on low-index facets is the subject of continuing discussion. An investigation of the CO₂RR over size-controlled Cu nanoparticles has also shown that the population of low-coordination Cu surface sites increases with decreasing Cu particle size (from 15 to 2 nm) and that low coordination sites result in higher H₂ and CO selectivities and lower hydrocarbon (CH₄ and C₂H₄) selectivities.²³ By contrast, a comparative investigation of

CO₂RR on (100)-, (111)-, and (751)-preferentially orientated Cu thin-film electrodes revealed that both Cu(100) and (751) surfaces with Cu coordination numbers from 6 to 8 promote C–C bond formation relative to more highly coordinated sites on Cu(111) surfaces.²⁴ More recently, a study of the electrochemical reduction of CO has suggested that highly porous Cu electrode could favor C₂₊ oxygenates selectivity with a large suppression of competitive HER; however, the underlying causes for these observations remains unclear.²⁵

Oxide-derived Cu with enriched surface grain boundaries have also been reported to be efficient for reducing CO₂ to multicarbon products.^{26–29} It has been hypothesized that this enhanced C₂₊ product selectivity on oxide-derived Cu could arise from surface Cu⁺ retained under reaction conditions and stabilized by subsurface oxygen^{28,30} or adsorbed halide anions,^{31,32} e.g., I[–], based on *quasi in situ* spectroscopic experiments. However, this hypothesis has been challenged by recent DFT calculations,^{33,34} ¹⁸O labeling,³⁵ electrochemical Raman spectroscopy,^{36,37} and *in situ* synchrotron XAS and XRD experiments,^{38,39} suggesting the near-surface oxygen and/or

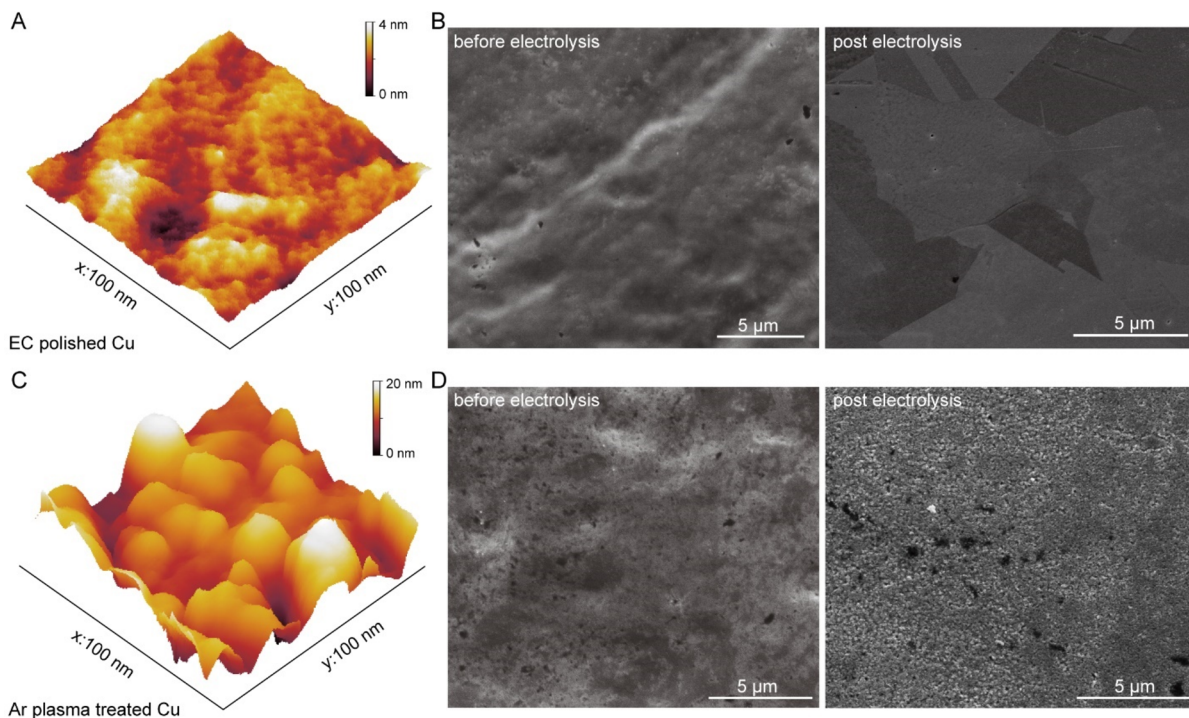


Figure 1. Cu electrodes characterization. AFM and SEM images of electrochemically polished Cu foil acquired before (A and B) and after a 10 min plasma pretreatment in Ar (C and D). Images A and C show the reconstructed 3D topography of the surface obtained from AFM scans; images B and D show typical SEM images of Cu foils taken before and after 1 h CO₂RR electrolysis at -1.0 V in CO₂-saturated 0.1 M CsHCO₃.

Table 1. Determined Surface Roughness of Post Electrolysis Cu Foils with Different Plasma Pretreatments

electrode	electropolished	N ₂ plasma		Ar plasma		O ₂ plasma		
		10 min	5 min	10 min	20 min	5 min	10 min	20 min
roughness	1.00	1.57	1.46	2.58	3.60	1.91	2.92	3.93

Cu(I) species are not sufficiently long-lived to be present under reaction conditions at highly reducing cathode potentials.⁴⁰ A more plausible interpretation for the enhanced C₂₊ selectivity of oxide-derived Cu is offered by a recent theoretical analysis of sites present on a rough Cu nanoparticle. This work concludes that surface twin boundaries in the oxide-derived Cu associated with concave defects with respect to Cu(100) planes serve as active sites for C–C bond formation by stabilizing OC–COH species, one of the precursors to C₂₊ products.⁴¹

In this study, we investigate the effects of plasma pretreatment of polished Cu foils in different gas atmospheres on their CO₂RR activity and selectivity. Because O₂-plasma treatment will cause both chemical and physical modifications to Cu surfaces, we pretreated Cu by Ar⁺ ion bombardment in an Ar plasma in order to isolate the effects of surface roughening. These studies show that the changes in the distribution of CO₂RR products is attributable to changes in Cu surface topography created by the plasma pretreatments. Roughened Cu surfaces containing a high proportion of under-coordinated Cu sites that exhibit stronger CO adsorption energies than more highly coordinated sites present of planar surfaces. Consistent with this finding, the fraction of CO formed from CO₂ released as CO decreases and the fraction converted to hydrocarbons and alcohols increases with increasing roughness. Our experimental efforts are supported by an analysis of the distribution of sites on a simulated roughened surface of Cu. This work shows that the roughened surface contains a much higher proportion of under-coordinated sites, and in particular sites that adsorb OCCOH, a

suggested precursor to C₂₊ products,⁴¹ more strongly than do the more highly coordinated sites present on a Cu(100) surface.

Surface Characterization before and after Plasma Pretreatment. Panels A and B of Figure 1 show atomic force microscopy (AFM) and scanning electron microscopy (SEM) images, respectively, of the electrochemically polished Cu foil. Figure 1C presents the AFM topography image of Cu foil after 10 min of Ar plasma treatment. In contrast to the polished foil, the Ar plasma-treated foil is much rougher, exhibiting extensive ridges and valleys.⁴² The arithmetic surface roughness factor (S_a) increases from 1.28 nm for the electropolished foil to 7.68 nm after Ar⁺ sputtering based on an AFM scan of a 500 × 500 nm² area (Figure S1). SEM images of Ar plasma treated Cu foil before and after CO₂RR electrolysis are shown in Figure 1D. In agreement with the topography determined by AFM, the 2D projection exhibits a surface structure covered with pits and islands generated by Ar⁺ bombarding and redeposition of Cu atoms. Similar surface roughening effects were also observed for N₂ and O₂ plasma-treated Cu foils (Figure S2).

To further quantify the surface roughness of Cu foil electrodes, we determined the electrochemically active surface area (ECSA) by measuring the double-layer capacitances and then calculating the relative roughness of plasma-pretreated Cu compared to electropolished Cu (Table 1 and Figure S3). In general, all of the plasma treatments increased the roughness of Cu, and prolonged pretreatment time led to a more roughened surface. For the same pretreatment duration, Cu foils treated in either an Ar or O₂ plasma were rougher than that exposed to an

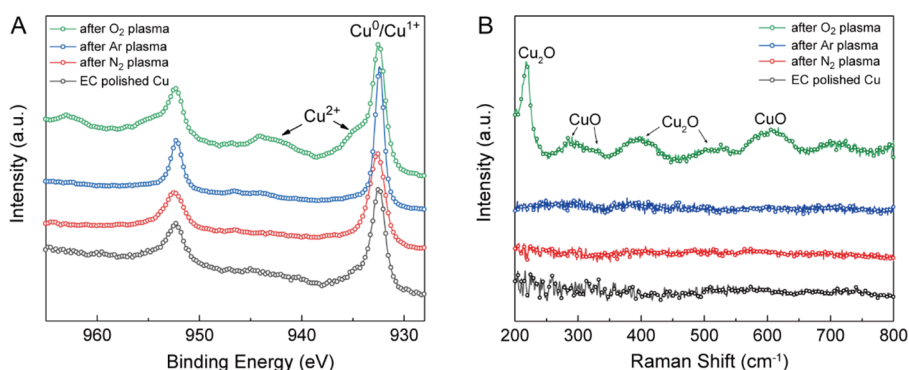


Figure 2. Chemical state characterization of Cu electrodes. *Ex situ* (A) XPS and (B) Raman spectra recorded on Cu foils after 10 min plasma treatments under different gaseous atmospheres.

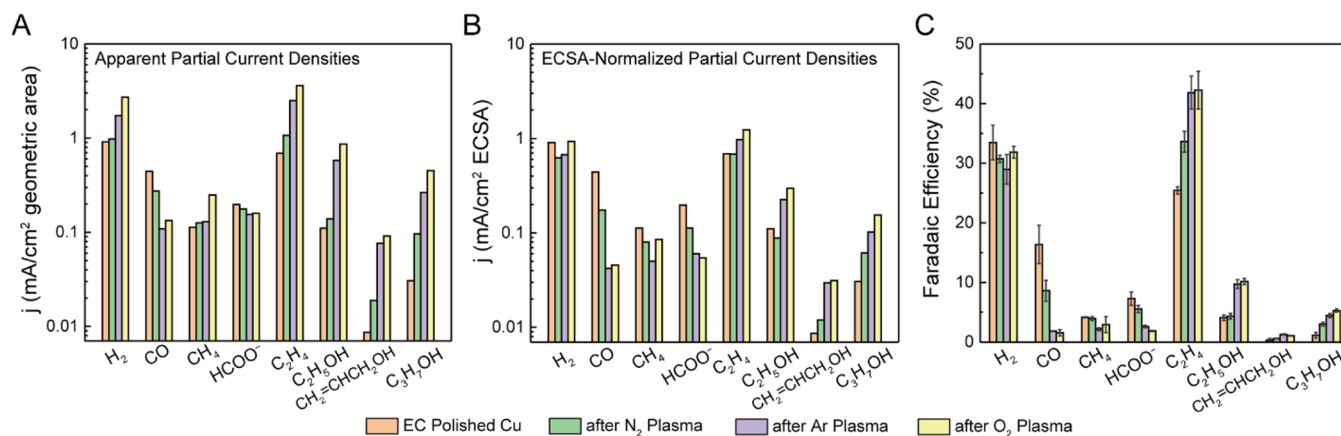


Figure 3. Electrochemical reduction of CO₂ over Cu foil electrodes after 10 min plasma pretreatments under different gaseous atmospheres. (A) Geometric area normalized, (B) ECSA-normalized partial current densities, and (C) selectivity of major products during CO₂RR over Cu foil electrodes after 10 min of plasma pretreatments under different atmosphere.

N₂ plasma. These differences are likely due to the larger ion size of Ar⁺, the more aggressive etching of O²⁻,^{43,44} as well as the pronounced surface structural rearrangement caused by removal of oxygen from copper oxide during electrochemical reduction.^{26,34} We also note that plasma pretreatment in an N₂ plasma for more than 10 min did not increase the surface roughness of Cu further, in contrast to what was observed for pretreatment in an O₂ plasma.

Figure 2A shows the *ex situ* core-level X-ray photoelectron spectra (XPS) of the Cu 2p region after plasma treatments in different atmospheres. All four samples show the predominant peak at 932.4 (Cu 2p_{3/2}) and 952.2 eV (Cu 2p_{1/2}), corresponding to Cu(0) or Cu(I). For the O₂ plasma-treated Cu, two other satellite peaks show up at ~935 and ~944 eV, which are assigned to Cu(II) species. Evidence for Cu(II) cation was also obtained from the Cu LMM region of the Auger spectrum (see Figure S4). Raman spectra of these four Cu samples are shown in Figure 2B. No obvious Raman scattering feature is observed for electropolished Cu or Cu exposed to an N₂ or Ar plasma. By contrast, features for both Cu₂O (~218, 402, and 526 cm⁻¹) and CuO (~290, 332, and 617 cm⁻¹) are observed after O₂ plasma pretreatment,⁴⁵ in agreement with XPS spectra and Raman studies reported previously.^{28,32}

Electrochemical Activity and Selectivity. The electrochemical CO₂RR performance of plasma-pretreated Cu foils was evaluated by 1 h chronoamperometric electrolysis at -1.0 V vs RHE, for which the CO₂ consumption rate is below 17.0 nmol s⁻¹ cm⁻² (Figure S5).⁴⁶ The 0.1 M CsHCO₃ was employed as

the supporting electrolyte, based upon our previous studies showing that Cs⁺ cations enhance the field stabilization of the intermediates critical to the formation C₂ products^{13,14,47} (further evidence for influence of cation identity is given in Figure S6). The superficial current densities shown in Figure 3A all increase with plasma pretreatment, in the order of the increasing ECSA. To account for this effect, Figure 3B shows the current densities for all four samples normalized by the ECSA. The ECSA-normalized current densities for H₂ and CH₄ are not strongly changed by plasma pretreatment, whereas those for CO and HCOO⁻ decrease and those for all C₂₊ product increases in the order no pretreatment < N₂ plasma pretreatment < Ar plasma pretreatment < O₂ plasma pretreatment. We note in particular that the rate of CO evolution decreases by more than an order of magnitude upon Ar or O₂ plasma-pretreated Cu compared to that for electropolished Cu, and the rate of C₃ products generation—allyl alcohol and *n*-propanol—increases by a factor of 3–5. The FEs of the principal products of the CO₂RR generated on electropolished and plasma pretreated Cu are illustrated in Figure 3C. After 10 min of plasma pretreatment, the selectivity toward hydrogen evolution (HER) and methane generation does not change very much, while the FEs for CO and HCOO⁻ decrease in the order of polished Cu > N₂ plasma pretreated Cu > Ar plasma pretreated Cu ≈ O₂ plasma pretreated Cu. By contrast, the FEs for C₂₊ products—C₂H₄, C₂H₅OH, and *n*-propanol—follow the reverse trend.

The similarity of both C₂₊ product selectivities and specific activities of Cu after Ar or O₂ pretreatment is particularly

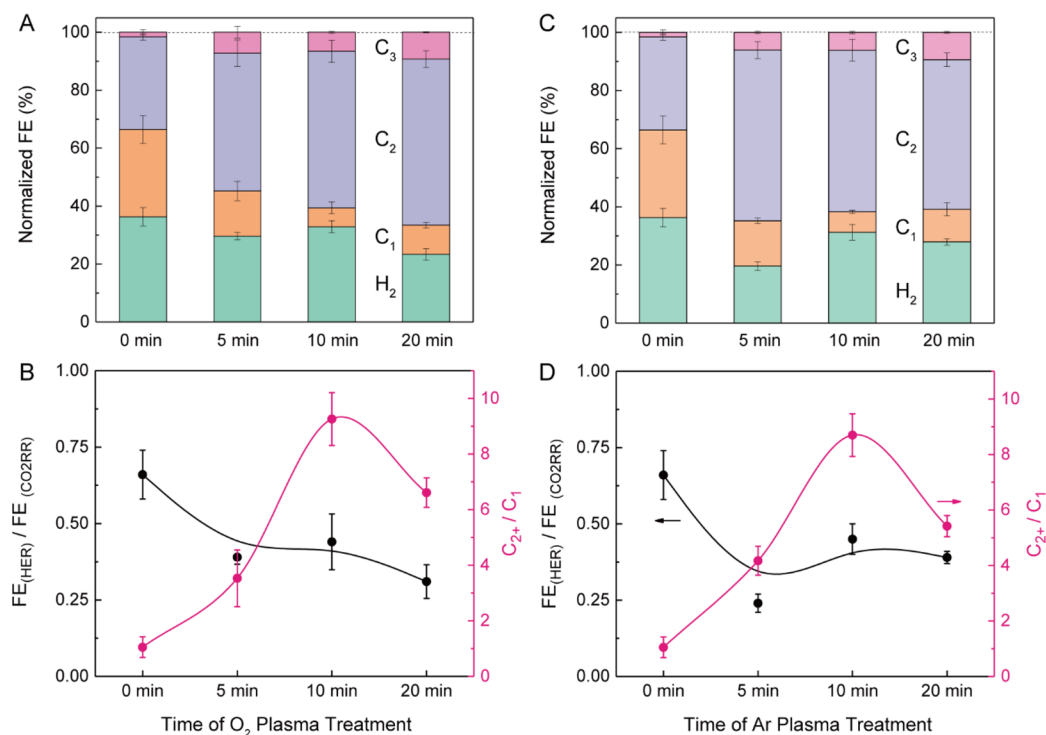


Figure 4. CO₂RR products distribution as a function of plasma treatment time. Both (A and B) O₂ and (C and D) Ar atmosphere plasma pretreatments were plotted. The sum of as-determined overall FEs are varied from 92.2% to 98.9% and all normalized to 100% in panels A and C for comparison.

interesting. Therefore, we investigated the effects of the duration of plasma pretreatment in Ar and O₂ on the distribution of CO₂RR products. Figure 4 shows these results for durations of 0, 5, 10, and 20 min, with the electropolished sample referenced as 0 min. The normalized FEs (referenced to 100%) are given in Figure 4A,C. The FE for HER is the cyan column, the FE for C₁ products (HCOO⁻, CO and CH₄) the orange column, the FE for C₂ products (C₂H₄ and C₂H₅OH) the purple column, and the FE for C₃ products (allyl alcohol and *n*-propanol) the magenta column. The selectivity ratio of HER/CO₂RR (black line) and the C₂₊/C₁ ratio (red line) are illustrated in Figure 4B,D. We observe similar product distributions throughout the time-course for both O₂ and Ar plasma pretreatments, which could largely rule out the (sub)surface oxygen effect. HER decreases to below 30% after 5 min of plasma exposure in both cases and levels off with prolonged exposure. The FE for C₁ products also decreases from 0 to 10 min, and it then increases again because of enhanced methane evolution after a 20 min pretreatment. Over the same pretreatment time, the FE for C₃ alcohols reaches ~10%, whereas the FE for C₂₊ products rises to ~60%. The ratio of C₂₊/C₁ reaches a maximum value of ~9 after 10 min of pretreatment in either an Ar or O₂ plasma.

Because the Cu surface roughness increases with prolonged plasma-treatment time, we assessed whether the changes shown in Figure 4 correlate with surface roughness. The choice of which product ratios to plot is guided by recent theoretical studies of the CO₂RR mechanism, which suggest that 2e⁻ products (HCOO⁻ and CO_g) are produced via the adsorbed intermediates *HCOO and *COOH, respectively, whereas both CH₄ and C₂₊ products are produced via the reduction of *CO.^{5,6,48–51} Therefore, as illustrated in Figure 5A, *CO serves as the key intermediate to CH₄ and C₂₊ products. Panels B and C of Figure 5 plot the selectivity ratios for C₂₊ products versus C₁

products (HCOO⁻, CO, and CH₄) and versus CH₄ alone, respectively. The C₂₊/C₁ ratio reaches a value of ~9 for a roughness of ~3 and then decreases thereafter, regardless of whether Cu is pretreated in an Ar or O₂ plasma. By contrast, the C₂₊/CH₄ ratio increases monotonically to a value of 30 in the case of Ar plasma pretreatment but reaches a maximum value of ~20 and then decreases in the case of O₂ pretreatment.

Figure 5D shows that the fraction of CO produced by the CO₂RR converted to CH₄ and C₂₊ products increases from 0.68 for a roughness of 1.0 to 0.96 for a roughness of 1.5. Further increasing the roughness to ~4.0 increases this ratio to 0.98. The trend observed in Figure 5D suggests that the binding energy for CO increases monotonically with surface roughness, consistent with temperature-programmed desorption experiments²⁶ and recent theoretical calculations on simulated Cu particle surface.^{41,52} The theoretical work also predicts that the stability of *OC-COH, a critical precursor to C₂ products, increases with surface roughness and in particular with the formation of concave sites at grain boundaries between Cu(100) and Cu(111) surfaces. The downturn in the ratio of FEs for C₂₊ product formation and CH₄ formation seen in Figure 5C when the roughness exceeds ~3.0 for O₂ plasma pretreated Cu might be attributable to the formation of subnanometric Cu clusters (i.e., surface dimer and trimer adatoms) that could serve as active sites for selective CO₂-to-CH₄ conversion.⁵³ Finally, Figure 5E shows the surface roughness dependence of the ratio of FEs for C₂₊ oxygenated products to that for C₂₊ hydrocarbons. The selectivity to C₂₊ oxygenated products (ethanol, allyl alcohol, and *n*-propanol) increases by a factor of 2 with increasing roughness, most likely because of the lower surface coverage by *H, as reflected in the HER trend of Figure 4, and the reduced likelihood of hydrogenating C-C intermediates on the roughened Cu surfaces. Similarly, the selectivity to HCOO⁻

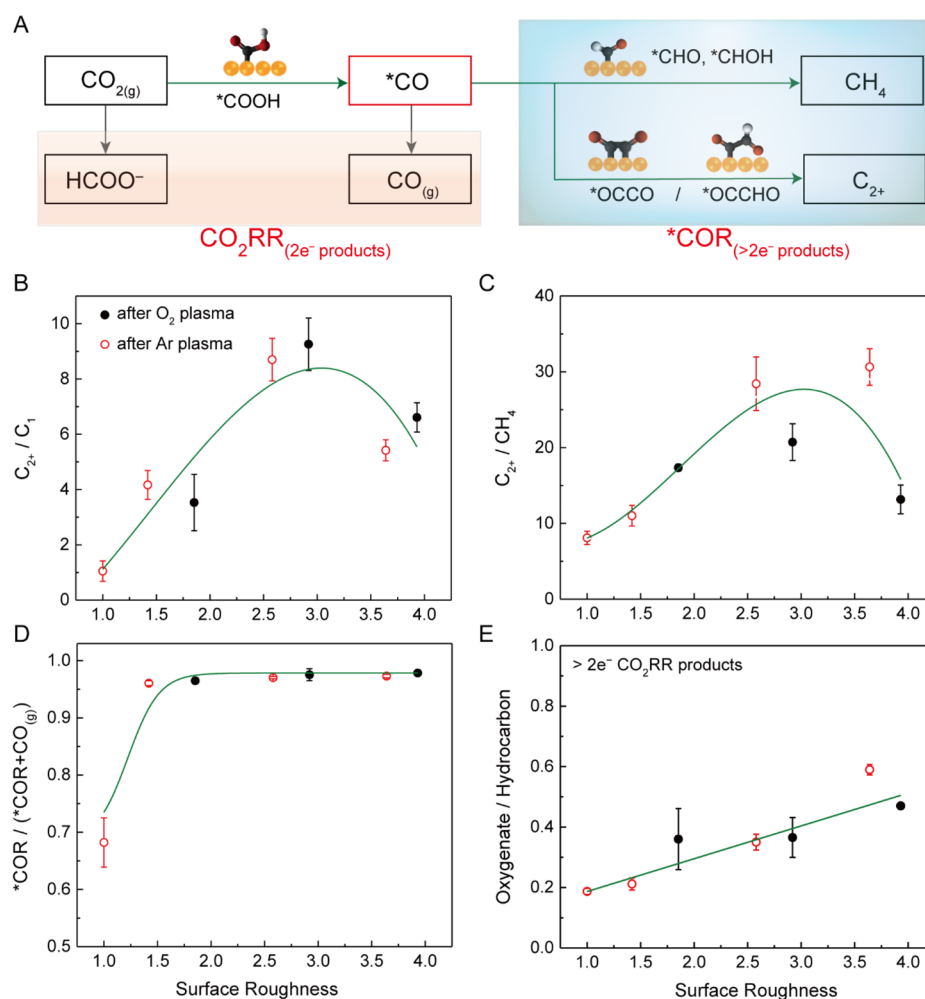


Figure 5. CO₂RR products distribution as a function of Cu surface roughness. (A) Simplified flowchart of CO₂RR mechanism leading to C₁ and C₂₊ product generation. * corresponds to adsorbed species, > 2e reduction products of CH₄ and C₂₊ that go through the reduction of *CO intermediate (*COR) are marked in light blue. (B) C₂₊/C₁ ratio, (C) C₂₊/CH₄ ratio, (D) *COR/(CO_(g) + *COR) ratio, and (E) oxygenate/hydrocarbon (only >2e reduction products) ratio.

decreases as the roughness increases from 1.0 to 2.6 and levels off thereafter (Figure S7). We suggest that this trend is due to the lower availability of terrace sites on roughened Cu surfaces, which are needed to bond bidentate *HCOO, the precursor to HCOOH and hence HCOO⁻ upon desorption of HCOOH into the alkaline electrolyte.⁶

A further factor that can contribute to enhancing the ratio of C₂₊/CH₄ on roughened versus smooth Cu surfaces is the higher local pH near the surface of the roughened Cu; therefore, we considered the possible effect of pH on the observed results. The higher pH near the roughened surface can be ascribed to the higher rate of OH⁻ generation rate per geometric electrode area. We note that this proposal is consistent with the control experiment presented in Figure S8, illustrating the effects of increased bicarbonate concentration, and the earlier findings of Hori et al., who reported a 4-fold enhancement in the C₂₊/CH₄ ratio upon increasing the surface pH from 8.5 to 9.5.⁵⁴ In our studies, we observed only a ~2.6-fold increase in the C₂₊/CH₄ ratio because of an estimated increase in the surface pH change from 9.76 to 9.88 for Cu treated for 5 min versus 10 min in an Ar plasma. The pH for these experiments was estimated assuming a mass-transfer boundary layer thickness of ~40 μm and a bicarbonate concentration of 0.1 M.^{46,55} We also note that it has

been reported that the increase in surface pH as a consequence of electrolyte polarization should suppress CH₄ production but keep C₂₊ activity constant at a given potential.^{56,57} Figure 3B shows that while the intrinsic activity (ECSA-normalized partial current density) for forming CH₄ decreases following plasma pretreatment, the intrinsic activities for producing C₂₊ products, especially oxygenated products, increases. These observations suggest that the enhanced C₂₊ selectivity observed after plasma treatment is not attributable to the small change in the local pH.

Table S1 compares our results on the effects of surface roughness with those previously reported. Various methods have been used to roughen the surface of Cu. These include reduction of copper oxides produced by thermal and/or plasma oxidation, potential cycling in halogen-containing solutions, dendrite growth, etc. Regardless of the method used to achieve roughening, increased surface roughness generally leads to an enhanced C₂₊ FE and C₂₊/C₁ ratio. Most authors have attributed these effects of enhanced surface roughness to a greater abundance of under-coordinated surface sites and defective sites that bind *CO strongly and promote its further reduction to C–C bond formation. Unfortunately, these results cannot be compared with those reported here because of the

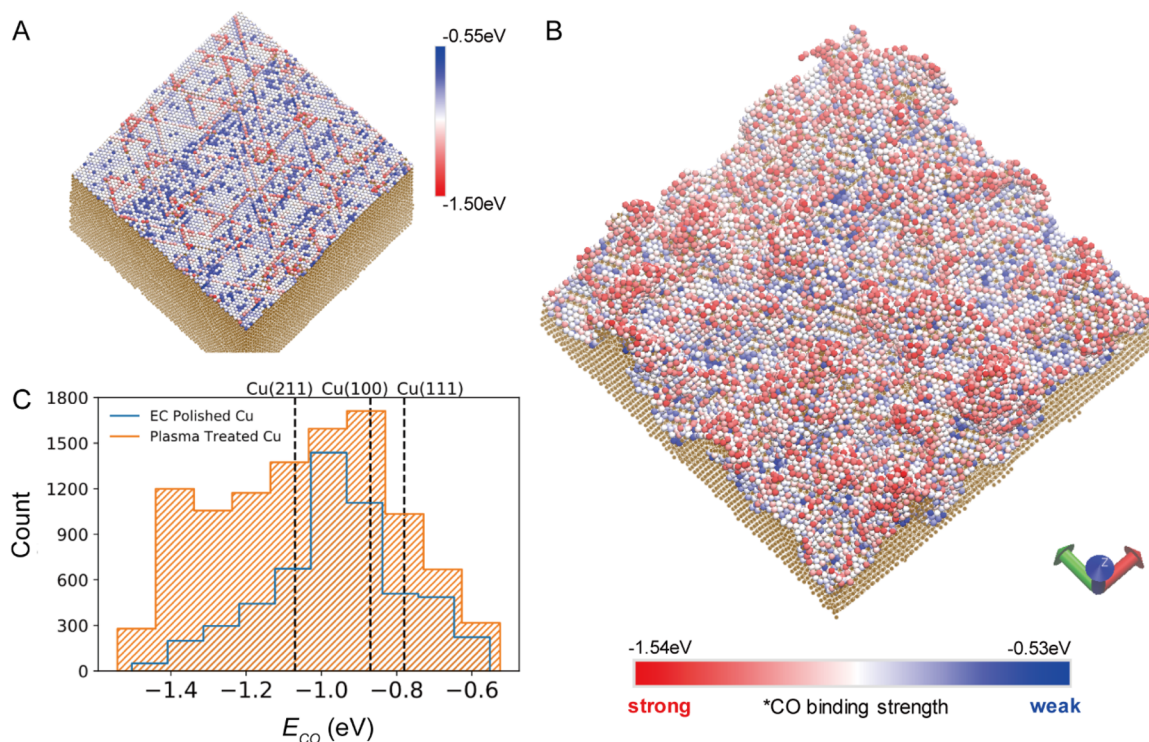


Figure 6. Active site visualization by ReaxQM-Machine Learning. Images of the computationally produced Cu surface of (A) electro-polished and (B) after Ar plasma bombardment and (C) predicted distribution of CO adsorption energies, ΔE_{CO} . The three dashed lines indicate the CO adsorption energies on Cu(111), (100), and (211).

large difference in the modes of sample preparation and conditions for their investigation.

Theoretical Simulation and Analysis. Insights into why surface roughening causes an increase in the formation of C_{2+} products can be gained by atomic-level analysis of the Cu surface created by computational simulation of Ar^+ bombardment roughening of a Cu(111) surface followed by thermal relaxation of the resulting surface. The simulation of roughening begins with a $17.6 \times 17.4 \text{ nm}^2$ surface that is 10.1 nm thick, which contains ~ 5422 total surface atoms (see Figure 6A). Figure 6B and the Supporting Movie 1 show that after bombardment of this surface with 1300 Ar^+ cations, the Cu surface exposes 10 433 atoms, corresponding to an increase in surface roughness of 1.93. The simulated untreated Cu surface and the plasma-treated surface resemble the topography and roughness of the experimental observations, as seen in Figure 1.

The ReaxQM-Machine Learning approach was used to predict the distribution of CO binding energies, ΔE_{CO} , for both simulated electrochemically polished Cu and the Ar plasma roughened surfaces.^{41,58} The electropolished Cu, shown in Figure 6A, is dominated by close-packed low index sites, and therefore, the mean values of ΔE_{CO} are close to those of Cu(111), (100), and (211) surfaces (Figure 6C). By contrast, a much higher population of stronger CO binding sites, ranging from -1.10 to -1.54 eV (red sites along the ridges in Figure 6B and the columns left of the Cu(211) line in Figure 6C), appear after plasma pretreatment. The increased number of strong CO binding sites is consistent with the experimental observation of a reduced selectivity to CO formation following plasma pretreatment to roughen the catalyst surface. Because the adsorption of *CO and *H are expected to compete on the surface of Cu, the stronger adsorption of CO on roughened Cu is expected to result in a reduction of the $\text{*H}/\text{*CO}$. This reasoning would

explain the increase in the formation of oxygenated relative to hydrocarbon products observed with increasing roughening seen in Figure 5E.

In our previous work we have employed the formation energy of *OC-COH (ΔE_{OCCOH}) as a descriptor for the selectivity of C_{2+} products formed during the CO_2RR and have shown that the mean value of ΔE_{OCCOH} decreased from $\sim 1.35 \text{ eV}$ for random surface sites to $\sim 0.50 \text{ eV}$ for surface square sites.³⁹ We also noted that concave defects located at the Cu(100) plane next to a (111) step exhibited the most favorable values of ΔE_{OCCOH} for C_{2+} selectivity. In the present study, we identified the surface square sites on the simulated Cu surfaces for comparison, as shown in Figure 7A,B. The percentage of square sites to overall surface atoms increased from 20.9% on the polished Cu surface to 38.5% on the surface produced by Ar plasma pretreatment. Figure 7C shows the atomic structures of randomly chosen square sites occurring on the simulated surface produced by plasma treatment. The sites are similar to the Cu(100) configuration but have abundant neighboring step sites. As noted, the significantly increased percentage of surface square sites favors C-C bond formation and provides a rationale for the enhanced C_{2+} selectivity that correlates with increased surface roughening caused by plasma pretreatment.

In conclusion, the present study investigated the effects of surface topography on the activity and product selectivity of electrochemical CO_2 reaction, demonstrating the distribution of products produced by CO_2RR on metallic Cu changes with Cu surface topography created by plasma pretreatment. The activity and product selectivity of electrochemically polished Cu was compared with those of roughened Cu surfaces prepared by different plasma pretreatments. The differences in overall CO_2RR activity from different plasma pretreatments are attributable to changes in the electrochemically active surface

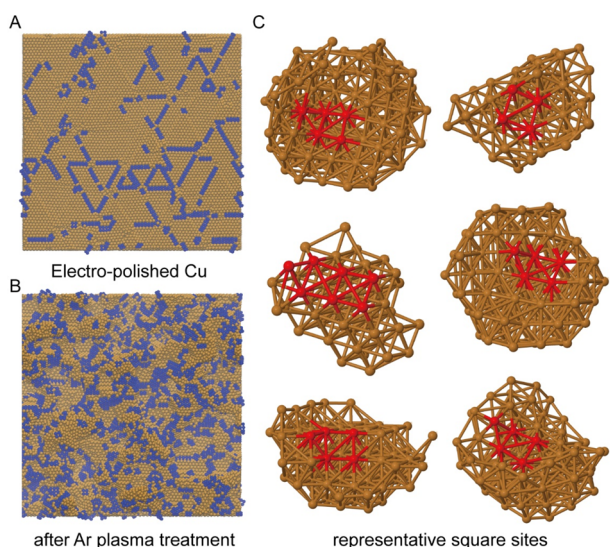


Figure 7. Visualization of square sites on computationally simulated Cu surfaces. (A) Electropolished Cu and (B) Ar plasma pretreated Cu. The surface square sites are marked in blue. (C) Zoomed-in atomic structure of randomly chosen square sites on the plasma-treated Cu surface, both the square sites and the neighboring Cu step sites (if any) are highlighted in red.

area. Of particular note, we observe that with increasing surface roughness, the ratio of current densities for CO_2RR to HER, the ratio of the current densities for COR to the sum of the current densities for CO formation and COR, and the ratio of current densities for formation oxygenates to hydrocarbons all increase. These trends are interpreted based on an atomic-level analysis of the topography of roughened Cu surface. We find that CO formed by the electrochemical reduction of CO_2 binds more strongly on the roughened surface than on the electrochemically polished surface, suggesting that the ratio of adsorbed H to CO decreases with increasing surface roughness. These trends explain why the fraction of CO produced by CO_2 reduction converted to reduced products and the fraction of these products appearing as oxygenates rather than hydrocarbons increase with increasing surface roughness. A further effect of increasing surface roughness is an increase in the fraction of final products (oxygenates and hydrocarbons) containing two or more C atoms. Our analysis shows that surface roughening increases the fraction of square sites similar to those on a Cu(100) surface but having abundant neighboring step sites. Prior work has shown that such sites enhance the formation of C–C bonds required to form C_{2+} products during the electrochemical reduction of CO_2 . Finally, we find that increased roughening of the Cu surface increases the ratio of current densities for C_{2+} to C_1 (CO, HCOO^- , and CH_4) products up to a maximum value of ~ 9 for a surface roughness of ~ 3 . Most of this downturn is ascribed to the enhanced formation of CH_4 on surfaces with roughness in excess of 3. The enhanced formation of methane on highly roughened surfaces is attributed to the formation of surface dimer and trimer clusters on the Cu surface that have been shown by theoretical calculation to serve as active sites for selective conversion of CO_2 to CH_4 . In summary, the results of the present work highlight the importance of surface topography and defect sites on the observed CO_2 reduction activity and selectivity and suggest that rational surface structure engineering could contribute to the development of Cu

electrocatalysts exhibiting a high selectivity to multicarbon products.

Supporting Information

The Supporting Information is available free of charge at <https://pubs.acs.org/doi/10.1021/acscenergylett.0c00482>.

Experimental and computational details; additional Cu electrode characterization by AFM, SEM, Auger, XPS, and electrochemical double-layer capacitance measurements; Cu surface roughness dependence of CO_2 consumption rate and $\text{FE}_{\text{formate}}/\text{FE}_{(\text{CO}_g+\text{COR})}$ ratio during chronoamperometric electrolysis; online differential electrochemical mass spectrometry study on the cation effect; control experiments study on anion buffering effect and extra surface roughening at the presence of iodide anion (PDF)

Supporting Movie 1: Cu surface after Ar plasma bombardment (MP4)

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Notes

The authors declare no competing financial interest.

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