

Center for Advanced Materials

# CAM

# REPORT

Submitted to Molecular Crystals and Liquid Crystals

**Influence of Degree of Polymerization on Phase Separation and Rheology of a Thermotropic Liquid Crystal Polymer**

K.R. Amundson, D.S. Kalika, M.-R. Shen, X.-M. Yu, M.M. Denn, and J.A. Reimer

July 1987

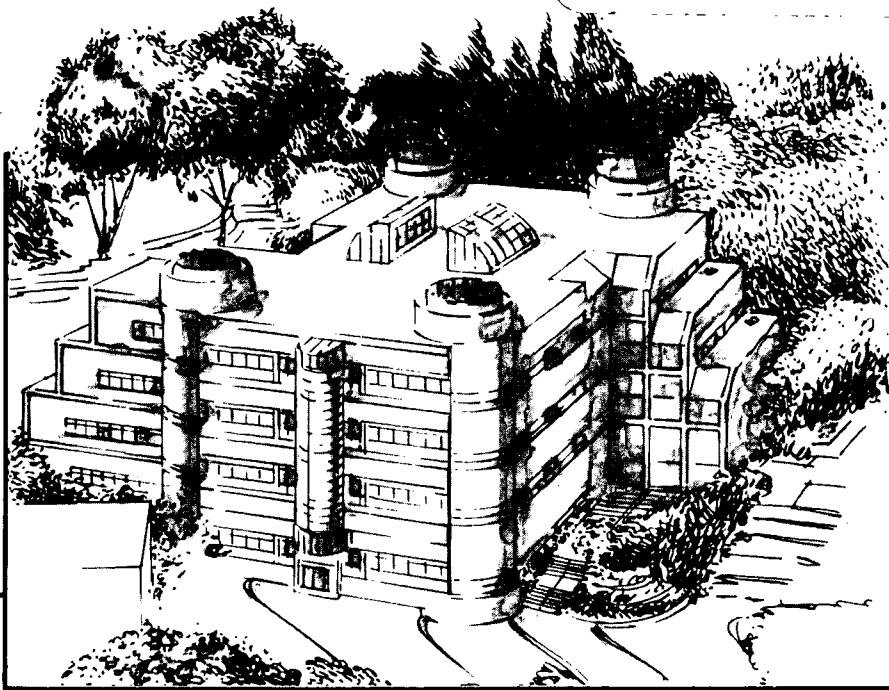
RECEIVED  
LAWRENCE  
BERKELEY LABORATORY

OCT 19 1987

LIBRARY AND  
DOCUMENTS SECTION

## For Reference

Not to be taken from this room



**Materials and Chemical Sciences Division**

**Lawrence Berkeley Laboratory • University of California**

ONE CYCLOTRON ROAD, BERKELEY, CA 94720 • (415) 486-4755

Prepared for the U.S. Department of Energy under Contract DE-AC03-76SF00098

LBL-23827  
c.1

## **DISCLAIMER**

This document was prepared as an account of work sponsored by the United States Government. While this document is believed to contain correct information, neither the United States Government nor any agency thereof, nor the Regents of the University of California, nor any of their employees, makes any warranty, express or implied, or assumes any legal responsibility for the accuracy, completeness, or usefulness of any information, apparatus, product, or process disclosed, or represents that its use would not infringe privately owned rights. Reference herein to any specific commercial product, process, or service by its trade name, trademark, manufacturer, or otherwise, does not necessarily constitute or imply its endorsement, recommendation, or favoring by the United States Government or any agency thereof, or the Regents of the University of California. The views and opinions of authors expressed herein do not necessarily state or reflect those of the United States Government or any agency thereof or the Regents of the University of California.

INFLUENCE OF DEGREE OF POLYMERIZATION ON PHASE  
SEPARATION AND RHEOLOGY OF A THERMOTROPIC LIQUID  
CRYSTAL POLYMER

KARL R. AMUNDSON, DOUGLASS S. KALIKA, MAN-RUO SHEN,  
XIAO-MING YU, MORTON M. DENN, and JEFFREY A. REIMER

Center for Advanced Materials, Lawrence Berkeley  
Laboratory, and Department of Chemical Engineering,  
University of California, Berkeley, California 94720  
U.S.A

ABSTRACT Dynamic rheometry and both high-temperature and solution NMR spectroscopy measurements were carried out on a series of thermotropic liquid crystal copolymers of 60 mole percent hydroxybenzoic acid and 40 mole percent poly (ethylene terephthalate); the members of the series differed only in the extent of solid phase polymerization. The viscosity decreased with increasing degree of polymerization, reaching a broad minimum after about five hours of solid phase polymerization. Solution NMR spectra showed only a slight deviation from chain randomness. Melt NMR spectra at 265°C showed that all members of the series contained both isotropic and nematic phases, with the relative amounts depending on polymerization time.

INTRODUCTION

The development of macroscopic structure during the processing of thermotropic liquid crystal polymers is influenced by the unique microstructure of these materials. Some liquid crystal polymers exhibit time scales for orientation development and relaxation that are very long relative to the times associated with the response of flexible-chain polymers; Moore and Denn<sup>1</sup>, for example, observed characteristic times for the relaxation of orientation on the order of ten minutes following alignment in a magnetic field. The orientation development

K. R. AMUNDSON, D. S. KALIKA, ET AL.

and its manifestation through the macroscopic rheology is undoubtedly associated in part with the "inter-domain" material; that is, the material in regions below the level of resolution of optical microscopy where the transition between the optically-observable regions of high local orientation occurs.

It is believed that some thermotropic liquid crystalline polymers are biphasic in the melt, and that one of these phases is isotropic. Nicely and coworkers<sup>2</sup> have recently reported on the use of proton wide-line NMR spectra to show evidence of the existence of an isotropic phase in copolymers of hydroxybenzoic acid and poly(ethylene terephthalate). We summarize here rheological and NMR experiments on a series of 60 mole percent HBA/40 mole percent PET copolymers having a range of molecular weights. The viscosity of this series shows an initial decrease as a function of molecular weight, suggesting the presence of two phases and a changing nematic fraction with increasing molecular weight. This observation is confirmed using proton wide-line NMR on the melt. Proton NMR spectra of solutions show little change in the chain sequencing with molecular weight.

#### PREPARATION

A copolymer of 60 mole percent hydroxybenzoic acid and 40 mole percent poly(ethylene terephthalate) was synthesized by the reaction of poly(ethylene terephthalate) and p-acetoxybenzoic acid at 275°C according to the preparation method described by Jackson and Kuhfuss.<sup>3</sup> Samples were then polymerized in the solid phase at 210°C under reduced pressure (less than 1 mm Hg) for times ranging up to 36 hours. The synthesis was repeated three times; the observations reported here are for the third batch synthesized, in which solid phase polymerization was carried out for up to 12 hours.

The inherent viscosity was measured in a solution of 60 volume percent phenol/40 volume percent tetrachloroethane at concentrations of approximately 0.60 gm/100 ml. The inherent viscosity is shown in Figure 1 as a function of time of solid phase polymerization. The inherent viscosity of the earlier batches was uniformly lower, but showed the same increase with time of solid phase polymerization.

## PHASE SEPARATION AND RHEOLOGY

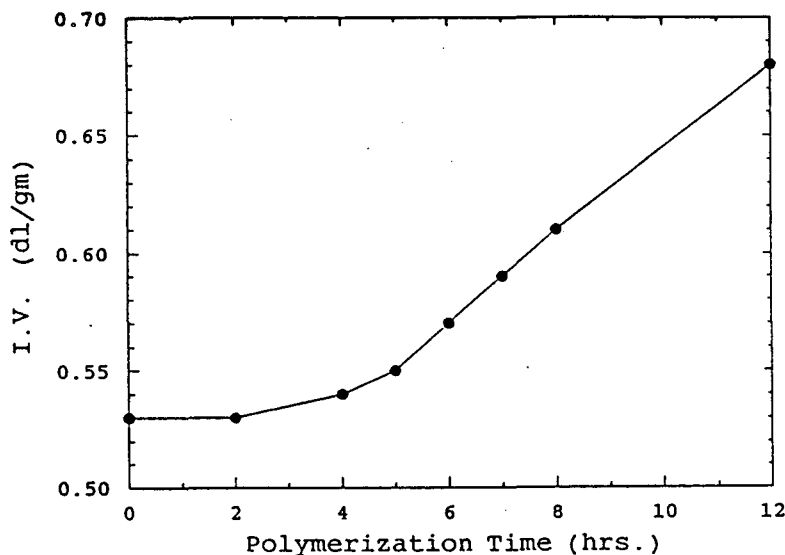


FIGURE 1. Inherent viscosity as a function of solid phase polymerization time.

RHEOLOGY

The rheological properties of the polymers were measured at 265°C using a Rheometrics Mechanical Spectrometer model 705M in the parallel plate arrangement at 5 percent strain. All samples were dried for 21 hours under vacuum at 120°C, and then loaded into the rheometer at the temperature of the experiment under a dry nitrogen atmosphere. Each sample was allowed 20 minutes to come to thermal equilibrium, during which time the gap was set at 1 mm. Frequency sweeps were carried out over the range 0.1 rad/sec to 100 rad/sec, and the complex viscosity ( $\eta^*$ ) and dynamic moduli ( $G'$ ,  $G''$ ) were measured.

Full frequency sweeps are shown in Figure 2 for the lowest and highest molecular weight polymer samples of the third batch. Data at selected frequencies are plotted as a function of solid phase polymerization time in Figure 3. The open symbols are data from replicate runs and represent the range of samples taken from the same batch, while the lines connect the average values; there is substantial variation within the batch, but the variation is less than the change with increasing inherent viscosity. The greatest change occurs over the first few hours of solid phase polymerization, during which time there is little

K. R. AMUNDSON, D. S. KALIKA, ET AL.

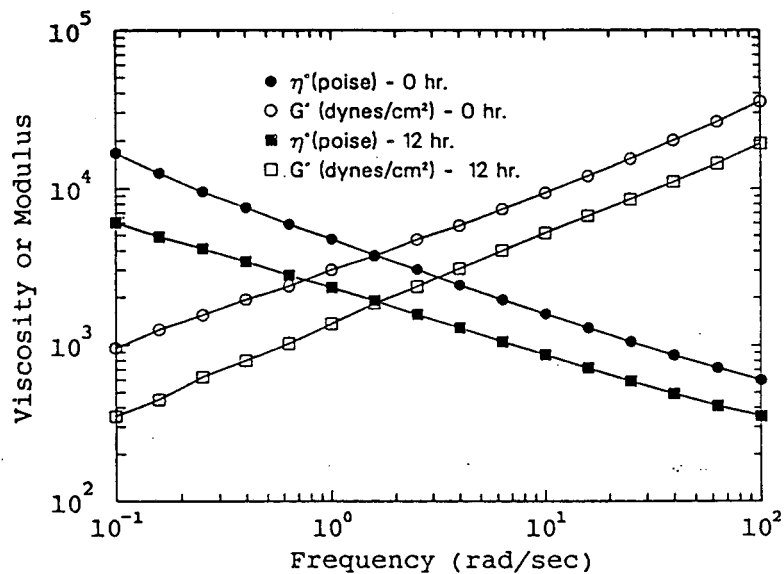


FIGURE 2. Complex viscosity ( $\eta^*$ ) and storage modulus ( $G'$ ) for zero- and twelve-hour solid phase polymerization.

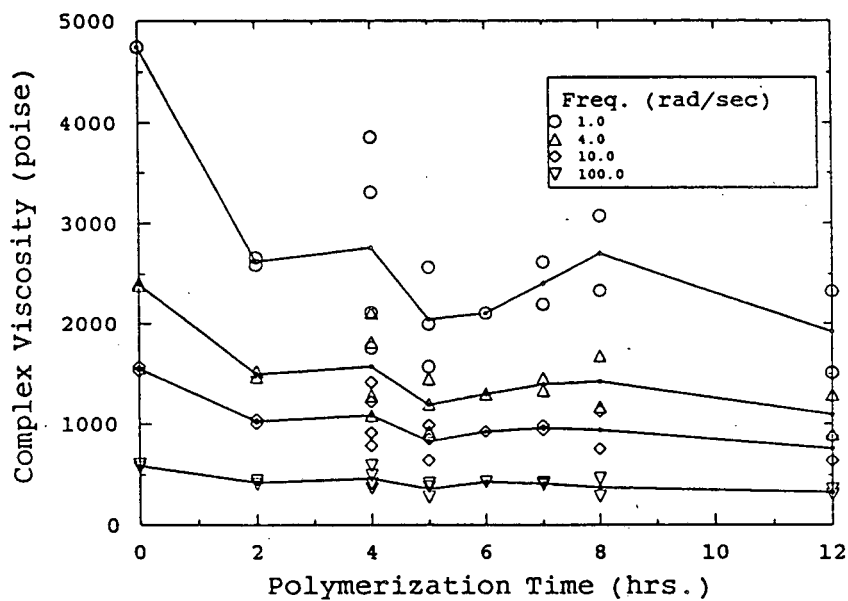


FIGURE 3. Complex viscosity at selected frequencies as a function of time of solid phase polymerization.

## PHASE SEPARATION AND RHEOLOGY

change in inherent viscosity. There is an apparent minimum in the viscosity at all frequencies at about 5 hours of solid phase polymerization. A similar minimum was observed at about the same polymerization time with the first two series that were synthesized, despite the fact that these two series had somewhat lower inherent viscosities and complex viscosities. This result differs from that reported by Fujiwara and coworkers<sup>4</sup> on a series of molecular weights of the same composition.

SOLUTION NMR

Polymer samples from the third batch were dissolved in a 3-to-1 deuterated chloroform/trifluoroacetic acid solution (3 wt%). Proton spectra were obtained on a Bruker AM-500 500 MHz spectrometer. Peak assignments were reported previously by Nicely et al.<sup>2</sup> Peak areas were determined by fitting of the spectra to a theoretical lineshape. Each peak was modeled as a Lorentzian-convoluted Gaussian; the Lorentzian character accounts for the homogeneous linewidth and the Gaussian character accounts for the inhomogeneous (i.e., chemical shift) broadening of each peak. The relevant portion of a typical proton spectrum and an optimal fit is shown in Figure 4.

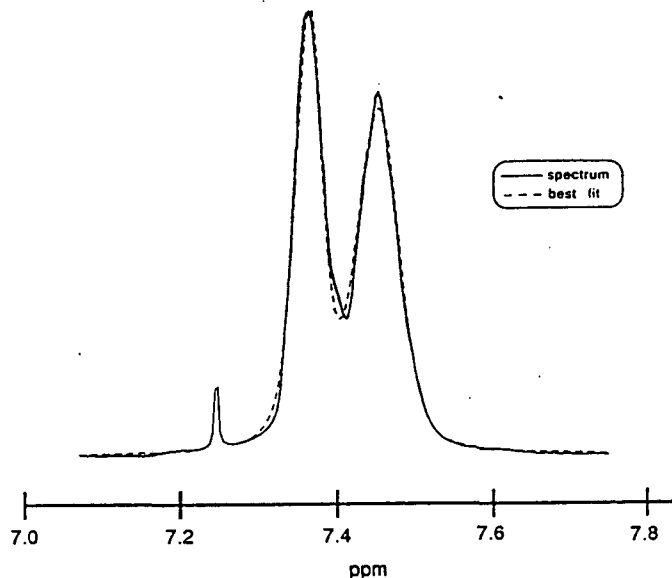


FIGURE 4. Solution proton spectrum and optimal fit for a sample with seven hours of solid phase polymerization.

K. R. AMUNDSON, D. S. KALIKA, ET AL.

The preference factor measured for a sample of 60 percent HBA / 40 percent PET supplied by Eastman was 1.27, which is close to the "best fit" value of 1.3 reported by Nicely and coworkers<sup>2</sup> for a series of Eastman polymers with monomer compositions ranging from 20 to 80 mole percent HBA. The preference factors for various times of solid phase polymerization of the third batch synthesized in this study are shown in Table 1. The proton results were corroborated for one sample using carbon-13 NMR. All values are below 1.3, indicating a small deviation from randomness.

TABLE I Preference factor from solution NMR.

Solid phase polymerization time (hours)	Preference factor
0	1.20
2	1.21
4	1.20
5	1.26
7	1.22
8	1.21
12	1.25

#### HIGH-TEMPERATURE LIQUID NMR

Polymer samples were dried for two days in a vacuum oven at 120°C and then sealed in 5mm NMR tubes. Broadline proton spectra were obtained on a homebuilt spectrometer operating at 180 MHz with a 4.3 Tesla superconducting solenoid. In order to insure thermal equilibrium of the sample, data were recorded at 10 and 18 minutes after reaching 265°C. Spectra were also recorded 25, 45, 65 and 95 minutes after reaching 265°C in earlier experiments so that transient behavior could be observed.

A typical proton spectrum is shown in Figure 5. The center peak is associated with protons in regions of the polymer sample that experience rapid isotropic motion. Two narrow peaks can be resolved here, one associated with aromatic protons and one with aliphatic protons. The broad pairs of peaks to the sides of the center peak are



## PHASE SEPARATION AND RHEOLOGY

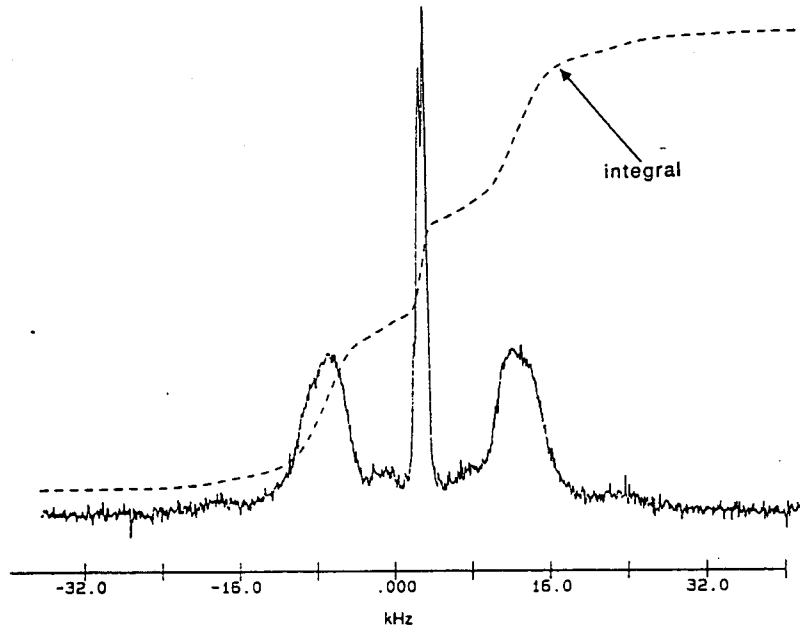


FIGURE 5. Melt proton spectrum at 265°C for a sample with eight hours of solid phase polymerization.

indicative of highly anisotropic motion and are associated with protons in nematic domains. The fraction of isotropic phase was determined by integration of the spectrum. The isotropic fraction for the 60 percent HBA / 40 percent PET sample supplied by Eastman was 0.22; this is smaller than the value of 0.38 for this composition and temperature read from the graph in Nicely et al.<sup>2</sup>

The isotropic fraction is shown as a function of the time of solid phase polymerization in Figure 6. There is a monotonic increase from 0.09 to 0.14, followed by a slight decrease; the maximum appears to correspond to the broad minimum in the viscosity curve, although trends are very sensitive to the large variation in properties within each batch. All samples have a smaller isotropic fraction than the Eastman sample. The second batch, which had a lower inherent viscosity for all times of polymerization, shows a similar trend, with isotropic fractions that are uniformly slightly higher. The trend in the data is difficult to rationalize.

The second moments of the melt spectra, divided by the nematic fraction, are plotted in Fig. 7 as a function of the isotropic fraction. The ordinate is a measure of the

K. R. AMUNDSON, D. S. KALIKA, ET AL.

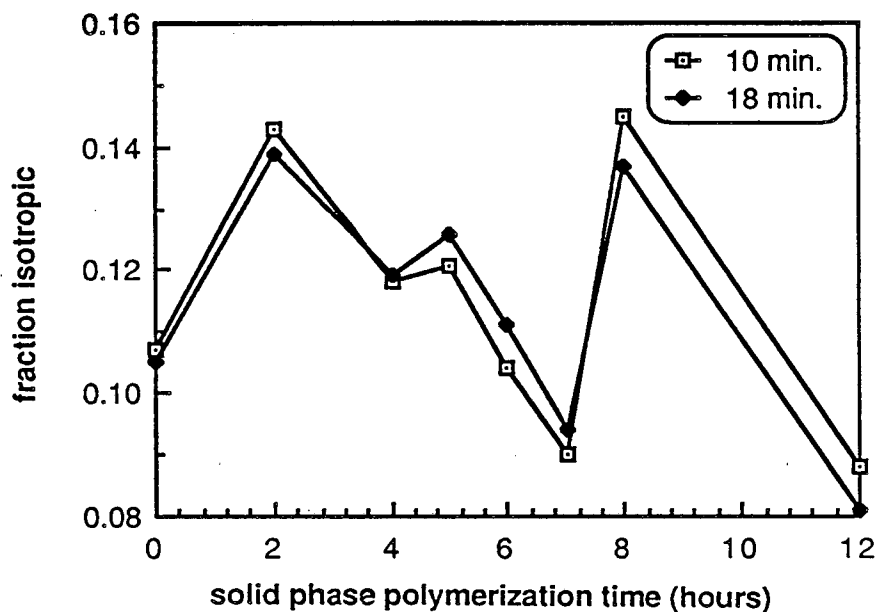


FIGURE 6. Isotropic fraction as a function of time of solid phase polymerization, third batch synthesized.

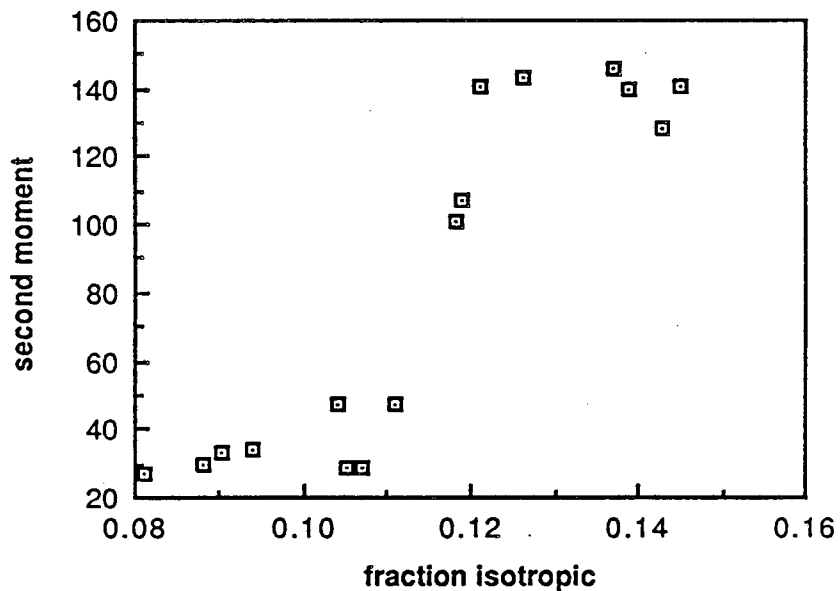


FIGURE 7. Second moments of the melt spectra, divided by nematic fraction, as a function of isotropic fraction.

## PHASE SEPARATION AND RHEOLOGY

average orientation of the nematic phase. The correlation is excellent, and suggests that the nematic orientation should also correlate with the viscosity. There is too much sample-to-sample variation to establish such a relation, however.

### CONCLUSIONS:

The rheological properties of this liquid crystalline polymer, and hence the development of orientation, are sensitive to the presence of an amorphous phase in the predominantly-nematic material. The trend of the data suggests that the shear viscosity is reduced by increasing the isotropic fraction when the material is predominately nematic, and that the average nematic orientation is increased; this would be consistent with a picture of nematic domains and isotropic material "between" these domains, with the highly-oriented nematic regions remaining largely undeformed during oscillatory rheometry. Regardless of the mechanism, however, it is likely that the long time scales of thermotropic liquid crystal polymers will be significantly affected by small variations in the isotropic fraction, and sensitivity to the phase diagram can be expected to have a major influence on processing behavior.

### ACKNOWLEDGMENT

This work was supported by the Director, Office of Energy Research, Office of Basic Energy Sciences, Materials Science Division of the U.S. Department of Energy under contract number DE-AC03-76SF00098, supplemented by gift funds from E. I. DuPont de Nemours and Company, Inc. K. R. Amundson was a Berkeley Fellow, and D. S. Kalika was an N.S.F. Fellow.

### REFERENCES

1. R. C. Moore and M. M. Denn, in The Path to High Modulus Polymers with Stiff and Flexible Chains, edited by A. E. Zachariades and R. S. Porter (Marcel Dekker, New York, in press).
2. V. A. Nicely, J. T. Dougherty and L. P. Renfro, Macromolecules, 20, 573 (1987).

K. R. AMUNDSON, D. S. KALIKA, ET AL.

3. W. J. Jackson and H. F. Kuhfuss, J. Polym. Sci., Polym. Phys. Ed., 14, 2043 (1976).
4. K. Fujiwara, M. Takahashi, and T. Masuda, paper presented at the Third Annual Meeting of the Polymer Processing Society, Stuttgart, W. Germany, April 7 - 10, 1987.

*LAWRENCE BERKELEY LABORATORY  
TECHNICAL INFORMATION DEPARTMENT  
UNIVERSITY OF CALIFORNIA  
BERKELEY, CALIFORNIA 94720*