

# Lawrence Berkeley National Laboratory

## Lawrence Berkeley National Laboratory

### Title

THE BARRIERS TO RING ROTATION IN 1,1',4,4'-TETRA-t-BUTYLURANOCENE AND 1,1',3,3'-TETRA-t-BUTYLFERROCENE

### Permalink

<https://escholarship.org/uc/item/4vk362sx>

### Author

Luke, Wayne D.

### Publication Date

1980-10-01



# Lawrence Berkeley Laboratory

UNIVERSITY OF CALIFORNIA

## Materials & Molecular Research Division

Submitted to the Journal of the American Chemical Society

THE BARRIERS TO RING ROTATION IN 1,1',4,4'-TETRA-t-BUTYLURANOCENE AND 1,1',3,3'-TETRA-t-BUTYLFERROCENE

Wayne D. Luke and Andrew Streitwieser, Jr.

October 1980

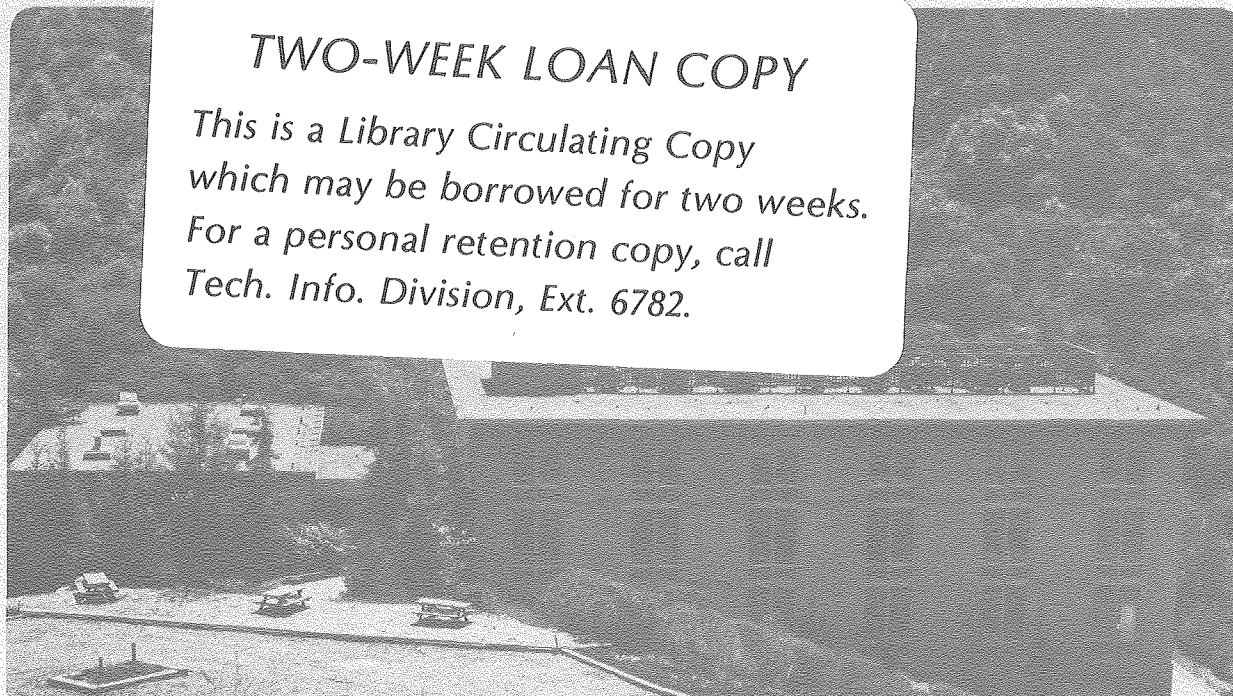
RECEIVED  
LAWRENCE  
BERKELEY LABORATORY

DEC 11 1980

LIBRARY AND  
DOCUMENTS SECTION

### TWO-WEEK LOAN COPY

*This is a Library Circulating Copy  
which may be borrowed for two weeks.  
For a personal retention copy, call  
Tech. Info. Division, Ext. 6782.*



LBL-11722 c, 2

## DISCLAIMER

This document was prepared as an account of work sponsored by the United States Government. While this document is believed to contain correct information, neither the United States Government nor any agency thereof, nor the Regents of the University of California, nor any of their employees, makes any warranty, express or implied, or assumes any legal responsibility for the accuracy, completeness, or usefulness of any information, apparatus, product, or process disclosed, or represents that its use would not infringe privately owned rights. Reference herein to any specific commercial product, process, or service by its trade name, trademark, manufacturer, or otherwise, does not necessarily constitute or imply its endorsement, recommendation, or favoring by the United States Government or any agency thereof, or the Regents of the University of California. The views and opinions of authors expressed herein do not necessarily state or reflect those of the United States Government or any agency thereof or the Regents of the University of California.

The Barriers to Ring Rotation in 1,1',4,4'-tetra-t-Butyluranocene  
and 1,1',3,3'-tetra-t-Butylferrocene.

Wayne D. Luke<sup>1</sup> and Andrew Streitwieser, Jr.\*

Department of Chemistry, University of California,  
Berkeley, California 94720

and

Materials and Molecular Research Division,  
Lawrence Berkeley Laboratory, Berkeley, California 94720

Abstract

The <sup>1</sup>H-NMR spectra of 1,1',4,4'-tetra-t-butyluranocene show a singlet for the t-butyl protons and three singlets for the ring protons. Coalescence occurs at -65 to -85°C with splitting of each peak at lower temperatures, corresponding to  $\Delta G^\ddagger$  of rotation of 8.3 Kcal mol<sup>-1</sup>. Coalescence was also observed for 1,1',3,3'-tetra-t-butylferrocene and a corresponding  $\Delta G^\ddagger = 13.1$  Kcal mol<sup>-1</sup> was obtained.

The energy barrier to ring rotation in sandwich organometallic complexes is generally quite small. The rotational barrier in ferrocene has been measured by several different methods; gas phase electron diffraction:  $0.9 \pm 0.3$  Kcal mol<sup>-1</sup>,<sup>2-5</sup> IR:  $0.9$  Kcal mol<sup>-1</sup>,<sup>6</sup> wide-line <sup>1</sup>H NMR second moment measurements:  $1.8 \pm 0.2$  Kcal mol<sup>-1</sup>.<sup>7-11</sup> Similarly, second moment studies have shown the corresponding barrier in nickelocene, cobaltocene and chromocene to be  $1.8 \pm 0.2$  Kcal mol<sup>-1</sup>,<sup>7</sup> in ruthenocene,  $2.3 \pm 0.2$  Kcal mol<sup>-1</sup>,<sup>10,11</sup> in bisbenzene-chromium,  $4.56$  Kcal mol<sup>-1</sup>,<sup>11</sup> and  $0$  Kcal mol<sup>-1</sup> in uranocene.<sup>8</sup> Appropriate substitution of the rings would be expected to raise this barrier; however, data on barriers to internal rotation in substituted metallocenes are limited. Second moment studies on several crystalline mono- and 1,1'-disubstituted ferrocenes have indicated internal rotation barriers of  $0.9-5.0$  Kcal mol<sup>-1</sup>.<sup>12-16</sup> A gas phase electron diffraction study has recently shown the barrier to be  $1.0 \pm 0.3$  Kcal mol<sup>-1</sup> for decamethylferrocene.<sup>17</sup> An equally low barrier has been assumed for 1,1',3,3',5,5',7,7'-octamethyluranocene from X-ray structural data.<sup>18</sup> We now report measurement of the barrier to ring rotation in the unsymmetrically tetra-t-butyl substituted metallocenes 1,1',4,4'-tetra-t-butyluranocene, 1, and 1,1',3,3'-tetra-t-butylferrocene, 2, by variable temperature <sup>1</sup>H NMR.

Variable temperature <sup>1</sup>H NMR spectra of 1,1',4,4'-tetra-t-butyluranocene<sup>19</sup> show that we were successful in "freezing out" ring rotation on the NMR time scale (Fig. 1). Above  $-30^{\circ}\text{C}$ , where ring rotation is rapid, the spectrum consists of four singlets with an area ratio of 18:2:2:2, corresponding to the t-butyl protons and three

sets of non-equivalent ring protons. In the low temperature spectrum below  $-80^{\circ}\text{C}$ , the compound is conformationally rigid on the NMR time scale, and the spectrum consists of eight singlets with area ratios of 9:9:1:1:1:1:1:1, corresponding to two kinds of t-butyl groups and six distinct sets of ring protons. Labeling the three ring proton resonances in the high temperature spectrum A-C, from low field to high field, the B resonance coalesces first, followed by virtually simultaneous coalescence of the A and C resonances and eventual coalescence of the t-butyl group as the temperature is lowered.

The coalescence temperature is dependent on the difference in frequency of the resonance signals in the low temperature spectrum of the conformationally 'frozen' molecule. The standard equation at coalescence, eq. 1, implies that the larger the difference in frequency  $\nu$ , the higher the coalescence temperature.

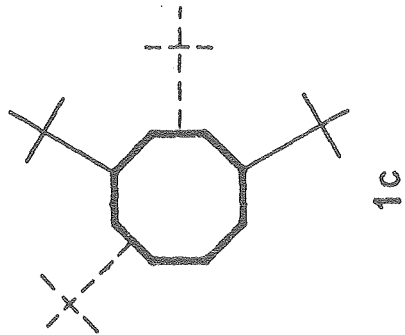
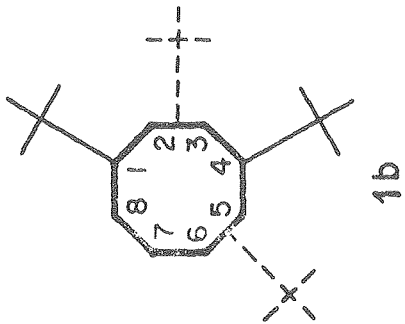
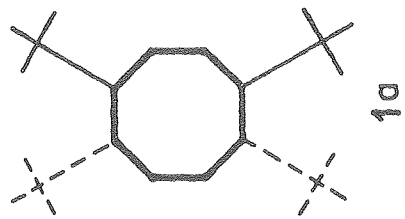
$$\text{At coalescence: } k_c = \frac{\pi(\Delta\nu)}{\sqrt{2}} \quad (1)$$

Because of the paramagnetic uranium center, all of the observed resonances are highly temperature dependent but the isotropic shifts are linear in  $T^{-1}$ .<sup>20</sup> These linear  $T^{-1}$  dependences for the ring protons both before and after coalescence are shown in Figs. 2 and 3. All sets could be readily extrapolated to the coalescence temperatures to obtain each  $\Delta\nu$  (Table 1). The assignment of the primed resonances below coalescence in Fig. 3 was made on the basis that the average of each assigned pair best fits the extrapolated value of the

Figs 2,3

corresponding unprimed resonance. Thus, the uranocene system is a unique example in which the uranium center serves effectively as an internal paramagnetic shift reagent to clearly resolve all of the nonequivalent proton resonances both in the limit of rapid and slow ring rotation. Accordingly,  $\Delta G^\ddagger$  for internal rotation can be calculated from four distinct sets of coalescence data (Table 2) to give an average value of 8.3 Kcal mol<sup>-1</sup> at -70°C. If the corresponding  $\Delta S^\ddagger$  derives principally from hindering the rotations of the two t-butyl groups passing each other in the transition state, we can estimate a contribution of 2-3 e.u. per group or a total  $\Delta S^\ddagger$  of about -5 e.u.<sup>21</sup> Correspondingly,  $\Delta H^\ddagger$  is about 7 Kcal mol<sup>-1</sup>.

We cannot definitively assign the specific ring proton resonances; however, observation of six ring proton resonances at low temperature eliminates la as the low temperature conformation of 1,1',4,4'-tetra-t-butyluranocene since it contains a plane of symmetry. Inspection of molecular models suggests that lb or a similar structure is the most reasonable conformation of the two rings and undoubtedly corresponds to the low temperature conformation. Moreover, this structure does suggest some partial assignments of the frequencies. Ring protons H-2 and H-3 differ little in the equilibrating structures lb and lc and may be assigned to A' and A" which show the smallest  $\Delta\delta$  and are also furthest downfield; the presence of a t-butyl group directly beneath in the ring below appears to produce a downfield shift.<sup>20</sup> On this basis B" and C" correspond to H-7 or H-8 in lb but no further detail seems appropriate at this time.



XBL 809-11920



Our success in measuring the rotational barrier in 1,1',4,4'-tetra-*t*-butyluranocene suggested that the rotational barrier in the ferrocene analog, 1,1',3,3'-tetra-*t*-butylferrocene, could be measured in a similar manner. One would expect the barrier in the ferrocene analog to be larger since the ring-ring distance in ferrocene is 0.52Å smaller than in uranocene. Surprisingly, although 1,1',3,3'-tetra-*t*-butylferrocene has been prepared several times,<sup>22-24</sup> and its structure has been elucidated by single crystal X-ray diffraction,<sup>25</sup> the ring rotational barrier has not been reported. In fact, we could not find any report in the literature concerning the determination of ring rotational energy barriers of substituted ferrocenes in solution by variable temperature <sup>1</sup>H NMR.

At 30°C, the <sup>1</sup>H NMR spectrum of 1,1',3,3'-tetra-*t*-butylferrocene consists of one *t*-butyl resonance (1.218 ppm, s, 18H) and two ring proton resonances split by 1,3 coupling; (3.822 ppm, t, 2H, J=2.45 Hz, 3.097 ppm, d, 4H, J=1.46 Hz), indicating rapid ring rotation relative to the NMR time scale. Below -35°C two distinct *t*-butyl resonances are observed at 1.231 ppm and 1.206 ppm (Fig. 4). Unfortunately, the resonances of the non-equivalent ring proton in the conformationally 'frozen' molecule could not be resolved. At -60°C the high field ring proton resonance is a broad singlet while the low field resonance is a broad doublet and thus, only the *t*-butyl resonance can be used to calculate  $\Delta G^\ddagger$ .

Taking its coalescence temperature at -27.5°C, the calculated  $\Delta G^\ddagger$  for ring rotation is 13.1 Kcal mol<sup>-1</sup>. As expected, both the coalescence temperature is higher and  $\Delta G^\ddagger$  is larger in the ferrocene than in the corresponding uranocene.

Acknowledgement. This work was supported in part by the Division of Nuclear Sciences, Office of Basic Energy Services, U.S. Department of Energy, under contract no. W-7405-Eng-48, and by NSF grant no. CHE78-24084.

### Experimental

$^1\text{H}$  NMR spectra were run on the Berkeley 180 MHz FT spectrometer equipped with a variable temperature probe. All uranocene spectra were run in toluene- $\text{d}_8$ . The temperature of the probe was monitored by a pre-calibrated thermocouple 5 mm from the sample tube, and could be held to  $\pm 0.3^\circ\text{C}$  over the dynamic temperature range. The data for 1,1',4,4'-tetra-t-butyluranocene are shown in Fig. 1 and 2. The least squares regression lines of Fig. 2 are summarized in Table 2.

1,1',4,4'-tetra-t-Butyluranocene was prepared by M. J. Miller.<sup>19</sup> 1,1',3,3'-tetra-t-Butylferrocene was prepared by the method of Leigh<sup>22</sup> by treatment of ferrocene with 4 eq. of t-butyl chloride and 4 eq. of  $\text{BF}_3 \cdot \text{Et}_2\text{O}$  in refluxing hexane for 24 h. The tri-t-butylferrocene was removed by reduced pressure short path distillation followed by two recrystallizations from  $\text{CH}_2\text{Cl}_2$ ; yield 23%, mp  $196\text{--}197^\circ\text{C}$  (lit.<sup>22</sup>  $198^\circ\text{C}$ ).

## References

- 1) Current address, Eli Lilly and Co., Tippecanoe Laboratories,  
Box 685, West Lafayette, Indiana, 47902.
- 2) Haaland, A.; Nilsson, J. E. Chem. Comm. 1968, 88.  
~~~~~
- 3) Haaland, A.; Nilsson, J. E. Acta. Chem. Scand. 1968, 22, 2653.  
~~~~~
- 4) Bohn, R.; Haaland, A. J. Organomet. Chem. 1966, 5, 470.  
~~~~~
- 5) Haaland, A. Fortschritte der Forschung, 1975, 53, 1.  
~~~~~
- 6) Rocquet, F.; Berreby, L.; Marsault, J. P. Spectrochim.  
Acta Part A 1973, 29, 1101.  
~~~~~
- 7) Makova, M. K.; Karimov, Y. S.; Kochetkova, N. S.; Leonova, E. V.  
Teor. Eksp. Khim. 1972, 8, 259.  
~~~~~
- 8) Anderson, S. E. J. Organomet. Chem. 1974, 71, 263.  
~~~~~
- 9) Clark, A. H. in "Internal Rotation in Molecules" W. J. Orville-  
Thomas, ed., J. Wiley and Sons, New York, 1974, p. 325.
- 10) Holm, C. H.; Ibers, J. A. J. Chem. Phys. 1969, 30, 885.  
~~~~~
- 11) Campbell, A. J.; Fyfe, C. A.; Harold-Smith, D.; Jeffrey, K. R.  
Mol. Cryst. Liq. Cryst. 1976, 36, 1.  
~~~~~
- 12) Mulay, L. N.; Attalla, A. J. Am. Chem. Soc. 1963, 85, 702.  
~~~~~
- 13) Mulay, L. N.; Fox, M. E. Naturweiss, 1962, 49, 446.  
~~~~~
- 14) Mulay, L. N.; Attala, A. Naturweiss, 1963, 50, 151.  
~~~~~
- 15) Sorriso, S.; Cardaci, G.; Murgia, S. M. J. Organomet. Chem.  
1972, 44, 181.  
~~~~~
- 16) Sorriso, S. J. Organomet. Chem. 1979, 179, 205.  
~~~~~
- 17) Almenningen, A.; Haaland, A.; Samdal, S.; Brunvoll, J.; Robbins,  
J. L.; Smart, J. C. J. Organomet. Chem. 1979, 173, 293.  
~~~~~

- 18) Hodgson, K. O.; Raymond, K. N. Inorg. Chem. 1973, 12, 458.  
~~~~~
- 19) Miller, M.; Streitwieser, A., Jr. J. Org. Chem., submitted.
- 20) Luke, W. D.; Streitwieser, Jr., A. in "Lanthanide and Actinide Chemistry and Spectroscopy", Edelstein, N. M., ed., ACS Symposium Series 131, Washington, D.C. (1980), p. 93.
- 21) Based on Table 27-11 of G. N. Lewis and M. Randall, "Thermodynamics", rev. by K. S. Pitzer and L. Brewer, 2nd ed., McGraw-Hill Book Co., New York, N. Y. (1961), p. 444. For a large group like t-butyl 1/Q is small and we use the left side of the table. Reasonable values of barrier heights in the middle portion of the table give entropy changes of the order of 2-3 e.u. We thank Professor K. S. Pitzer for a helpful discussion.
- 22) Leigh, T. J. Chem. Soc. 1964, 3294.  
~~~~~
- 23) Nesmeyanov, A. N.; Kochetkova, N. S. Dokl. Akad. Nauk. SSSR 1957, 117, 92.  
~~~~~
- 24) Nesmeyanov, A. N.; Kochetkova, N. S.; Vitt, S. V.; Bondarev, V. B.; Kovshov, E. I. J. Dokl. Akad. Nauk. SSSR 1964, 156, 99.  
~~~~~
- 25) Kaluski, Z. L.; Gusev, A. I.; Kalinin, A. E.; Struchkov, Y. T. Zh. Strukt. Khim. 1972, 13, 950.  
~~~~~

Table 1. Least squares linear regression vs  $T^{-1}$  of the variable temperature  $^1\text{H-NMR}$  data for 1,1',4,4'-tetra-*t*-butylurano-cene.

| <u>Resonance</u>  | <u>Slope x 10<sup>3</sup></u> | <u>Intercept</u> | <u>r<sup>2</sup></u> |
|-------------------|-------------------------------|------------------|----------------------|
| Ring A            | -7.72                         | -5.49            | 0.9985               |
| Ring B            | -14.73                        | 3.26             | .9970                |
| Ring C            | -15.42                        | 2.95             | .9983                |
| <i>t</i> -butyl   | -5.09                         | 5.04             | .9980                |
| Ring A'           | -3.73                         | -13.66           | .9991                |
| Ring A''          | -8.17                         | -3.85            | .9986                |
| Ring B'           | -9.40                         | -4.14            | 1.0000               |
| Ring B''          | -19.31                        | 17.37            | .9985                |
| Ring C'           | -11.87                        | -3.04            | .9999                |
| Ring C''          | -17.41                        | 10.25            | .9989                |
| <i>t</i> -butyl'  | -3.86                         | 0.17             | 1.0000               |
| <i>t</i> -butyl'' | -5.76                         | 9.31             | 1.0000               |

Table 2. Calculation of rotation barrier in 1,1',4,4'-tetra-t-butyluranocene.

| <u>Proton</u> | <u>Coalescence<br/>temp, °C</u> | <u>Isotropic shift<sup>a</sup><br/>ppm</u> |        | <u><math>\Delta G^\ddagger</math><br/>kcal mol<sup>-1</sup></u> |
|---------------|---------------------------------|--|--------|---|
| t-butyl       | -85                             | -20.37                                     | -21.33 | 8.61  |
| ring A        | -70                             | -33.49                                     | -47.31 | 8.25  |
| ring B        | -65                             | -49.36                                     | -75.46 | 8.20  |
| ring C        | -70                             | -61.52                                     | -75.52 | 8.25  |
|               |                                 | Average                                    |        | 8.3 ± 0.2   |

(a) Extrapolated to coalescence temperature from shifts of the conformationally "frozen" molecule at lower temperatures.

Fig. 1.  $^1\text{H}$  NMR spectra of 1,1',4,4'-tetra-*t*-butyluranocene as a function of temperature. The upper group of spectra show only the ring protons. The spectra at low temperatures show the *t*-butyl group as well.

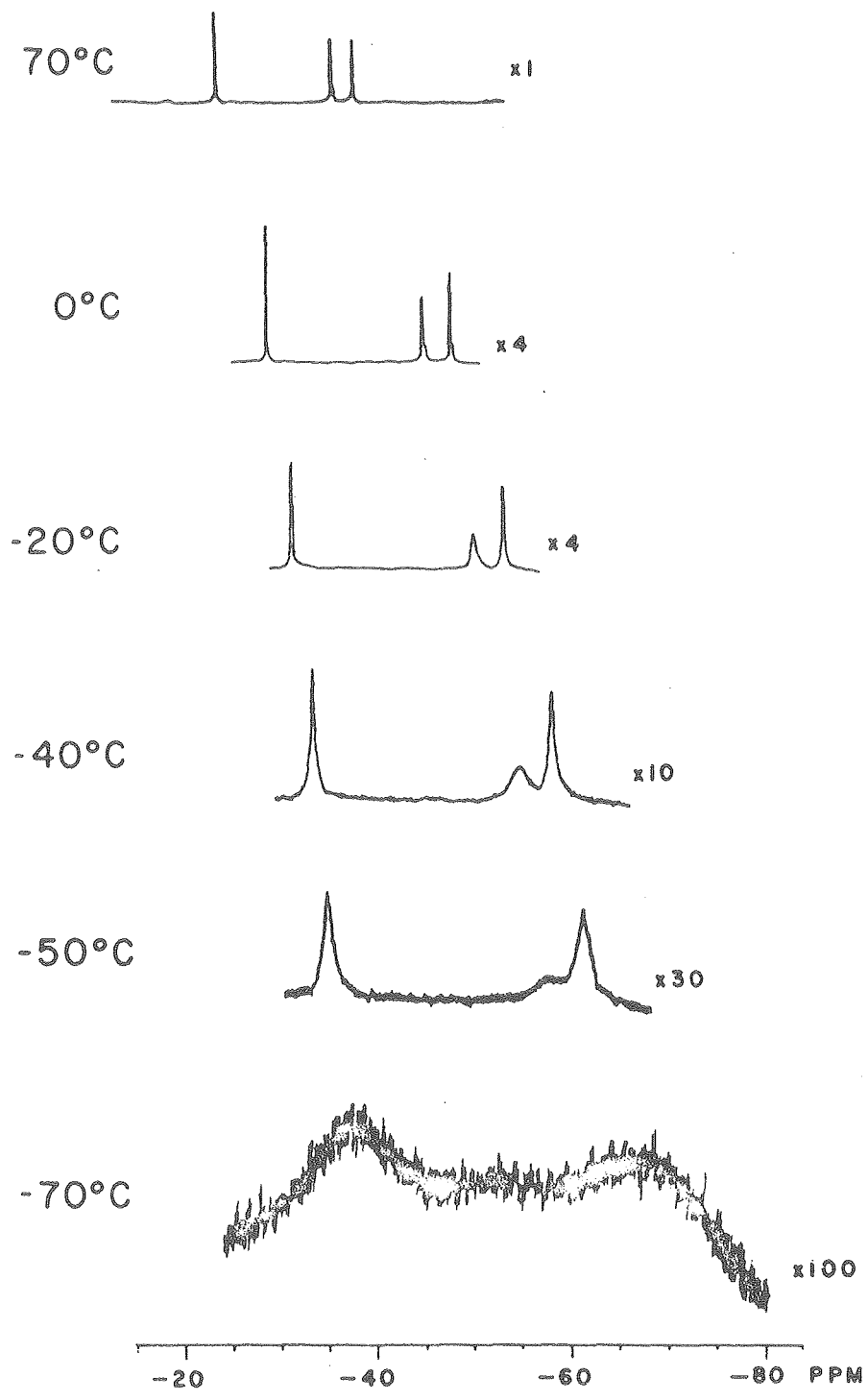
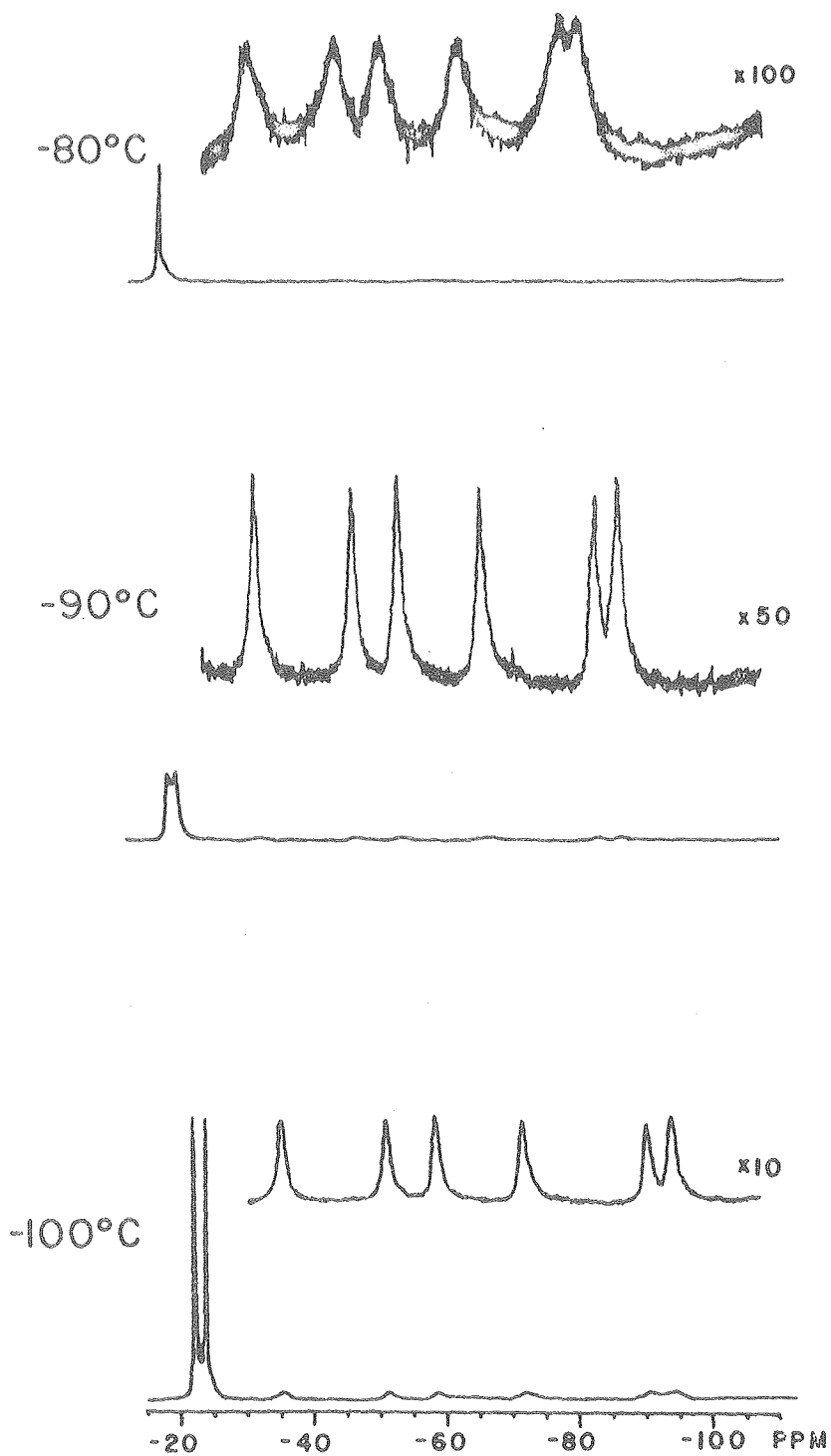


Figure 1 (top)

XBL 809-11922

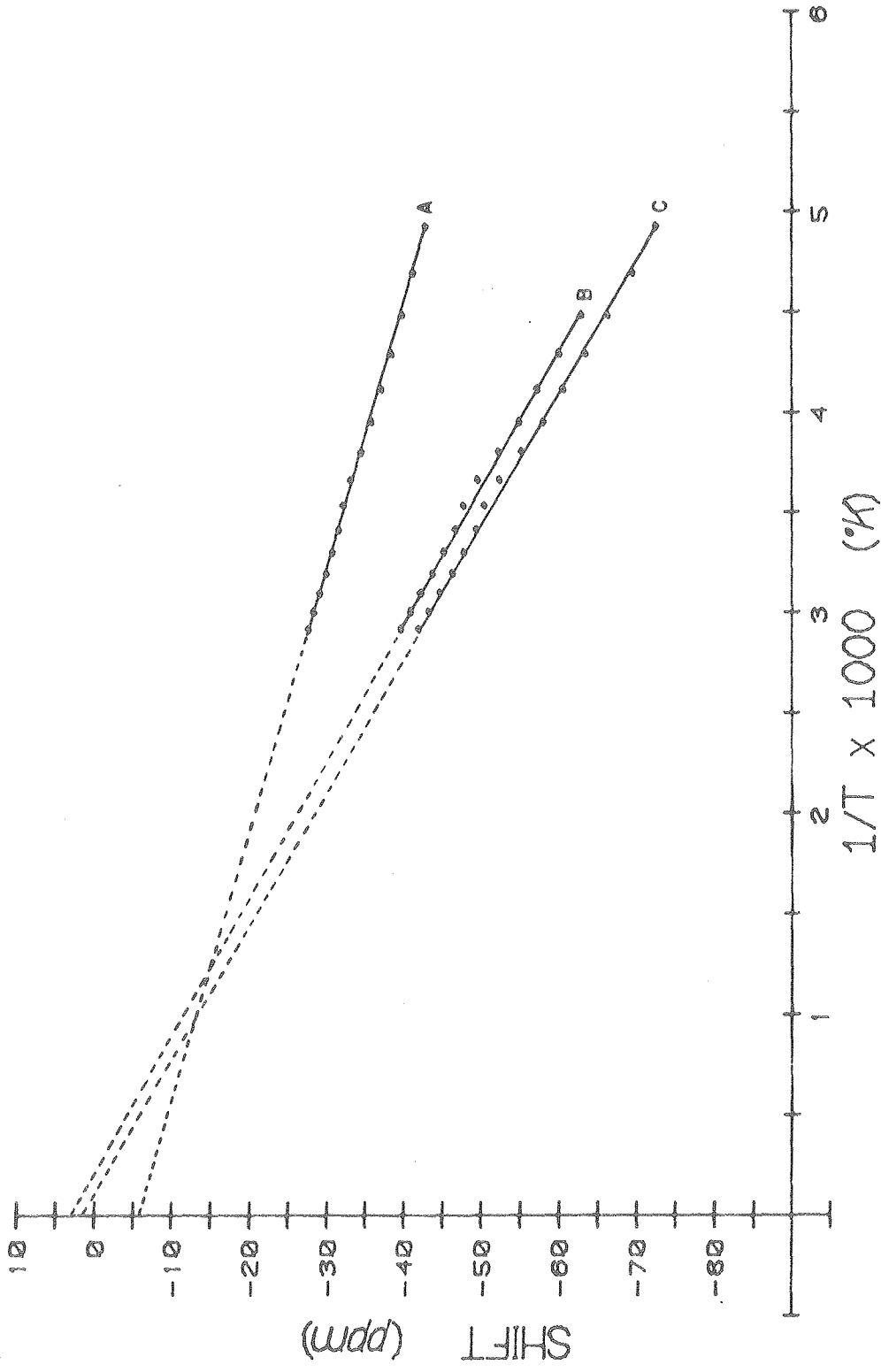




XBL 809-11921

Figure 1 (bottom)

Fig. 2. Temperature dependence of ring proton  $^1\text{H}$  NMR shifts above coalescence showing linearity in  $1/T$  and extrapolation to  $1/T = 0$ .



XBL 809-11923

Figure 2

Fig. 3. Proton shifts before and after coalescence.

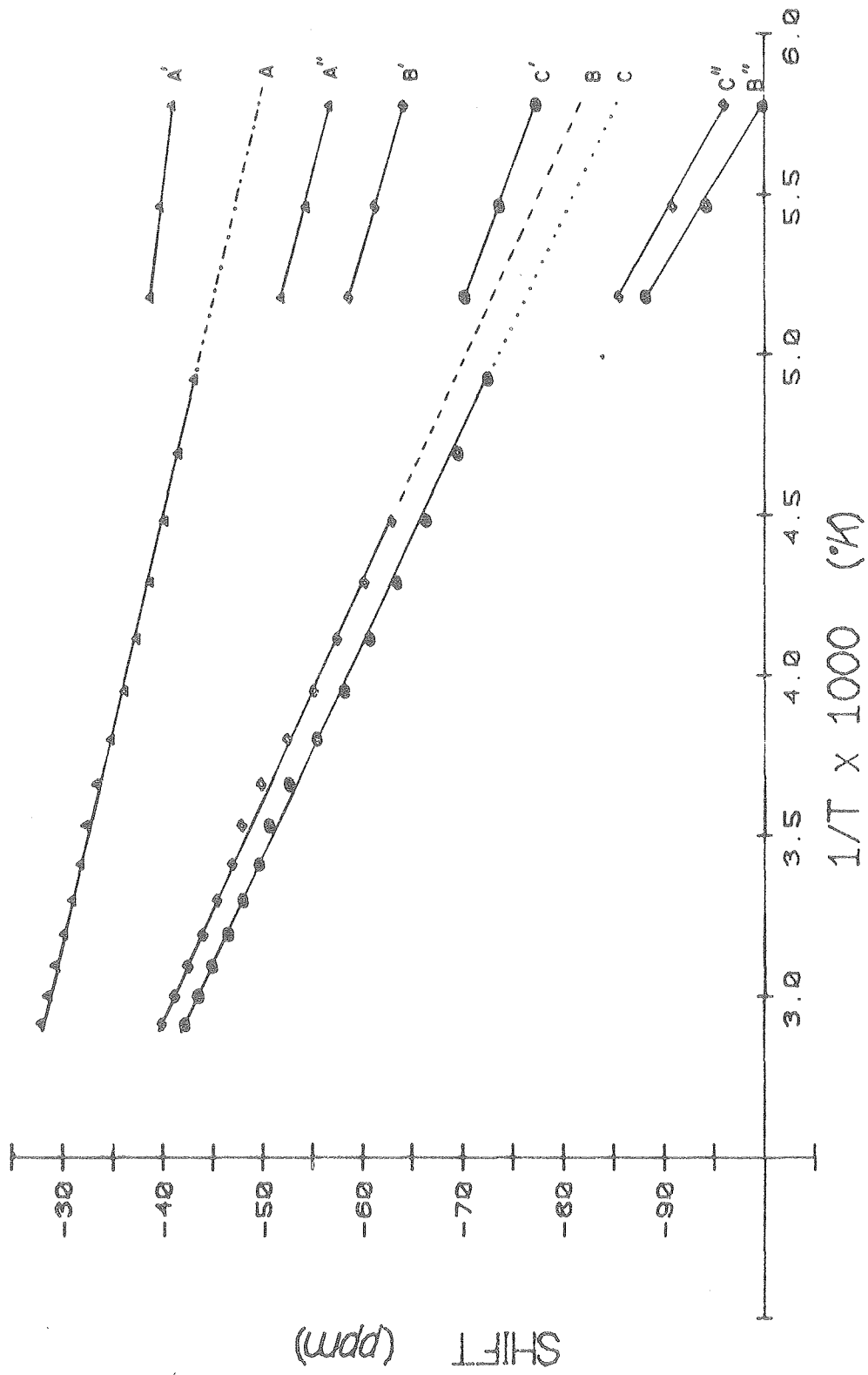
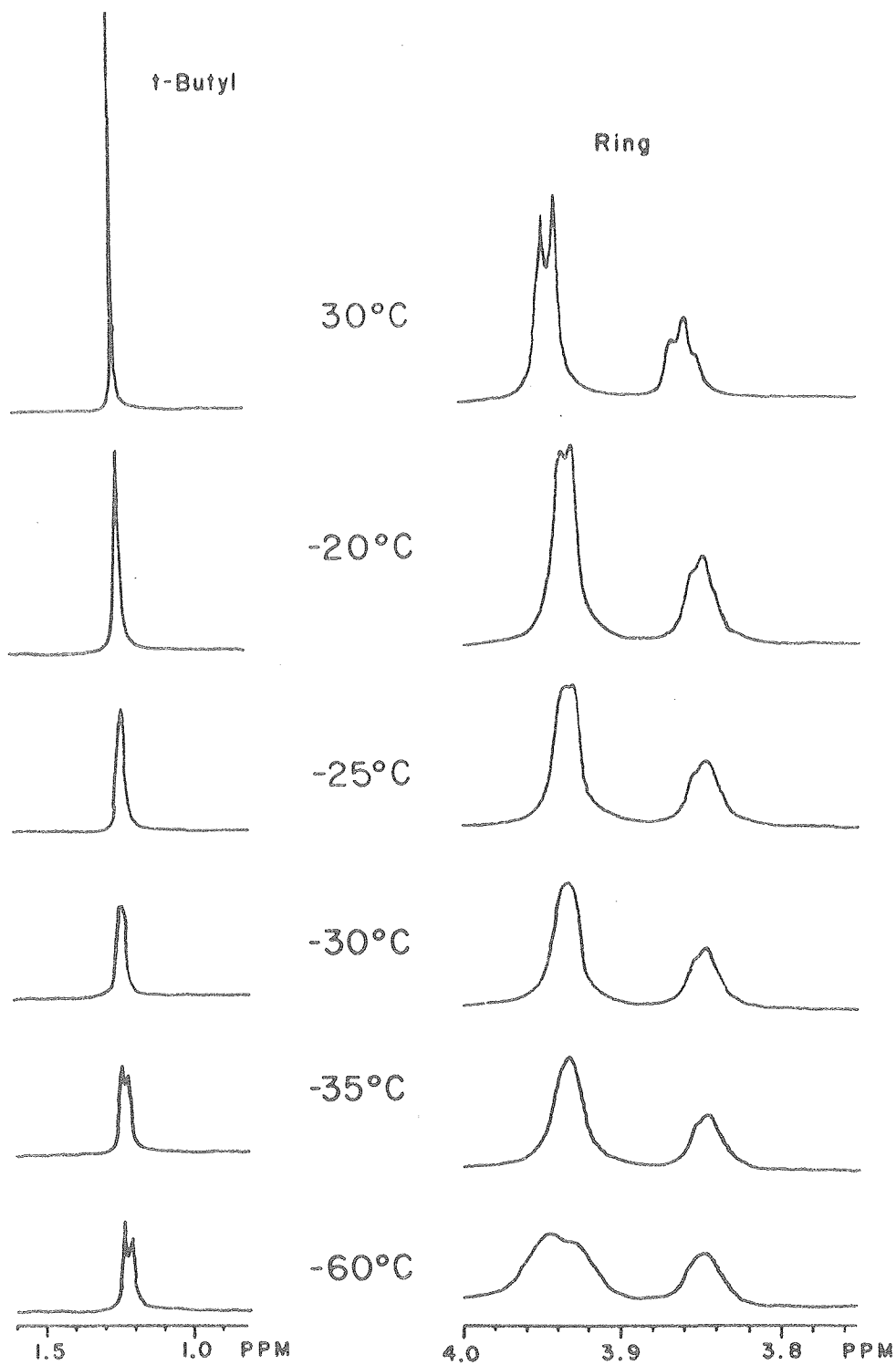


Figure 3

XBL 809-11924

Fig. 4.  $^1\text{H}$  NMR spectra of 1,1',3,3'-tetra-*t*-butyl-ferrocene as a function of temperature.



XBL 809-11925

Figure 4