Title
Real-time high-resolution X-ray imaging and nuclear magnetic resonance study of the hydration of pure and Na-doped C3A in the presence of sulfates

Permalink
https://escholarship.org/uc/item/2df633g9

Author
Kirchheim,, A. P.

Publication Date
2011-01-19

DOI
10.1021/ic101460z
Real-time high-resolution X-ray imaging of the hydration of pure and Na-doped C₃A in different suspensions


ᵃDepartment of Civil Engineering
Federal University of Rio Grande do Sul, Porto Alegre, RS, Brazil
ᵇCenter for X-ray Optics, Lawrence Berkeley National Laboratory, Berkeley, CA USA
ᶜNMR Core lab, King Abdullah University of Science and Technology, Thuwal 23955-6900, Kingdom of Saudi Arabia,
ᵈDepartment of Chemical & Biomolecular Engineering, University of Melbourne, Victoria 3010, Australia
ᵉDepartment of Civil and Environmental Engineering
University of California, Berkeley, CA, USA

ABSTRACT
This study details the differences in real-time hydration between pure (cubic) and Na-doped (orthorhombic) C₃A. First, pure phases were synthesized in the laboratory to develop an independent benchmark for the reactions, then the reactions of others phases were isolated. Because the kinetics of this reaction moves extremely quickly, most methods are not adequate to study the reactions in the early phases. Here, a high-resolution full-field soft X-ray imaging technique operating in the X-ray water window was used to capture the mechanism of the C₃A hydration. The results show that there are differences in the hydration mechanism of each type of C₃A, dependent on the concentration of sulfates ions in solution. The reactions with cubic C₃A (pure) seems to be more susceptible to higher concentrations of sulfate ions by forming smaller ettringite needles at a slower pace than the orthorhombic C₃A (Na-doped) sample.

Keywords: tricalcium aluminate, hydration, hydration products, ettringite, image analysis

1. Introduction
The properties of clinkers are strongly affected by the quantity and composition of the clinker phases [1], which can drive the kinetic process of hydration. The hydration of C₃A is an important stage in the overall hydration of cement because it has considerable influence on its setting and early hardening. The hydration of C₃A is far more strongly affected by environment conditions than that of C₃S and, in particular, by the presence of other substances with which it can react [8]. The reaction of C₃A with water is a primary consideration. Calcium aluminate hydrates (e.g., C₃AH₆, C₄AH₁₉, and C₂AH₈) form quickly and liberate large amounts of heat [9]. If the very rapid and exothermic hydration of C₃A is allowed to proceed unhindered in cement, then the setting occurs too quickly and the cement paste does not develop strength. Therefore, calcium sulfate is added to slow down the C₃A hydration. In the presence of gypsum, C₃A forms needle-like ettringite. Thus, for practical purposes, it is not the hydration reaction of C₃A alone that is important but the hydration reaction of C₃A in presence of gypsum (sulfates). The reactivity of C₃A in clinker and the availability of sulfate in solution control the setting characteristics of concrete, making these compounds ideal for addressing probable setting problems.

Because of the high contents of C₃A in some cement, special attention must be paid to the mineralogy and reactivity of this phase. Industrial clinkers and their phases contain minor elements such as Na₂O, K₂O, SO₃, and Cl, which may be attributed to several sources, such as raw materials, fuel and/or the lining of the kilns. Alkalis can be incorporated into a number of phases of the clinker and often Na₂O is

* Corresponding author: Tel.: +55 51 3308 3518; Fax: +55 51 3308 3321. E-mail address: anapaula.k@gmail.com
taken up by the C$_3$A [2]. When C$_3$A is synthesized in the presence of alkalis, its crystal lattice changes and the formation of the other phases occur [3]. There are several series of solid solutions with the general formula Na$_2$xCa$_{3-x}$Al$_2$O$_6$, including a cubic C$_3$A with 0<x<0.10, an orthorhombic C$_3$A with 0.16<x<0.20, and a monoclinic C$_3$A with 0.20<x<0.25 [4]. The most important solid solutions series occurring in this system (Na$_2$xCa$_{3-x}$Al$_2$O$_6$) are the orthorhombic and monoclinic [5].

In Portland cements, the cubic or orthorhombic C$_3$A are found alone or in combination; the monoclinic modification has not been observed. The cubic phase is often finely grained and closely mixed with dendritic crystals of ferrite. The orthorhombic modification occurs as a prismatic, dark interstitial material, and it is sometimes pseudo tetragonal [4]. In the orthorhombic structure, the replacement of Na$^+$ at a calcium site can only occur at the Ca$_5$ site, which causes the polyhedron to expand [6]. In industrial clinkers, a more typical mechanism of formation of this phase is the metastable extension of the solid solution series accompanied by rapid cooling [5]. In cement, higher percentages of Na$_2$O in the raw materials lead to formation of orthorhombic C$_3$A [7].

This paper aims to clarify the early hydration of cubic and orthorhombic C$_3$A using advanced in-situ technique by means of soft X-ray microscopy, which allows studying hydration in a wet environment with nanometer spatial resolution. The present study uses pure components of cubic C$_3$A (pure) and orthorhombic C$_3$A (Na-doped) synthesized in a laboratory. This allows for isolating and observing the effects of each crystalline form of the aluminate in a given solution, which would not be possible if cement particles (containing several different phases) were used. In addition, different concentrations of sulfate ions were analyzed to understand how each aluminate works in different conditions. This study is part of a larger work [10] that includes the study of these materials in other conditions and using other techniques. Also some previous results were published by Kirchheim et al [11].

2. Experimental Section

2.1 Materials

Samples of tricalcium aluminate (C$_3$A)—orthorhombic (Na-doped) and cubic (pure)—were obtained from Construction Technology Laboratories, Inc., Skokie, IL. Both compounds were synthesized in a laboratory by heating a stoichiometric blend of reagent grade CaCO$_3$ and alumina (Al$_2$O$_3$) in an electric furnace at 1400 °C for 1 hour, followed by quenching in air. The orthorhombic C$_3$A was prepared from reagent grade CaCO$_3$, Al$_2$O$_3$, and Na$_2$CO$_3$ in stoichiometric proportion similar to that reported by Regourd et al [12], with the latter slightly added in excess to account for alkali volatilization during the synthesizing process. After synthesizing, each material was processed in a ceramic mill to 325 mesh.

XRD, using a PANalytical Xpert Pro and thermal analysis (TA) TA Instruments, model SDT 2960, were done to confirm phase composition, purity and conditions of the sample. Comparison of the peak positions were made with data from Regourd et al [12] (ICDD PDF No.38-1429 for the cubic sample and ICDD PDF No.26-0958 for the orthorhombic one). The TA analysis showed that the sample were keep no hydrated by the air moisture because it presented a low loss of mass – less then 1% and 8%, respectively for cubic and orthorhombic sample.

Particle size distributions of the samples were done in a CILAS 1180 Laser Granulometer, in isopropyl alcohol of 99.5% of purity based liquid media; the results are presented in Table 1. The lower size of this equipment was 40nm and was applied the Fraunhofer Approximation.

**Table 1**: particle size distribution of the sample by means of laser granulometry

<table>
<thead>
<tr>
<th></th>
<th>Cubic C$_3$A (µm)</th>
<th>Orthorhombic C$_3$A (µm)</th>
</tr>
</thead>
<tbody>
<tr>
<td>10% diameter</td>
<td>5,74</td>
<td>1,64</td>
</tr>
<tr>
<td>50% diameter</td>
<td>24,31</td>
<td>16,04</td>
</tr>
<tr>
<td>90% diameter</td>
<td>45,00</td>
<td>36,34</td>
</tr>
<tr>
<td>mean diameter</td>
<td>25,16</td>
<td>18,43</td>
</tr>
</tbody>
</table>

2.2 Sample preparation
2.2.1 Imaging with Soft X-Ray Microscopy

Soft X-ray microscopy operating in the water window, i.e. at a photon wavelength of about 2.4nm, combines high spatial resolution in the tens of \( \text{nm} \) range with the ability to penetrate several \( \mu \text{m} \) of aqueous solution, making it an ideal \textit{in-situ} technique to study wet and nano-structured materials. The high spatial resolution of this microscopy technique is well suited for observing the formation of ettringite crystals around the calcium aluminate grains and in solution, and to see the differences between the hydrates formed from cubic and orthorhombic \( \text{C}_3\text{A} \). Furthermore, because the sample is kept wet at all times, it is possible to observe the progression of hydration over time.

Real-time hydration studies with soft X-Ray transmission microscopy was done at the full field soft X-ray microscopy end-station XM-1 located at the beam line 6.1.2 at the Advanced Light Source (ALS) in Berkeley CA. The microscope is jointly operated by the Center for X-ray Optics and the ALS and the X-ray optical setup of XM-1 is described elsewhere [13]. Recent achievements in using Fresnel zone plates (FZP) for X-ray optics allow for a spatial resolution down to 15nm [14, 15]. Synchrotron radiation emitted from a bending magnet is passed through the first FZP, a condenser zone plate, which provides both a partially coherent hollow-cone illumination of the specimen, and, in combination with a pinhole, serves as linear monochromator to select the proper wavelength. After transmitting the sample, a micro zone plate (MZP) projects a full-field image onto an X-ray sensitive CCD camera. For this study, a 35nm MZP was used and imaged at a magnification of about 2000. The illumination time per image was between 1-14 seconds. Each image covers a field of view of about 10\( \mu \text{m} \).

The liquid media, composed of a saturated solution with calcium hydroxide (CH) and calcium sulfate dehydrate (gypsum, \( \text{C}_2\text{SH}_2 \)), was selected to provide sulfate ions to stimulate both the formation of ettringite and the chemical composition of the pore solution in fresh cement pastes. After mixing gypsum and calcium hydroxide for 24 hours the solution was filtered twice. To avoid carbonation, the solution was prepared and stored in a glove bag filled with nitrogen gas. Also to avoid the possibility of an alkali-silica reaction, polyethylene and Teflon beakers, flasks, and pipettes were used. Dispersions of both \( \text{C}_3\text{A} \) forms were prepared with the saturated solution, which presented a solution/aluminate ratio of 5 ml/g, 10 ml/g, and 50 ml/g. The amounts of reagents used to prepare the dispersions were 0.4g, 0.2g and 0.04g of \( \text{C}_3\text{A} \) and 2 ml of the saturated solution, respectively. The solid particles were hand-mixed for 50 seconds in the solution and then centrifuged for 15 seconds. Therefore, in this study the solution/aluminate ratio is called the initial solution/aluminate ratio (s/\( \text{a}_{\text{initial}} \)). During this first minute of hydration, it is expected that aluminum and calcium-bearing ions are released into the solution from the dissolution of the \( \text{C}_3\text{A} \) particles. The number of nuclei formed during the initial period will be affected by the solution/aluminate ratio; the number of nuclei formed initially may then play an important role in determining the growth rates of ettringite or other hydrates after dilution. To allow for sufficient transmission of the soft X-rays, a small droplet—containing small solid particles—was taken from the supernatant solution and assembled in a specialized sample holder, comprising of two silicon nitride windows, where the solution is squeezed in between to an appropriate thickness and then imaged. This procedure is described in detail in Silva and Monteiro [16] and Monteiro et al [17].

3. Results and discussion

The hydrated products were observed to form within a few minutes after mixing the cubic and three different initial solution/aluminates. Results are presented below.

3.1 Initial solution/aluminate ratio = 5 ml/g

- **Cubic \( \text{C}_3\text{A} \)**

  Figure 1 shows images of the cubic \( \text{C}_3\text{A} \) hydration in a saturated solution of calcium hydroxide and gypsum; the image of the sample was taken in the same specimens at two different positions. The images from Position 1, recorded from 18 minutes to 1 hour and 41 minutes, show the cubic \( \text{C}_3\text{A} \) particle
(indicated by arrows) covered by hydration products. The images of hydration evolution in Position 2, recorded from 50 minutes to 1 hour and 28 minutes, show an ettringite agglomeration. In a few minutes all particles completely hydrate inside their original boundaries. Other 3 specimens were analyzed and also presented the same morphology.

**Position 1**

(a) 18 min  (b) 56 min  (c) 1h 41 min

**Position 2**

(a) 50 min  (b) 1 h 8 min  (c) 1h 28 min

In all images, the products of hydration are fibrous-like. The morphology identified suggests that crystals are acicular ettringite, which precipitated instantly. Silva and Monteiro [16], however, observed that cubic C₃A particles in the same type of solution resulted in C₃A particles being covered by a continuous or discontinuous hydrated layer. X-ray diffraction and thermal analyses pointed by the authors revealed that the sample still contained 31% of Ca(OH)₂, which was originally used as a raw material for synthesizing the C₃A.

- **Orthorhombic C₃A**

Figure 2 shows large long fibers-like crystals [needles and hexagonal platelets that can be seen in a plan view (note the appearance of transparent hexagons) or a lateral view (note the long needles)], with a length of 3 to 5 μm, in a solution with several orthorhombic C₃A particles. The particles seem to be surrounded by a gel (indicated by the arrows). These particles in the form of needles and platelets...
disappear over time; after 49 minutes are no longer present in the sample. Note that the orthorhombic C₃A particles suffered substantial swelling during hydration (indicated by the circumference and the arrows). It is also observed an agglomeration/approximation of the particles over time.

![Figure 2 – in-situ soft x-ray images of hydrating orthorhombic C₃A particles in a saturated calcium hydroxide-gypsum solution, s/a_initial= 5 ml/g. Hydration time is indicated. Sample 1. Scale bar corresponds to 1 μm](image)

In the orthorhombic specimen, the reactions evolved as expected until at 50 minutes of hydration; no modifications were observed after this period. Although the aspect of the orthorhombic C₃A hydration in the same suspension (s/a_initial=5 ml/g) is completely different than the cubic C₃A hydration previously shown, it does not necessarily mean that the cubic is more reactive, however, it does mean it works in a different way, principally by means of the morphology of the particles and products formed. This differences can be explained based in Glasser and Marinho [20] ideas, which proposed that the hydration of Na-doped C₃A (orthorhombic C₃A) can be summarized as the broken off the Na-O bonds which lead to the formation of mobile Na⁺ ions where Ca²⁺ is only partially hydrolyzed in this zone. In the vicinity or the alumina-rich zone, more general hydrolysis occur and oxide ions, formerly linked to Ca²⁺ an Al³⁺, convert to OH⁻ groups, which together with appropriate cations are available to participate in the reconstituted surface film. Ions steadily flow into the solution, whose ionic strength and bascicity increases at least up to the point at which precipitations begins.

*b) Initial solution/aluminate ratio = 10 ml/g*
In order to evaluate the effect of higher sulfate ions concentrations in solution and to facilitate the observation of the cubic C₃A particles hydrating, the initial solution/aluminate ratio was increased to 10 ml/g.

- **Cubic C₃A**
  In a higher solution/aluminate ratio, unlike the images showed previously (figure 1), these allow us to observe the particles being hydrated over time more clearly and slower.
  From the beginning of the analysis (Figure 4), the particles are immediately covered by needle-like hydrates that grow in length and thickness over time (mainly between 53 minutes and 1 hour and 19 minutes). Changes in the ionic concentration over time might explain the instability of the ettringite and formation of large crystals [16].

**Position 1**

![Figure 4 – in-situ soft x-ray images of hydrating cubic C₃A particles in a saturated calcium hydroxide-gypsum solution, s/ainitial= 5 ml/g. Hydration time is indicated. Sample 1. Scale bar corresponds to 1 µm](image)

- **Orthorhombic C₃A**
  Figure 5 presents two positions in the same sample of orthorhombic C₃A particles hydrating in a saturated calcium hydroxide-gypsum solution. Position 1 shows images from 40 min to up 2 hours. Compared to cubic C₃A hydration, the orthorhombic C₃A seemed to **increase the aspect ratio of** ettringite crystals. The same gel that appeared in Figure 3 can be see again, a possible precursor of ettringite formation (See Figure 5 at Position 2). Some researchers [4, 24] believe that there is a change in the morphology of the ettringite at the end of the induction period. This is initially deposited as a gel-like material and, according
to this view, changes into a crystalline needle-like form. Unlike the crystalline form, ettringite gel is impermeable to water. In a mixture with cement and water, Billingham and Coveney [24] posed that an initial jump in the slurry consistency at early stages is due to the rapid formation of ettringite gel on mixing.

**Position 1**

- (a) 40 min
- (b) 1 h 27 min
- (c) 2 h 28 min

**Position 2**

- (a) 19 min
- (b) 43 min
- (c) 2 hours

Figure 5 – *in-situ* soft x-ray images of hydrating orthorhombic C₃A particles in a saturated calcium hydroxide-gypsum solution, s/a<sub>initial</sub> = 10 ml/g. Hydration time is indicated. Sample 1, Positions 1, 2. Scale bar corresponds to 1 µm

Both forms of C₃A in this suspension (s/a<sub>initial</sub>=10 ml/g) presented differences in morphology when compared to the previous suspension (s/a<sub>initial</sub>=5 ml/g); this can be attributed to the increased concentration of sulfates, which allow the ettringite formation close to the C₃A particles. In summary: (a) the hydration of cubic C₃A hydration shows ettringite needles forming along the surface of the particle, which are smaller than those observed in orthorhombic C₃A; and (b) in orthorhombic C₃A hydration an intermediate gel stage occurred before the formation of long ettringite needles.

c) Initial solution/aluminate ratio = 50 ml/g

To complement the analysis of the effect of sulfate ions concentration and verify further the influence of the initial solution/aluminate ratio in the hydration reaction of cubic and orthorhombic C₃A, it was increased to s/a<sub>initial</sub>=50 ml/g.
- **Cubic C\textsubscript{3}A**

  The aspect of the cubic C\textsubscript{3}A particles hydrating in a higher initial solution/aluminate ratio is presented in Figure 6. Note how different this image is compared to the previous images. Here, it is possible to observe the needles growing and sometimes a complete dissolution of the C\textsubscript{3}A particles occurred, with precipitation of needle-like crystals of ettringite.

  **Position 1**

  (a) 19 min  (b) 53 min  (c) 1 h 1 min

  (d) 1 h 14 min  (e) 2 h 38 min  (f) 3 h 10 min

  Figure 6 – *in-situ* soft x-ray images of hydrating cubic C\textsubscript{3}A particles in a saturated calcium hydroxide-gypsum solution, s/a\textsubscript{initial} = 50 ml/g. Hydration time is indicated. Sample 1, Position 1. Scale bar corresponds to 1 µm

  Position 1 (Figure 6) shows the hydration of the particles starting at 11 minutes, which is first covered by hydrates and then followed by a rapid length and thickness growth up to the first 50 minutes after mixing. After this period, the C\textsubscript{3}A particle (which is completely covered by numerous needles-like hydrates) starts to be transformed in ettringite. From these images, the trough-solution mechanism of ettringite formation can be seen; basically there is interface dislocation of outer products to inner products. The cubic C\textsubscript{3}A particle—initially with a dense structure—at the end of 3 hours is fully transformed into a structure of ettringite needles. The C\textsubscript{3}A particle is completely consumed. The original boundary of the C\textsubscript{3}A particle increases and the observation space is mainly occupied by ettringite needles.

- **Orthorhombic C\textsubscript{3}A**

  Figure 8 presents orthorhombic C3A particles hydrating at 54 minutes to 4 hours in a saturated calcium hydroxide-gypsum solution. Note that the orthorhombic samples exhibit almost identical behavior to that observed for the cubic sample in Figure 7. If the same particle
size is considered, a layer of small needles is readily visible growing in the surface of the particle and some swelling of the particle [as indicate the arrows in image (a) and (c)]. One key difference between the cubic sample and the orthorhombic sample is the appearance of several long needles that form close and around the particle and in solution. Needles can also be observed far away from the particle. Comparing Figs 9 a-c, suggests that some initial products formed away from the surface of the C3A are lost - rather than formed - over time. The growth of acicular hydrates occurs with greater intensity after 1 hour; the reactions apparently stabilized after 2 hours.

**Position 1**

Figure 8 – *in-situ* soft x-ray images of hydrating orthorhombic C₃A particles in a saturated calcium hydroxide-gypsum solution, s/a_{initial} = 50 ml/g. Hydration time is indicated. Sample 1, Position 1. Scale bar corresponds to 1 µm

The through-solution mechanism (dissolution/precipitation) of ettringite formation can be seen in Figure 9. A comparison between the orthorhombic particles shown in Figure 9 with the cubic particles shown in Figure 6 shows that the orthorhombic C₃A form larger and thicker needles-like crystals over time. Observing Figure 8 and 9 we can notice the variation in behavior between smaller and larger grains.

**Position 2**
Contrary to the results obtained in the previous specimens ($s/a_{\text{initial}} = 5$ and $10$ ml/g), in this suspension ($s/a_{\text{initial}} = 50$ ml/g) the through-solution mechanism of ettringite formation can be observed for both specimens (cubic and orthorhombic), varying quite strongly when the initial solution/aluminate ratio is increased. In summary, the through-solution rates can be distinguished from ion exchange and absorption.

Finally, it is interesting to note that in all positions of both the cubic and orthorhombic samples for these conditions, the reactions happen mostly after 50 minutes of hydration. This is explained by the higher interparticle distance and the higher sulfate ions concentration. In a cement matrix a high water cement ratio has been shown to lead to better diffusion of the ions in the pore solution leading to the formation of more ettringite crystals, especially away from the surface of cement particles [9].

Any crystallization process comprise three basic steps: achievement of supersaturation; formation of crystal nuclei, and growth of particles [25]. The difference in chemical potentials becomes the driving force for dissolution (in the case of undersaturation) or precipitation (in the case of supersaturation). Once nuclei form in a supersaturated solution, they begin to grow by accretion and, as a result, the concentration of the remaining materials drops. Thus, there is a competition for material between the processes of nucleation and crystal growth. The more rapid the nucleation, the larger the number of nuclei formed before relief of supersaturation occurs, resulting in final crystal with smaller size. Brown and LaCroix [26] determined that the nucleation in ettringite formation was diffusion controlled. It has also been suggested that the $C_3A$ surface may serve as a vehicle for ettringite nucleation [27]. Silva and
Monteiro [16] also observed the ettringite growth on or near the $C_3A$ surface. On the other hand, Minard et al. [23] assumed that while the slow process of ettringite formation could be limited by ettringite growth rate, its dependence on the surface area of $C_3A$ indicated that it is limited by dissolution of $C_3A$.

Minard et al [23] described that even with different $C_3A$ reactivity, the time necessary to exhaust the totality of the sulfate ions does not evolve in a linear way with the quantity of gypsum initially introduced. In the present research, the hydration of cubic $C_3A$ (pure) seems to be more susceptible to the presence of higher concentration of sulfate ions – considering less $C_3A$ particles in solution for the same initial sulfate concentration. The crystal growth of the hydration products from the cubic sample was more rapid with a larger number of nuclei, which leads to a smaller crystal size, depending on the concentration of sulfate ions. The higher the sulfate concentration, the smaller the size of ettringite needles, the faster the conum of the entire particle of $C_3A$. All orthorhombic $C_3A$ (Na-doped) samples were characterized by the growth of bigger ettringite needles, with the apperance of an intermediate gel stage before the formation of long ettringite needles. For the orthorhombic samples, the higher the sulfate concentration, the smaller are the ettringite size, the longer are the time to exaust the reaction. Corroborating, but in different mixture proportions, Kirchheim et al [11] showed that, in a SEM study, the ettringite crystals, when based on the size and the consumption of $C_3A$ and gypsum, orthorhombic $C_3A$ pastes presented a faster rate of formation of crystals, which were also longer. Real time Rheometric measurements showed that orthorhombic $C_3A$ mixes presented higher increase of stiffness than cubic $C_3A$ at very early stages of hydration which can be adressed to a higher reactivity of the orthorhombic $C_3A$ in presence of sulfates.

5. Conclusion

Analyzing the hydration of cubic and orthorhombic $C_3A$ under certain conditions can be summarized as follows:

- The higher the initial solution/aluminate ratio ($s/a_{initial} = 50 \text{ ml/g}$), the greater the sulfate amount available in the reaction, the lower the number of particles of $C_3A$ in solution, which increased the sulfates/aluminates ratio. This allowed for observation of the complete dissolution of $C_3A$ particles and precipitation of ettringite crystals.
- For the $s/a_{initial} = 10 \text{ ml/g}$ solution, the formation of needles was observed near the surface of particles of cubic $C_3A$ and a gel appeared between the particles of orthorhombic $C_3A$, which was followed by the formation of ettringite needles on this gel.
- For the $s/a_{initial} = 5 \text{ ml/g}$ solution, the products of hydration formed almost instantaneously, for both forms of $C_3A$. The cubic $C_3A$ had grew, instantaneously, aciculares crystals. For the orthorhombic $C_3A$ samples, a gel and hexagonal platelets appeared close to the particles. All reactions had occurred before 50 minutes of hydration.

Soft X-ray microscopy has been shown to be an interesting in-situ techique to study real-time hydration of cement phases, particularly cubic and orthorhombic $C_3A$ providing valuable information about the morphology and growth rate of the hydration products.

Acknowledgments

The authors acknowledge the financial support of PNPD-CAPES (Programa Nacional de Pós-doutorado da Fundação Coordenação de Aperfeiçoamento de Pessoal de Nível Superior – Ministério da Educação – Brasil) and are grateful to Dong-Hyun Kim and Anne Sakdinawat (Center for X-ray Optics) for their assistance in acquiring the X-ray images. The operation of the microscope is supported by the Director, Office of Science, Office of Basic Energy Sciences, Materials Sciences and Engineering Division, of the U.S. Department of Energy under Contract No. DE-AC02-05-CH11231. Paulo Monteiro acknowledges the financial support given by KAUST.

References

DISCLAIMER

This document was prepared as an account of work sponsored by the United States Government. While this document is believed to contain correct information, neither the United States Government nor any agency thereof, nor The Regents of the University of California, nor any of their employees, makes any warranty, express or implied, or assumes any legal responsibility for the accuracy, completeness, or usefulness of any information, apparatus, product, or process disclosed, or represents that its use would not infringe privately owned rights. Reference herein to any specific commercial product, process, or service by its trade name, trademark, manufacturer, or otherwise, does not necessarily constitute or imply its endorsement, recommendation, or favoring by the United States Government or any agency thereof, or The Regents of the University of California. The views and opinions of authors expressed herein do not necessarily state or reflect those of the United States Government or any agency thereof or The Regents of the University of California.