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Title

Erratum: ^1H -NMR measurements on proton mobility in nano-crystalline YSZ (Physical Chemistry Chemical Physics (2013) DOI: 10.1039/C3CP53039F)

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Peer reviewed

A new mechanism for ozonolysis of unsaturated organics on solids: phosphocholines on NaCl as a model for sea salt particles

Federico Karagulian, A. Scott Lea, Christopher W. Dilbeck and Barbara J. Finlayson-Pitts

Phys. Chem. Chem. Phys., 2008, **10**, 528–541 (DOI: 10.1039/B712715D). **Amendment published 20th November 2012.**

There was an error in converting the data on the time dependence of the secondary ozonide (SOZ) absorbance (Fig. 10) to rates of SOZ formation as a function of O_3 (Fig. 12). This affects the values of the rate constants reported in the paper. However, the overall conclusions regarding the reaction mechanism and the effects of water and irradiation remain valid. Corrections for the kinetics are summarized below.

Kinetics section

Fig. 10 is correct but Fig. 12 should be replaced by the following Figure:

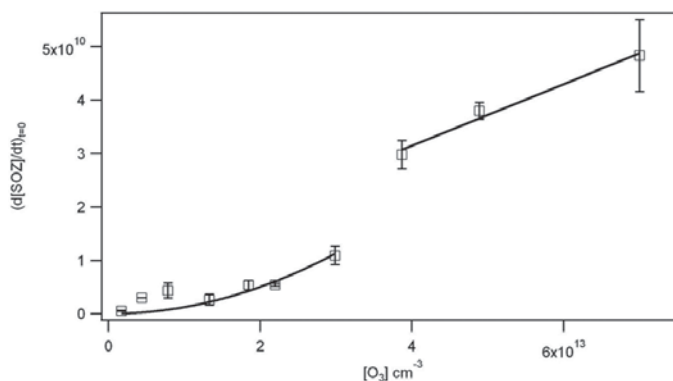


Fig. 12 Initial rates of formation of SOZ (molecule $\text{cm}^{-2} \text{s}^{-1}$) as a function of gas phase ozone concentration. Solid lines are the fits based on a first order dependence on ozone at high concentrations, and a second order dependence at low concentrations as described in the text.

As in the original paper, at high ozone concentrations, the rate of formation of SOZ is linear in ozone and approaches $k_1[O_3][OPPC]$. From the slope of the line at ozone concentrations from $(3.9\text{--}7.0) \times 10^{13}$ molecules cm^{-3} , $k_1[OPPC] = (5.8 \pm 1.6) \times 10^{-4}$ (2s). Using an OPPC concentration of 3.2×10^{13} OPPC cm^{-2} , a value for k_1 of $(1.8 \pm 0.5) \times 10^{-17}$ $\text{cm}^3 \text{molecule}^{-1} \text{s}^{-1}$ is derived. This corresponds to a reaction probability of $\gamma = 6 \times 10^{-8}$. However, because of the non-linearity of SOZ formation over the range of ozone concentrations studied, the effective reaction probability varies. For example, converting the measured rate of SOZ formation at an ozone concentration of 2.3×10^{13} molecule cm^{-3} (0.9 ppm) to an effective reaction probability gives $\gamma = 3 \times 10^{-8}$. This order of magnitude is similar to that expected from the gas phase ozonolysis of alkenes, not higher as originally stated. This is attributed to the OPPC forming a non-fluid layer on the NaCl substrate, rather than a fluid, more porous and dynamic film as for OPPC on an aqueous substrate.

Fitting the data in Fig. 12 at lower O_3 concentrations ($< 3 \times 10^{13}$ molecule cm^{-3}) to the form $k_1 k_4 [O_3]^2 [OPPC] / k_2$ as described in the original paper gives the fit shown in revised Fig. 12, from which a value is obtained for the ratio $k_4 / k_2 = (2.2 \pm 0.9) \times 10^{-14}$ $\text{cm}^3 \text{molecule}^{-1}$ (2s). Independent values of k_4 and k_2 , and hence the lifetime of POZ with respect to thermal decomposition, could not be derived from the data.

With water vapor addition

Fig. 13 shows the corrected dependence of SOZ on the water vapor concentration. The steep drop-off with increasing water indicates that the Criegee intermediate and/or the POZ are efficiently trapped by even small amounts of water. Eqn (VIII) of the original paper indicates that if water dominates the loss of the CI or POZ, the rate of SOZ formation should depend on the inverse water concentration to the first power for either CI or POZ reacting with water, or to the second power for both CI and POZ reacting with water. The dependence in Fig. 13 is best matched by variation of $d[\text{SOZ}]/dt$ with $[\text{H}_2\text{O}]^2$, suggesting that both the CI and POZ are trapped by water. Given the complexity of the mechanism, rate constants for the individual steps involved cannot be extracted.

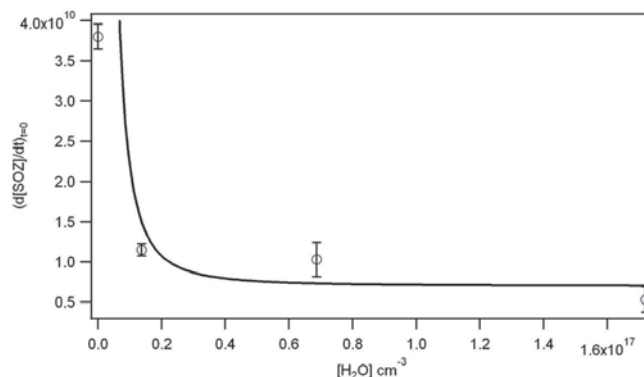


Fig. 13 Initial rate of formation of SOZ as a function of water vapor concentration at an O_3 concentration of 5×10^{13} molecule cm^{-3} . The line is of the form $(d[\text{SOZ}]/dt)_{t=0} = A + B/[\text{H}_2\text{O}]^2$.

Note that while the rate of SOZ formation decreases with water vapor concentration, the overall rate may not. As seen in Fig. 7 in the original paper, while the yield of SOZ decreases with water vapor, those of the carbonyl products increase. The loss of $-\text{CH}_2-$ groups at 2919 and 2850 cm^{-1} is also seen to increase with water vapor concentration, indicating a shift towards more volatile products such as the nine carbon aldehyde and/or acid that are removed in the gas stream.

Conclusions and atmospheric implications

Because unique values of the rate constants (except for k_1) cannot be derived from the data, the individual estimated lifetimes in Table 4 are withdrawn. For a value of $k_1 = (1.8 \pm 0.5) \times 10^{-17}$ cm^3 molecule $^{-1}$ s $^{-1}$ which can be derived from the revised data, the lifetime of OPPC with respect to the ozone reaction at 100 ppb O_3 is 6 hours. From the ratio $k_4/k_2 = (2.2 \pm 0.9) \times 10^{-14}$ cm^3 molecule $^{-1}$, under dry conditions 5% of the POZ will react with O_3 at 100 ppb O_3 , with the rest decomposing to form CI and RCHO. The lifetime will be proportionally shorter, and the fraction of POZ reacting directly with O_3 larger, as the ozone concentration increases. Given the rapid decline in SOZ in the presence of even small amounts of water vapor (revised Fig. 13), trapping of the CI and POZ by water under tropospheric conditions is likely to dominate the chemistry.

Resonance Raman studies of excited state structural displacements of conjugated polymers in donor/acceptor charge transfer complexes

Adam J. Wise and John K. Grey

Phys. Chem. Chem. Phys., 2012, **14**, 11273–11276 (DOI: 10.1039/C2CP41748K). **Amendment published 20th November 2012.**

Eqn (3) in our original paper was missing a factorial sign on the mode order (n_k) in the last term. The correct form should read:

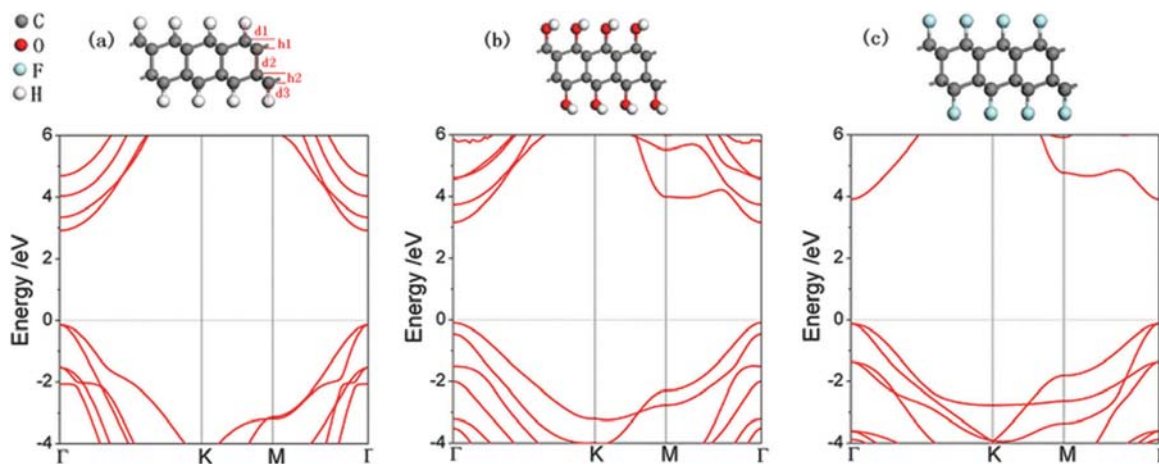
$$\langle \phi_f | \phi(t) \rangle = \prod_k \left\{ \exp \left[-\frac{\Delta_k^2}{2} (1 - \exp(-i\omega_k t)) - \frac{i\omega_k t}{2} \right] \times \dots \right. \\ \left. (1 - \exp(-i\omega_k t))^{n_k} \times \frac{(-1)^{n_k} \Delta_k^{n_k}}{(2^{n_k} n_k!)^{1/2}} \right\}$$

Diamondization of chemically functionalized graphene and graphene–BN bilayers

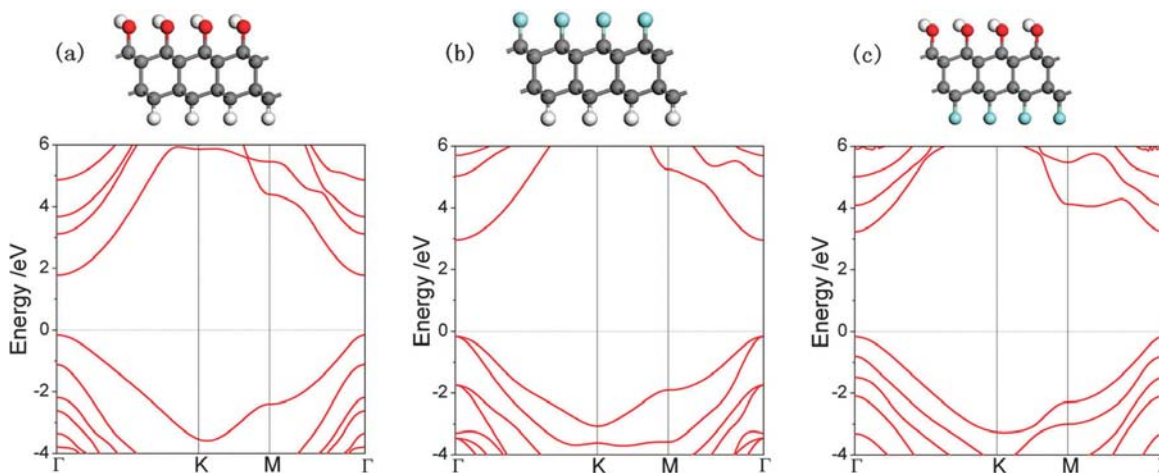
Long Yuan, Zhenyu Li, Jinlong Yang and Jian Guo Hou

Phys. Chem. Chem. Phys., 2012, **14**, 8179–8184 (DOI: 10.1039/C2CP40635G). *Amendment published 18th December 2012.*

On page 8181 the corrected Fig. 3 is given below.



On page 8182 the corrected Fig. 4 is given below.

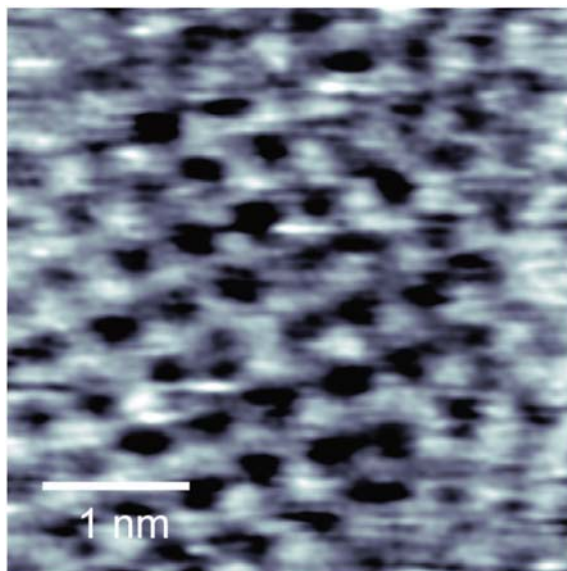


Hydration of hydrophilic thiolate monolayers visualized by atomic force microscopy

Takumi Hiasa, Kenjiro Kimura and Hiroshi Onishi

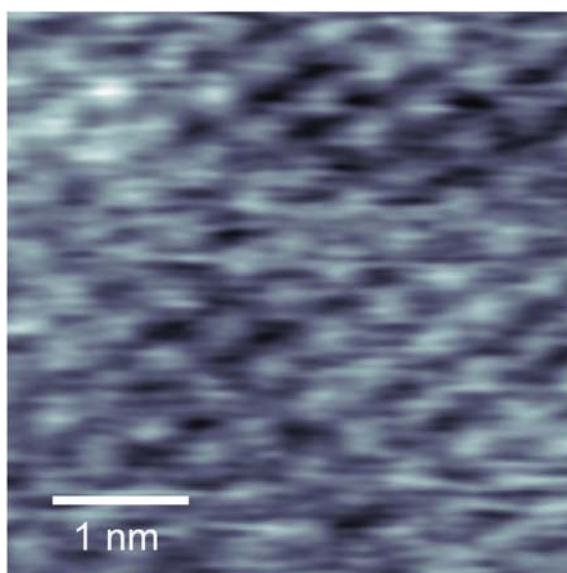
Phys. Chem. Chem. Phys., 2012, **14**, 8419–8424 (DOI: 10.1039/C2CP40252A). *Amendment published 19th December 2012.*

Fig. 5a in our original paper was published as follows:



This image was obtained at 16:35:05 on November 16, 2011. By checking related records in the laboratory notebook, the authors have now confirmed that the image was observed on a COOH-terminated thiolate monolayer which was brought out of n-octanol, dried with N₂ gas, and immersed in a 50 mM KCl aqueous solution. This was not the appropriate conditions for Fig. 5a.

The correct image for Fig. 5a should have been as follows:



This image was obtained at 12:25:43 on November 28, 2011. By checking records in the laboratory notebook, the authors have now confirmed that the image was observed on a COOH-terminated monolayer immersed in a 50 mM aqueous NaCl solution. Fig. 5a should have been prepared with this image. This correction does not affect any discussions and conclusions of the paper.

Mechanism underlying bioinertness of self-assembled monolayers of oligo(ethyleneglycol)-terminated alkanethiols on gold: protein adsorption, platelet adhesion, and surface forces

Tomohiro Hayashi, Yusaku Tanaka, Yuki Koide, Masaru Tanaka and Masahiko Hara

Phys. Chem. Chem. Phys., 2012, **14**, 10196–10206 (DOI: 10.1039/C2CP41236E). *Amendment published 21st December 2012.*

The values of the amounts of adsorbed protein (fibrinogen and BSA) in the above article are incorrect due to the error in converting frequency shift to mass change of the QCM sensor. The corrected values are 3 times higher than the reported values. The corrected figures are presented below. It should be noted that this correction does not affect our conclusion in this paper.

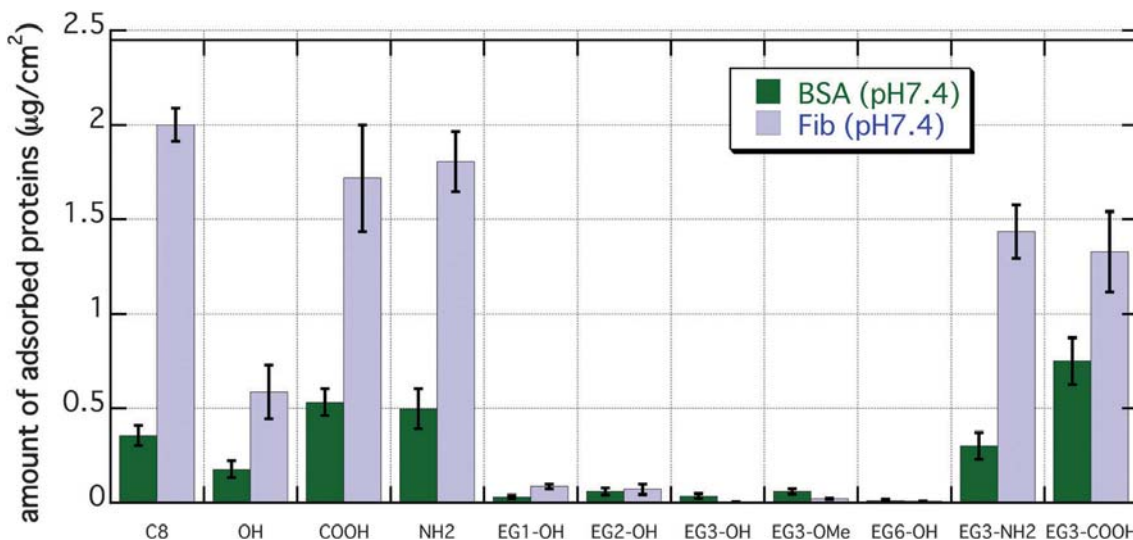


Fig. 2 The amounts of adsorbed BSA and fibrinogen in the PBS solution measured by QCM. The amount is defined in Fig. 1 and its caption. Error bars denote standard deviation ($n = 4$).

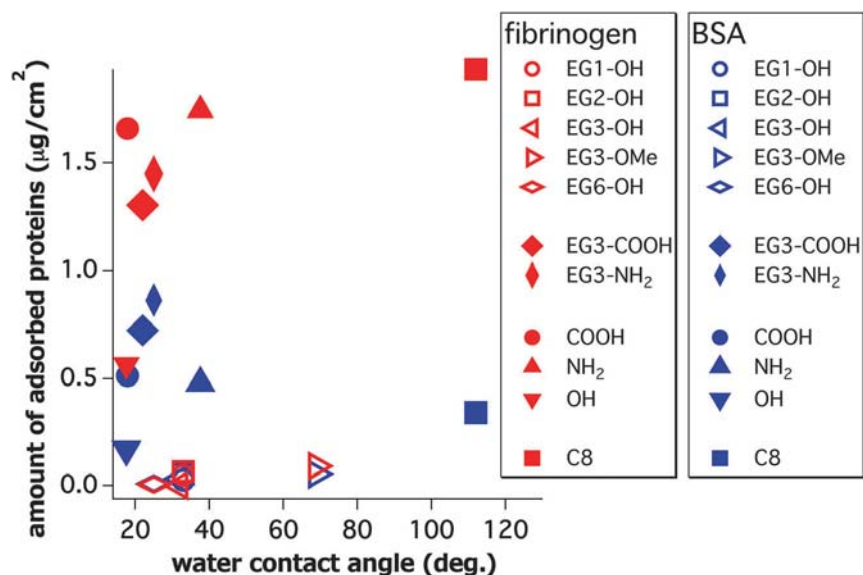


Fig. 3 The amounts of adsorbed BSA and fibrinogen in the PBS solution (same data presented in Fig. 2), plotted as a function of water contact angles.

Negative interference dominates collective transport of kinesin motors in the absence of load

Arthur R. Rogers, Jonathan W. Driver, Pamela E. Constantinou, D. Kenneth Jamison and Michael R. Diehl

Phys. Chem. Chem. Phys., 2009, **11**, 4882–4889 (DOI: 10.1039/B900964G). **Amendment published 1st February 2013.**

The correct referencing of the author shown as D. Kenneth Jamison should be Jamison DK.

A novel method to measure diffusion coefficients in porous metal–organic frameworks

Olexandra Zybaylo, Osama Shekhah, Hui Wang, Maxim Tafipolsky, Rochus Schmid, Diethelm Johannsmann and Christof Wöll

Phys. Chem. Chem. Phys., 2010, **12**, 8093–8098 (DOI: 10.1039/B927601G). **Amendment published 19th February 2013.**

Unfortunately, in our analysis an incorrect expression for a key formula has been used, which affects equations (3) and (4). Equation (3) should read:

$$\frac{M_f(t)}{M_\infty} = 1 - \frac{8}{\pi^2} \sum_{m=0}^{\infty} \frac{1}{(2m+1)^2} \exp\left(-\frac{D(2m+1)^2 \pi^2 t}{4L^2}\right)$$

Equation (4) should read:

$$\frac{M_f(t)}{M_\infty} \approx \frac{2}{\sqrt{\pi}} \sqrt{\frac{Dt}{L^2}}$$

In consequence, the correct value for the diffusivity of pyridine in HKUST-I amounts to $2.4 \times 10^{-18} \text{ m}^2 \text{ s}^{-1}$ and the correct value for the activation energy for diffusion is 0.71 eV (instead of 0.78 eV).

These corrections do not change any of the conclusions presented in the original paper, the agreement with the calculations is actually slightly better than before.

Acknowledgement: We thank Dr Lars Heinke, Karlsruhe, for drawing our attention to this error.

Molecular rearrangement reactions in the gas phase triggered by electron attachment

Benedikt Ómarsson, Elías H. Bjarnason, Sean A. Haughey, Thomas A. Field, Alexander Abramov, Peter Klüpfel, Hannes Jónsson and Oddur Ingólfsson

Phys. Chem. Chem. Phys., 2013, **15**, 4754–4766 (DOI: 10.1039/C3CP44320E). **Amendment published 12th March 2013.**

The incorrect journal title is displayed in Reference 9. The full correct reference is as follows:

B. Ómarsson, E. H. Bjarnason, S. Haughey, T. A. Field and O. Ingólfsson, *Chem. Phys. Lett.*, 2012, **539–540**, 7.

Additionally, the editor list of Reference 19 should read:

Ed. B. J. Berne, G. Ciccotti and D. F. Coker.

Ultrafast photoinduced dynamics of halogenated cyclopentadienes: observation of geminate charge-transfer complexes in solution

T. J. A. Wolf, O. Schalk, R. Radloff, G. Wu, P. Lang, A. Stolow and A.-N. Unterreiner

Phys. Chem. Chem. Phys., 2013, (DOI: 10.1039/C3CP44295K). **Amendment published 2nd April 2013.**

In the last line, fourth column of Table 2 the unit “fs” in brackets suggests that the values given are in units of femtoseconds. This is wrong; the values in this column are all given in units of picoseconds as indicated by the “ps” in the table header. The general statements of the article are not altered in any way by this correction.

Quantum yields for the photolysis of glyoxal below 350 nm and parameterisations for its photolysis rate in the troposphere

Robert J. Salter, Mark A. Blitz, Dwayne E. Heard, Tamás Kovács, Michael J. Pilling, Andrew R. Rickard and Paul W. Seakins

Phys. Chem. Chem. Phys., 2013, 15, 4984–4994 (DOI: 10.1039/c3cp43597k). **Amendment published 1st May 2013.**

Reference 1 refers to the companion paper and has been published. The correct reference is: R. J. Salter, M. A. Blitz, D. E. Heard, M. J. Pilling and P. W. Seakins, *Phys. Chem. Chem. Phys.*, 2013, 15, 6516. This paper can be found here: C3CP43596B.

Pressure and temperature dependent photolysis of glyoxal in the 355–414 nm region: evidence for dissociation from multiple states

Robert J. Salter, Mark A. Blitz, Dwayne E. Heard, Michael J. Pilling and Paul W. Seakins

Phys. Chem. Chem. Phys., 2013, 15, 6516–6526 (DOI: 10.1039/c3cp43596b). **Amendment published 1st May 2013.**

This paper is part 2 of 2 and has a companion paper. This paper is reference 4, R. J. Salter, M. A. Blitz, D. E. Heard, T. Kovács, M. J. Pilling, A. R. Rickard and P. W. Seakins, *Phys. Chem. Chem. Phys.*, 2013, 15, 4984–4994 and can be found here: C3CP43597K.

In situ atomic force microscopy of zeolite: A dissolution

L. Itzel Meza, Michael W. Anderson, Ben Slater and Jonathan R. Agger

Phys. Chem. Chem. Phys., 2008, 10, 5066–5076 (DOI: 10.1039/b804857f). **Amendment published 2nd May 2013.**

Our paper on the dissolution of zeolite A using *in situ* atomic force microscopy *Phys. Chem. Chem. Phys.*, 2008, 10, 5066–5076 was a full length paper expanding a previous communication *Chem. Commun.*, 2007, 2473–2475. Due to an oversight the Chemical Communications paper was not cited in the PCCP publication.

Folding kinetics of a polymer

Štěpán Růžička, David Quigley and Michael P. Allen

Phys. Chem. Chem. Phys., 2012, **14**, 6044–6053 (DOI: 10.1039/C2CP00051B). **Amendment published 9th May 2013.**

In our original article, a convergence problem resulted in an averaging error in computing the entropy from a set of Wang–Landau (WL) Monte-Carlo simulations. Here we report corrected results for the freezing temperature of the homopolymer chain as a function of the range of the non-bonded interaction and find that the previously reported forward-flux sampling and brute-force simulation results are in agreement with the revised Wang–Landau calculations. Please find more details of the amendment by Bart Vorselaars, Štěpán Růžička, David Quigley and Michael P. Allen in this link: <http://www.rsc.org/suppdata/cp/c2/c2cp00051b/paper.pdf>.

Conjugate acene fused buckybowls: evaluating their suitability for p-type, ambipolar and n-type air stable organic semiconductors

Uppula Purushotham and G. Narahari Sastry

Phys. Chem. Chem. Phys., 2013, **15**, 5039–5048 (DOI: 10.1039/c3cp44673e). **Amendment published 22nd May 2013.**

We have found typographical errors in the reorganization energy data in Table 3 and Fig. 5. This effect the text as follows:

Page 5045, right column, line 20: “The λ_+ values clustered in the range of 0.01 eV to 0.03 eV” is incorrect and should read “The λ_+ values clustered in the range of 0.01 eV to 0.25 eV”.

Page 5045, right column, line 22: “The λ_+ values are lower in boron doped compounds by 0.08 eV to 0.18 eV compared to undoped bowls” is incorrect and should read “The λ_+ values are lower in boron doped compounds by 0.02 eV to 0.11 eV compared to undoped bowls”. However, none of the changes mentioned here alter the concepts and interpretations of the results of the manuscript.

The Table 3 and Fig. 5 should read as follows.

Table 3 Energies of HOMO, LUMO orbitals, HOMO-LUMO orbital energy gap (HLG) and reorganization energies λ_+ (hole transport), λ_- (electron transport) of boron and nitrogen doped acene fused buckybowls at B3LYP/6-311+G(d)//B3LYP/6-31G level of theory (all values are given in eV)

Str.	HOMO	LUMO	HLG	λ_+	λ_-
Cor-B1-Cor-B	-5.17	-4.90	-0.27	0.08	0.19
Cor-B2-Cor-B	-5.11	-4.87	-0.24	0.07	0.07
Cor-B3-Cor-B	-5.06	-4.83	-0.23	0.05	0.04
Cor-B4-Cor-B	-4.98	-4.77	-0.21	0.01	0.09
Cor-B5-Cor-B	-4.93	-4.73	-0.20	0.08	0.06
Sum-B1-Cor-B	-4.98	-4.61	-0.37	0.07	0.09
Sum-B1-Sum-B	-4.81	-4.35	-0.46	0.07	0.08
Sum-B2-Cor-B	-4.91	-4.58	-0.32	0.10	0.08
Sum-B2-Sum-B	-4.73	-4.33	-0.40	0.07	0.07
Sum-B3-Cor-B	-4.84	-4.55	-0.29	0.06	0.07
Sum-B3-Sum-B	-4.67	-4.31	-0.36	0.06	0.06
Sum-B4-Cor-B	-4.77	-4.44	-0.33	0.07	0.05
Sum-B4-Sum-B	-4.61	-4.29	-0.32	0.06	0.06
Sum-B5-Cor-B	-4.74	-4.38	-0.36	0.08	0.06
Sum-B5-Sum-B	-4.55	-4.27	-0.28	0.06	0.10
Cor-B1-Cor-N	-3.34	-3.01	-0.34	0.14	0.09
Cor-B2-Cor-N	-3.34	-3.05	-0.29	0.12	0.06
Cor-B3-Cor-N	-3.35	-3.09	-0.26	0.09	0.03
Cor-B4-Cor-N	-3.38	-3.14	-0.24	0.10	0.03
Cor-B5-Cor-N	-3.40	-3.18	-0.22	0.09	0.14
Sum-B1-Cor-N	-3.13	-2.71	-0.42	0.25	0.22
Sum-B1-Sum-N	-2.95	-2.44	-0.52	0.14	0.11
Sum-B2-Cor-N	-3.15	-2.77	-0.38	0.14	0.11
Sum-B2-Sum-N	-3.00	-2.51	-0.49	0.12	0.12
Sum-B3-Cor-N	-3.19	-2.83	-0.36	0.10	0.06
Sum-B3-Sum-N	-3.05	-2.59	-0.46	0.10	0.12
Sum-B4-Cor-N	-3.30	-2.91	-0.39	0.09	0.14
Sum-B4-Sum-N	-3.10	-2.69	-0.41	0.09	0.12
Sum-B5-Cor-N	-3.25	-2.93	-0.31	0.07	0.09
Sum-B5-Sum-N	-3.14	-2.77	-0.37	0.05	0.10

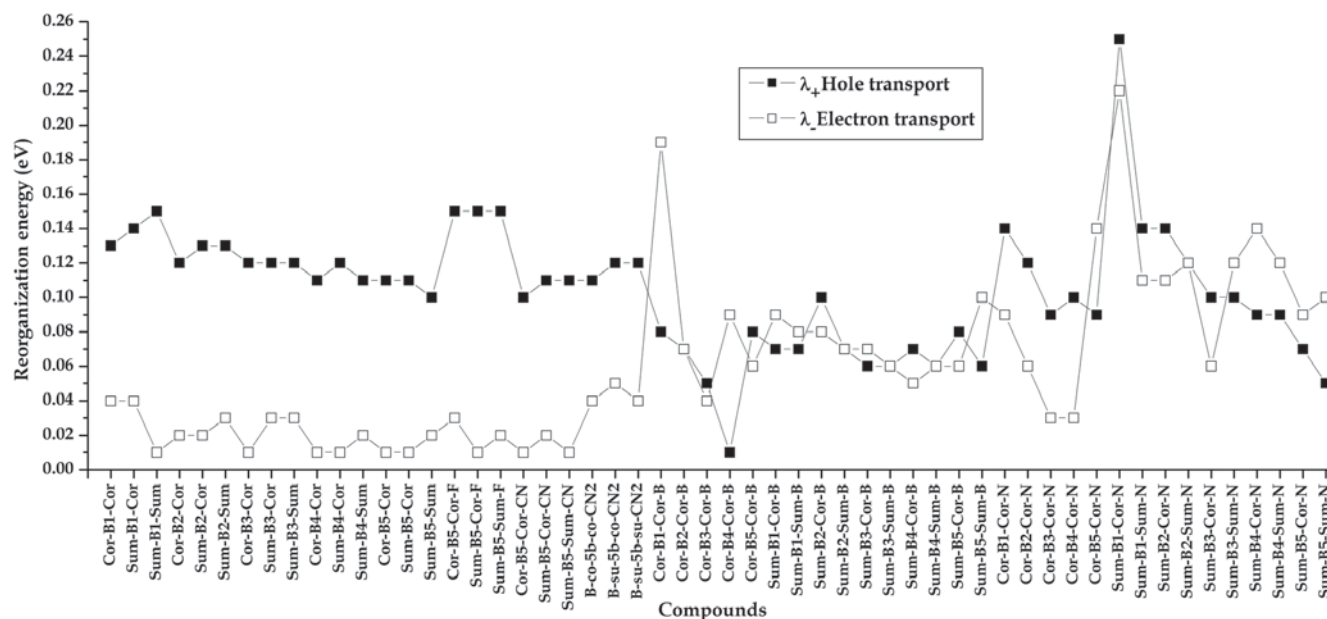


Fig. 5 The reorganization energy, hole transport (λ_+) and electron transport (λ_-) of acene fused buckybowls and their boron and nitrogen doped bowls at B3LYP/6-311+G(d)//B3LYP/6-31G level of theory.

***In situ* XRD studies of nanocrystallization of Fe-based metallic glass: Comparative study by reciprocal and direct space methods**

Jozef Bednarcik, Stefan Michalik, Vladimir Kolesar, Uta Rütt and Hermann Franz

Phys. Chem. Chem. Phys., 2013, **15**, 8470–8479 (DOI: 10.1039/C3CP44445G). **Amendment published 30th May 2013.**

Equation 13 contained an error. The '+' sign is in fact a '-' sign in the numerator. The correct form should read:

$$S(q) = 1 + \frac{I_{eu}^{coh}(q) - \langle f^2 \rangle}{\langle f \rangle^2} \quad (13)$$

Diffusion in binary gas mixtures studied by NMR of hyperpolarized gases and molecular dynamics simulations

R. H. Acosta, L. Agulles-Pedrós, S. Komin, D. Sebastiani, H. W. Spiess and P. Blümler

Phys. Chem. Chem. Phys., 2006, **8**, 4182–4188 (DOI: 10.1039/B609316G). **Amendment published 31st May 2013.**

There is a mistake in equation 1 that could result in a deviation of 13% from the correct value and up to 60% if $\varepsilon = \delta$. The original equation is:

$$b = \gamma^2 G^2 \left(\frac{2}{3} \delta^3 + \frac{\varepsilon^3}{30} - \frac{\delta \varepsilon^2}{6} \right)$$

However, the correct equation 1 should be:

$$b = \gamma^2 G^2 \left(\frac{2}{3} \delta^3 + \frac{\varepsilon^3}{30} - \frac{\delta \varepsilon^2}{6} + \delta^2 \varepsilon \right)$$

The incorrect formula has been replicated and affects the assumptions made in the recent publication R. H. Acosta, P. Blümmler, K. Muennemann and H. W. Speiss, *Progress in Nuclear Magnetic Resonance Spectroscopy*, 2012, **66**, 40–69 (DOI: 10.1016/j.pnmrs.2012.03.003).

Directly measuring reaction kinetics of $\cdot\text{QOOH}$ – a crucial but elusive intermediate in hydrocarbon autoignition

Judit Zádor, Haifeng Huang, Oliver Welz, Johan Zetterberg, David L. Osborn and Craig A. Taatjes

Phys. Chem. Chem. Phys., 2013, **15**, 10753–10760 (DOI: 10.1039/C3CP51185E). **Amendment published 18th June 2013.**

On page 10759, right column, 6th and 7th line it was stated that $k_2, \infty(298 \text{ K}) = (9.0 \pm 3.0) \times 10^{-12} \text{ cm}^3 \text{ molecule}^{-1} \text{ s}^{-1}$. In reality (as it can also be seen in Fig. 4) it is $k_2, \infty(298 \text{ K}) = (9.0 \pm 3.0) \times 10^{-13} \text{ cm}^3 \text{ molecule}^{-1} \text{ s}^{-1}$.

Highly soluble energy relay dyes for dye-sensitized solar cells

George Y. Margulis, Bogyu Lim, Brian E. Hardin, Eva L. Unger, Jun-Ho Yum, Johann M. Feckl, Dina Fattakhova-Rohlfing, Thomas Bein, Michael Grätzel, Alan Sellinger and Michael D. McGehee

Phys. Chem. Chem. Phys., 2013, **15**, 11306–11312 (DOI: 10.1039/C3CP51018B). **Amendment published 5th July 2013.**

In Table 1, the units of V_{OC} should be mV rather than V.

Plasmonic staining of DNA molecules with photo-induced Ag nanoparticles monitored using dark-field microscopy

Yuko S. Yamamoto, Ken Hirano, Tomomi Ishido, Takao Yasui, Norio Murase, Yoshinobu Baba and Tamitake Itoh

Phys. Chem. Chem. Phys., 2013, **15**, 10316–10320 (DOI: 10.1039/C3CP51494C). **Amendment published September 2013.**

In the Experimental section, “ $\text{AgNO}_3 (1 \times 10^{-3} \text{ mM})$ ” on the 18th line of page 2 should be changed to “ $\text{AgNO}_3 (1 \text{ mM})$ ”. The error affects only this part of the paper.

Modulation of the stability of amyloidogenic precursors by anion binding strongly influences the rate of amyloid nucleation

David Ruzafa, Francisco Conejero-Lara and Bertrand Morel

Phys. Chem. Chem. Phys., 2013, **15**, 15508–15517 (DOI: 10.1039/C3CP52313F). *Amendment published 18th October 2013.*

Equation no. 5 (page 15513) in the above article was incorrect. This equation should be as follows:

$$r_0 = \frac{k_F(Bx_N)^2(2 - Bx_N)}{K_A(1 - Bx_N)^2}$$

1H-NMR measurements on proton mobility in nano-crystalline YSZ

Judith Hinterberg, Alina Adams, Bernhard Blümich, Paul Heitjans, Sangtae Kim, Zuhair A. Munir and Manfred Martin

Phys. Chem. Chem. Phys., 2013, (DOI: 10.1039/C3CP53039F). *Amendment published 25th October 2013.*

The authors would like to add an additional acknowledgment as follows:

This work was supported by a collaborative Materials World Network (MWN) grant from the Deutsche Forschungsgemeinschaft (DFG) and the National Science Foundation (NSF).

The Royal Society of Chemistry apologises for these errors and any consequent inconvenience to authors and readers.

Additions and corrections can be viewed online by accessing the original article to which they apply.
