A Hybrid Catalyst-Bonded Membrane Device for Electrochemical Carbon Monoxide Reduction at Different Relative Humidities

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Abstract: A hybrid catalyst-bonded membrane device using gaseous reactants for carbon monoxide reduction (COR) reaction in the cathode chamber, an aqueous electrolyte for oxygen evolution reaction (OER) in the anode chamber, and an anion exchange membrane (AEM) for product separation was constructed. The Cu electrocatalyst was electrodeposited onto the gas diffusion layers (GDLs) and was directly bonded to the AEM by mechanical pressing in the test-bed. The impacts of relative humidity at the cathode inlet on the selectivity and activity of COR were investigated by computational modeling and experimental methods. At a relative humidity of 30%, the Cu-based catalyst in the hybrid device exhibited a total operating current density of 87 mA cm^{-2} at -2.0 V vs. Ag/AgCl reference electrode, a Faradaic Efficiency (FE) for C_2H_4 generation of 32.6%, and a FE for liquid based carbon product of 42.6%. Significant improvements in the partial current densities for COR were observed in relative to planar electrodes or flooded gas diffusion electrodes (GDEs). In addition, a custom test-bed was constructed to characterize the oxidation states of the Cu catalysts in real time along with product analysis though the backside of the GDLs via *operando* X-ray absorption (XAS) measurements.

Introduction

Gas diffusion electrodes (GDEs) have been recently used for the selective reduction of CO_2 and CO to ethylene, ethanol, and other carbon products at high current densities and Faradaic efficiencies.[1–7] GDE configurations, consisting of a three phase interface of reactant gas (CO or CO_2), liquid or polymer electrolyte, and solid electrocatalyst have contributed to the improved performances due to increased concentration of reactant gas at the electrocatalyst surface compared to the limited solubility in liquid electrolytes.[8] Vapor-fed GDEs also provide opportunities for the catalyst materials to operate under a wide range of pH conditions, and favorably steer selectivity through suppression of hydrogen evolution reaction (HER) by controlling the concentration of CO_2 and water vapor.[8]

CO is an important intermediate in CO_2R and is widely accepted as the first intermediate to multi-carbon products.[1,9–11] COR not only has shed light on mechanistic pathways of CO_2R, but also has the advantage of producing an overall more efficient CO_2R system in a two-step, cascade CO_2 reactor, in which the first catalytic reactor converts CO_2 into CO, and the second catalytic reactor converts CO into higher order reduction products such as ethanol or ethylene.[12] For instance, in some electrocatalyst performance regions with high cathodic overpotentials, a relative improvement in solar-to-fuel (STF) conversion efficiency as high as 54% can be obtained by using a two-step reactor.[12] Herein, a hybrid catalyst-bonded membrane device that contains gas only reactants in the cathode chamber and a liquid electrolyte in the anode chamber, separated by an anion exchange membrane (AEM) was constructed, and the impacts of the applied potential and the relative humidity in this hybrid device configuration were studied. Unlike many previously reported flow cells,[3,5,6,13,14] the electrodeposited Cu catalyst in the vapor-fed device was in direct physical contact with the AEM via mechanical pressing. In the hybrid device configuration, the

AEM not only directly facilitates the ionic transport between the cathode and anode chamber, but also controls and dictates the local reaction environments at the reaction sites of the Cu-catalyst for COR. In a slightly modified cell configuration (See Supporting Information Figure S1), *operando* X-ray absorption spectroscopy (XAS) measurements were carried out at device relevant operating current densities to understand the impact of Cu oxidation state on the selectivity of reduced CO products.

Results and Discussion

Figure 1a shows a schematic illustration of the hybrid catalyst-bonded membrane device. The hybrid catalyst-bonded membrane device consists of a cathode chamber and an anode chamber separated by an anion exchange membrane (AEM). In the anode chamber a flowing 1.0 M KOH aqueous electrolyte was used for oxygen evolution reaction (OER), while in the cathode chamber CO gas at a controlled relative humidity (RH) was introduced at a flow rate of 10.0 sccm. The relative humidity of CO was controlled using two mass flow controllers, which controlled one dry stream, and one stream flowing through a bubbler filled with deionized water, as shown in the Supporting Information (Figure S2). A traditional three-electrode configuration was used in the hybrid device with the Cu-GDE as the working electrode, Pt mesh as the counter electrode, and Ag/AgCl as the reference electrode. The Pt mesh was mechanically pressed against the AEM, and Cu/AEM/Pt formed the membrane-electrode-assembly (MEA) for COR in the hybrid device. As the cathode chamber does not contain any liquid electrolyte, the Ag/AgCl reference electrode was placed in the anolyte reservoir (Fig. 1a). The non-traditional three-electrode configuration, in which the reference electrode was placed at the opposite of the counter electrode, was validated experimentally to provide accurate potential points for the working electrode (Figure S3). Total cell resistances measured before and after bulk electrolysis ranged from 2-5 Ω as measured by electrochemical impedance spectroscopy (EIS) (Figure S4).

The Cu electrocatalyst layer was electrodeposited on graphite-based gas diffusion layers (GDL) with a microporous carbon layer in an aqueous solution containing 0.15 M $CuCl_2$, 1.0 M HCl, and 20% ethanol. Electrodepositions were performed potentiostatically by applying -0.50 V vs. Ag/AgCl until 4.5 C cm^{-2} of total charge was passed. During the deposition, the GDL served as the working electrode, copper mesh served as the counter electrode, and Ag/AgCl (sat. KCl) served

as the reference electrode. After electrodeposition, the Cu-GDE was dipped in deionized water several times to rinse off any residual deposition electrolyte, followed by drying in air. Figure 1b shows an optical image of the electrodeposited Cu catalyst on GDL. The morphology and chemical composition of the Cu catalyst was characterized by scanning electron microscopy (SEM), energy dispersive X-ray spectroscopy (EDX), and X-ray diffraction (XRD) before and after deposition. Cross sectional SEM/EDX images (Figure 1c) of the Cu-GDE show a dendrite type morphology and the Cu catalyst mainly deposited on the surface of the GDL with no penetration into the microporous layer of the substrate as shown by SEM/EDX (Figure S5). XRD patterns (Figure 1d) after deposition show peaks matching calculated polycrystalline Cu diffraction patterns, as well as a small amount of Cu_2O present in the bulk.

Prior to bulk electrolysis experiments the Cu-GDE was allowed to equilibrate under a flow of CO (10 sccm) at room temperature and controlled RH for 1 hour at open circuit voltage (OCV) conditions. During this period of time EIS measurements were taken to determine the cell resistance, and it was found over the course of an hour that the OCV typically changed from -0.2 V to -1.0 V, while the cell resistance reduced from ~30 Ω to ~4 Ω at equilibration due to wetting of the anion exchange membrane (Figure S6). Bulk electrolysis experiments were performed using a range of potentials from -1.6 V to -2.2 V vs. Ag/AgCl. Gas products from COR were measured every 6 min during bulk electrolysis by gas chromatograph (GC) (see SI for more details), while liquid products were collected after bulk electrolysis from the anode electrolyte and measured by high performance liquid chromatograph (HPLC). We note that the liquid products identified in the HPLC may not be the direct reduction products of the Cu catalyst in the hybrid device, since the Pt anode could oxidize some liquid products that transported across the AEM. Additional loss of liquid products may have resulted from escaping the cathode or anode chamber in the vapor form,

or being absorbed within the AEM. Nevertheless, from almost all measurements, >85% of electrons participating the reactions were accounted for COR products using GC and HPLC. Figure 2a shows the total operating current densities as a function of time at 100% RH. The total current densities increased for the first 5-10 minutes of operation and reached a relatively steady value after ~20 minutes of operation. Figure 2b shows the corresponding product distribution as a function of applied potentials for COR at 100% RH. At -2.0 V vs. Ag/AgCl, a total carbon product selectivity of 51.1% (23.5% toward generation of C_2H_4, and 27.6% toward generation of liquid based carbon products) was observed. The geometric partial current densities for C_2H_4 generation remained relatively stable between 25 mA cm^{-2} to 30 mA cm^{-2} during the course of testing for COR (Fig. 2c). In comparison, aqueous-based electrolytes exhibit geometric partial current densities < 1 mA cm^{-2} for COR due to the low solubility and low diffusion coefficient of CO, even with a range of catalyst morphologies and nano-structures.[15–18] Figure 2c also shows the geometric partial current density for a flooded GDE, in which the Cu catalyst layer was intentionally flooded with KOH electrolyte ahead of the bulk electrolysis measurements. The partial current density for C_2H_4 generation reduced to ~1-2 mA cm^{-2} , which was comparable to many reported aqueous-based systems.

The impacts of relatively humidity on the activity and selectivity of the Cu-GDE were carried out by experiment and by Multi-physics simulations. Figure 3a shows the linear sweep voltammetry (LSV) of Cu-GDEs at five different RHs from OCV to -2.2 V vs. Ag/AgCl reference electrode at a scan rate of 50 mV s^{-1}. The total operating current density exhibited an increasing trend as the RH increased from 5% to 100%. Current densities up to 350 mA/cm^{-2} were observed at -3.0 V vs. Ag/AgCl (Fig. S7), however at these high current density regions, the intense OER at the anode chamber, which was not optimized for this study, produced large amounts of O_2 bubbles and

caused noisy electrochemical signals. Figure 3b, 3c and 3d show the product distributions of COR at various applied potentials for three different RH conditions. The applied potential had the largest impact on product distribution and in general, high selectivity for COR products were observed at more negative potentials, with the highest selectivity of total COR products of ~75% at -2.0 V and RH 30%. Product distributions were impacted in the RH range of 5% to 30%, with little change in the distribution at RH values >50% (Figure S8).

The relatively small change in the activity and selectivity for COR at various RHs at the cathode feed was investigated by a Multi-physics model that accounted for electrochemical reactions and water transport in the vapor-fed test-bed. Figure 4a shows a schematic illustration of the vapor-fed test-bed in Multi-physics modeling, with the calculation domain of the 2D model for water-vapor transport available in the Supporting Information file. In the AEM domain, three mechanisms of the water transport were considered in the model, including diffusion, water electro-osmotic drag, and hydraulic permeation (see equations S1 to S3). The liquid-gas flow in GDL, MPL, and channel domains were described by two separate Darcy's equations for liquid phase and gas phase, respectively. The water saturation was predicted using Van Genuchten model which was expressed as capillary pressure. The Hertz–Knudsen–Langmuir equation was used to predict the interfacial mass transfer (evaporation and condensation) between liquid and gas phase.[19] The 2D numerical model was used to simulate local RH and water saturation in the catalyst layer as a function of the relative humidity at the cathode inlet. Figure 3b shows the average RH inside the cathode catalyst layer ranged from 88.5% to 99.3% while the external RH was controlled from 5% to 100%. The water saturation in the cathode catalyst layer was close to 0.52 regardless of controlled inlet RH. This is mostly related to the domination of water transport through the AEM by diffusion due to the aqueous-based electrolyte in the anode compartment.

The relatively un-changed local RH and water saturation in the catalyst layer resulted in limited impact to activity and selectivity for COR by changing the inlet RH in the system. To further test this hypothesis, an all vapor fed two-electrode configuration was employed by flowing 100% RH N_2 through the anode at 10.0 sccm, and 5% RH CO in the cathode. The current density was set to 10 mA/cm^2, the average current density when applying -1.6 V in a three-electrode configuration. While the Faradaic efficiency of HER was diminished to 7.5% in this all vapor configuration, the anion exchange membrane would dry out over a short period of time resulting in large cell resistances and increased cell voltages (Figure S9). Further improvements to water transport and controlling RH at the anode could lead to decreased HER and better selectivity of COR products.

A slightly modified GDE cell with an opening for the incident X-ray was constructed to accommodate *operando* XAS measurements. Figure 5a shows the experimental setup for *operando* XAS measurements at the Stanford Synchrotron Radiation Lightsource (SSRL). Multiple Cu oxidation states were observed by XANES measurements after deposition and during OCV conditions, and as the Cu-GDE was allowed to equilibrate an increase in the Cu_2O pre-edge was observed by XANES (Figure S10). There are several factors to consider during this equilibration period, including membrane wetting and oxidation state changes of the electrocatalyst which all play factors in the observed changes. The change in OCV correlates to the oxidation state change while decreased cell resistance is due to membrane wetting and increased conductivity of the membrane. Figure 5b and 5c show the XANES measurements of the Cu-GDEs at OCV and under various applied potentials for COR. Under OCV conditions, a mixed oxidation state of Cu was observed (Figure 2b) with peaks matching both Cu and Cu_2O reference XANES spectra. However, under applied potential for 1 hour, the Cu oxide was eventually reduced and peaks from the Cu-GDE matched metallic Cu (Figure 2c). Hence, after ~20 minutes of

operation when the total operating current density stabilized (Fig. 2a), metallic Cu, instead of Cu oxides, was responsible for COR in the vapor-fed GDEs from XANES measurements. The vapor-fed GDEs offered a unique platform, without the complications in the aqueous electrolyte, for investigating chemical and structural properties of the catalyst under device-relative operating conditions. In aqueous based cells, high current densities may result in bubble formation at the electrode surface blocking active sites and disrupting thee operando measurements. However, in the vapor-fed configuration there is no liquid at the electrode surface and only water vapor is needed. This enables spectroscopic probing of the system at relatively high current density with little interference.

Conclusion

In summary**,** we report a hybrid catalyst-bonded membrane device using only gas reactants for the electrochemical reduction of CO to ethylene, and various liquid products with a selectivity of up to 75% for C_{2+} products. The identified liquid products included acetic acid, propanol, and allyl alcohol. At a relative humidity of 30%, the Cu-based catalyst in the hybrid device configuration exhibited a total operating current density of 87 mA cm^{-2} at -2.0 V vs. Ag/AgCl reference electrode, a Faradaic Efficiency (FE) for C_2H_4 generation of 32.6%, and a FE for liquid based carbon product of 42.6% for CO reduction. Significant improvements in the partial current densities for COR were observed in relative to planar electrodes or flooded gas diffusion electrodes (GDEs). The local RH and water saturation at the catalyst layer were dictated by the diffusional water transport through the AEM and exhibited little change over a wide range of the cathode vapor feed from 5% to 100%. As a result, relative humidity at the cathode had little impacts on the product selectivity and activity in the hybrid catalyst-bonded membrane device. In addition, while different oxidation states of Cu were observed by *operando* XAS measurements in a custom test-bed, these were quickly reduced to metallic Cu and had no direct correlation on the selectivity of ethylene or H_2.

Supporting Information: Detailed descriptions of the Cu electrodepositon, cell design, cross sectional SEM/EDX data, relative humidity control schematic, details for numerical modelling for water-vapor transport, cell resistance before and after bulk electrolysis, CV scans of ferrocyanide in a GDE configuration, OCV and EIS data during equilibration, XANES data at equilibration, and SEM images before and after bulk electrolysis, can be found in the Supporting Information file.

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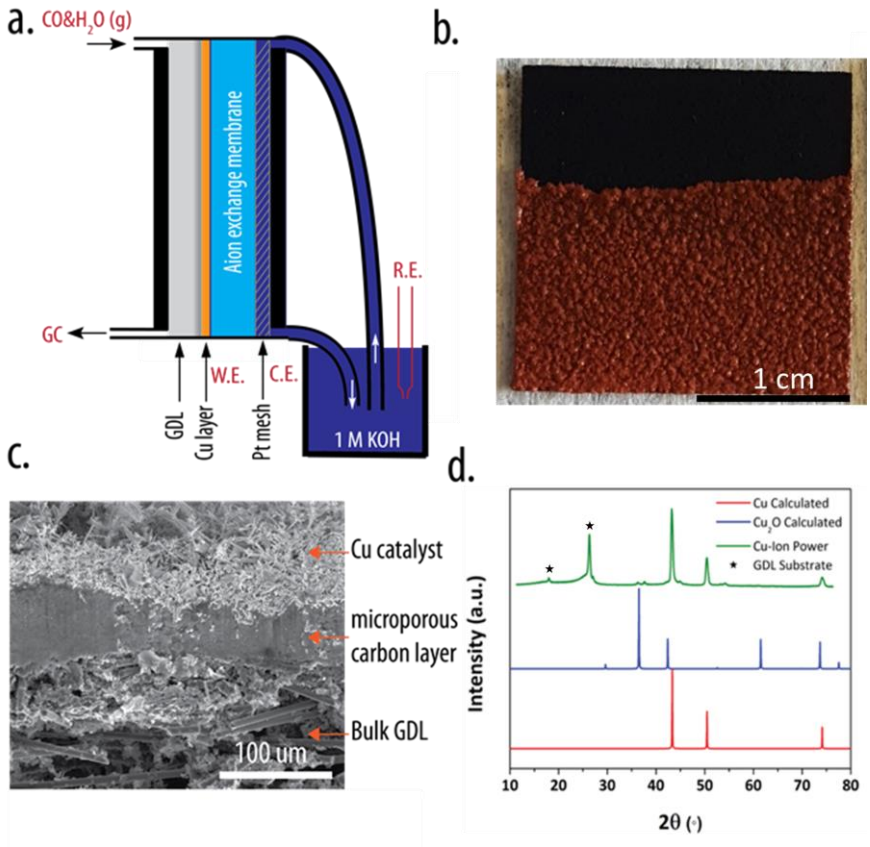

Figure 1. (a) A schematic illustration of the hybrid catalyst-bonded membrane device. (b) Optical image of the electrodeposited Cu catalysts on GDL. (c) Cross sectional SEM images of the Cu catalysts on GDL. (d) XRD pattern of the electrodeposited Cu catalyst on GDL.

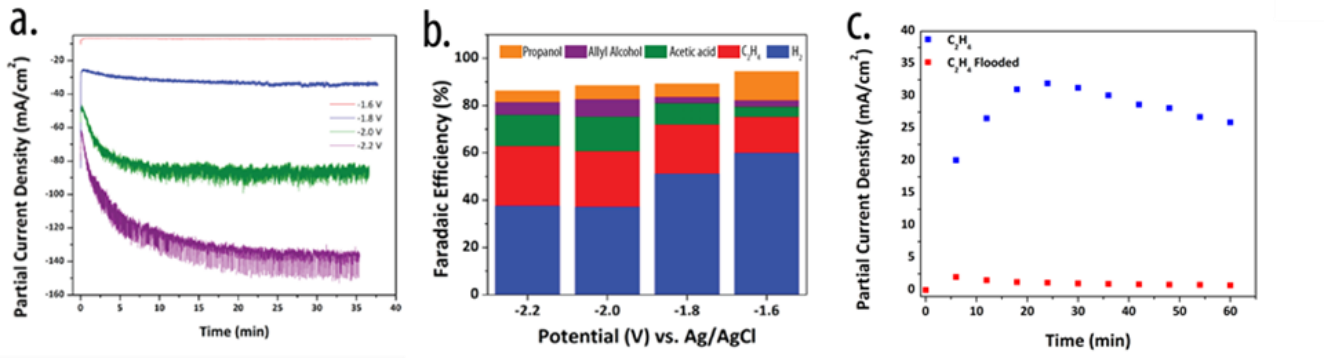

Figure 2. (a) Total operating current densities as a function of time at different applied potentials for COR at 100% RH. (b) Product distribution as a function of applied potentials for COR at 100% RH. (c) Partial current densities for C_2H_4 generation as a function of time for the flooded cell and the vapor cell at 100% RH.

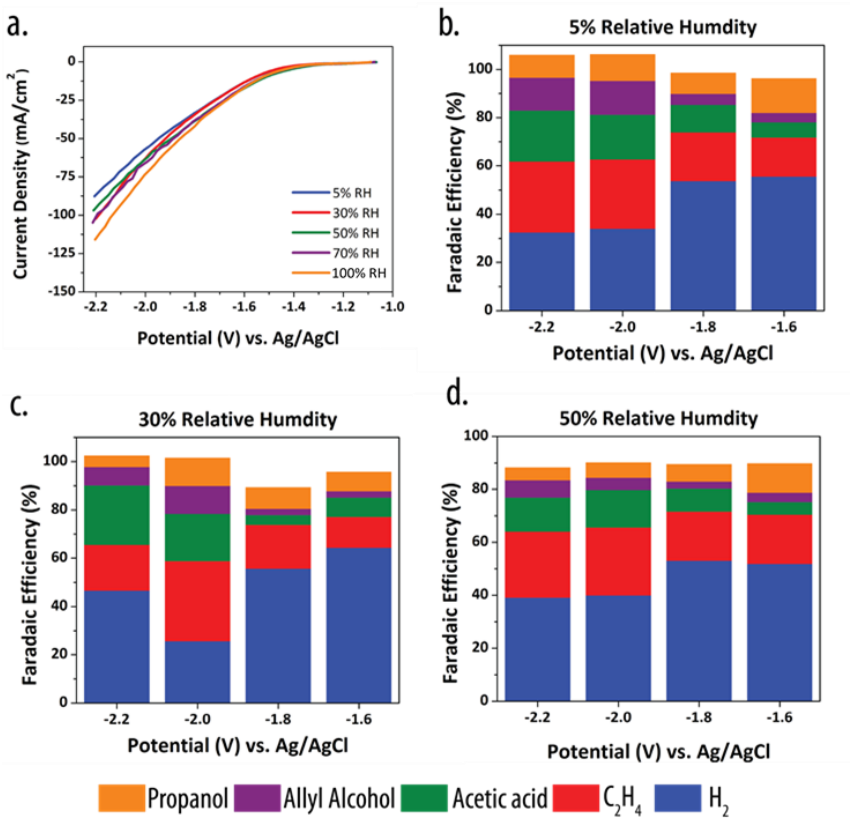

Figure 3. (a) LSV scans of Cu-GDE at various relative humidities. (b)-(d) Product distribution as a function of applied potential and controlled relative humidity for COR.

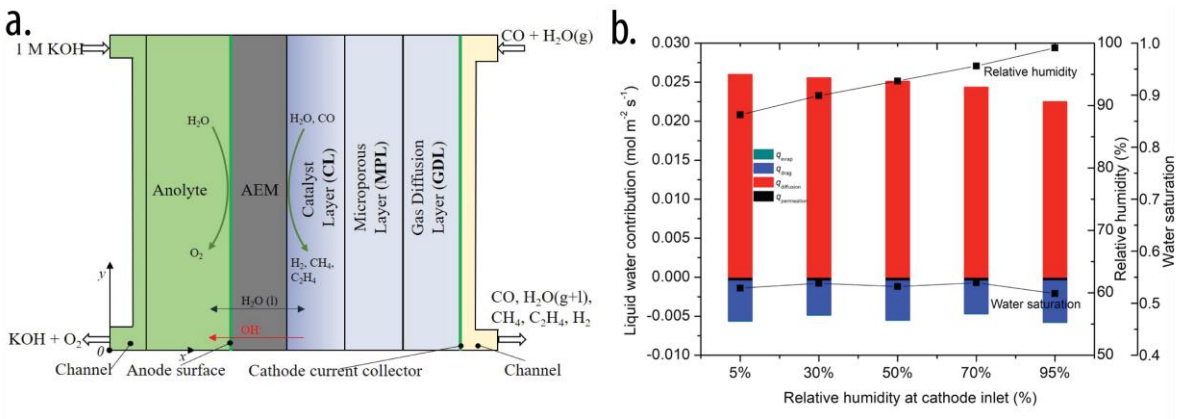

Figure 4. (a) A schematic illustration of the vapor-fed device in Multiphysics modeling (b) Simulated average relative humidity and water saturation in the cathode catalyst layer.

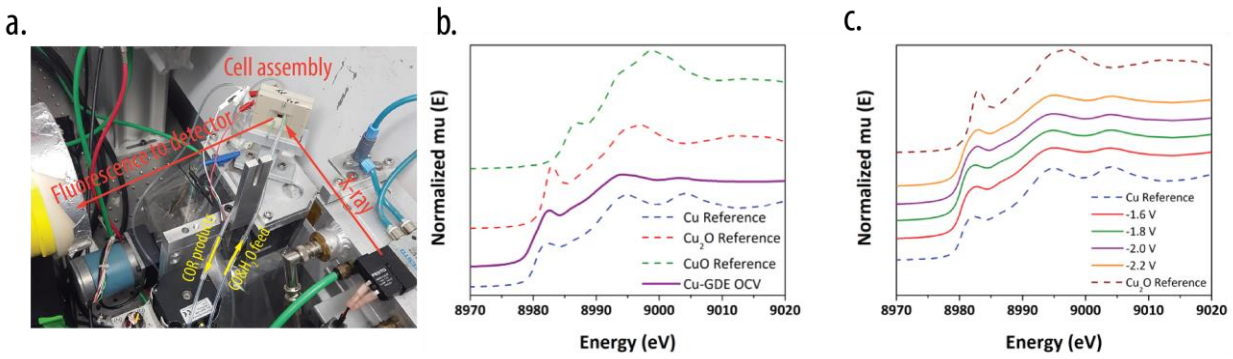

Figure 5. (a) Experimental setup for *Operando* XAS at SSRL (b) *Operando* XANES measurements at OCV and (c) under various applied potentials.

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