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Publication Date

1953-02-27

UNIVERSITY OF CALIFORNIA

Radiation Laboratory

Contract No. W-7405-eng-48

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February 27, 1953

Berkeley, California

OBSERVATIONS ON THE RADIATION DECOMPOSITION OF
SOME C¹⁴-LABELED COMPOUNDS ¹

(1) The work described in this paper was sponsored by the U. S. Atomic Energy Commission.

by

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ABSTRACT

Varying amounts of radiation decomposition during storage in the dry form have been found in the cases of C¹⁴-labeled valine, norvaline, norleucine, choline chloride, calcium glycolate and cholesterol. These data indicate that investigators using labeled organic compounds should make frequent checks of the purity of these compounds in order to exclude the possibility of the presence of decomposition products.

The work described in this paper was initiated by a recent observation that considerable radiation decomposition has occurred to a sample of choline-methyl-C¹⁴ chloride which had been synthesized in these laboratories last spring. This observation, together with a similar one made by Dr. Sidney Weinhouse on samples of C¹⁴-labeled calcium glycolate which were synthesized here four years ago, has illustrated the need for information on the rate of

decomposition of C¹⁴-labeled organic compounds. Since this group (Bio-Organic Group of the Radiation Laboratory) has carried out during the past six years the syntheses of several dozen labeled organic compounds, we have available a number of such compounds which could be examined for evidence of radiation decomposition. From these compounds we picked a list of twelve which had been exposed to large amounts of self-radiation since their syntheses and for which there are unequivocal data to show their high degree of purity at the time of synthesis. All of these compounds had been stored in dry, solid form in a dark cabinet at room temperature. The extent of the radiation decomposition was established by making paper chromatograms of each of the compounds and determining, by means of radioautographs of the chromatograms, if other radioactive constituents had appeared during storage. Similar paper chromatograms had been prepared at the times of synthesis and in all cases there had been no observable (or, at most, less than one percent) radioactive impurity. By comparing the old radioautographs with the ones recently prepared, we have measured the percentage of radioactive impurities which have appeared since the times of synthesis. In the case of choline, which had been stored in a sealed tube, we were able to determine the amount of a volatile radioactive product, namely, trimethylamine, which had formed during storage. None of the other compounds were stored in sealed tubes and any volatile decomposition products would not be detected unless they comprised a considerable fraction (at least ten percent) of the radioactivity of the freshly prepared compound. The data which were obtained are summarized in Tables I and II.

Table I
Compounds Showing More Than One Percent Radiation Decomposition

Compound	Date Prepared	Sp. Act. ($\mu\text{c.}/\text{mg.}$)	Reps ^(a) $\times 10^{-6}$	No. of non-volatile radioactive pro- ducts observed on paper chromatograms ^(b)	% Decom- position
Valine-4,4'-C ¹⁴ ₂ .HCl	Nov. 1951	6.0	10	1	1-2
Norvaline- β -C ¹⁴ .HCl	Jan. 1951	17.7	46	6	5
Norleucine-2-C ¹⁴	March 1952	17.4	20	4	2
Choline-methyl-C ¹⁴ ₁ chloride	May 1952	13.0	11	0 ^(c)	63
Calcium glycolate-1-C ¹⁴	Feb. 1949	5.8	30	1	13 ^(d)
Calcium glycolate-2-C ¹⁴	Feb. 1949	4.0	21	1	22 ^(d)
Cholesterol-4-C ¹⁴ (e)					

(a) These values were calculated as follows:

$$\text{Roentgens equivalent physical} = \frac{(k) (N) (\bar{E}) (\text{Reps/e.v.})}{(\text{Rep}) \quad \text{wt. in grams}}$$

where: k = fraction of radiation absorbed
(arbitrarily given the value of 1
in our calculations)

N = total number of events since the time
of synthesis

\bar{E} = average energy of the radiation
(= 50,000 e.v. for C¹⁴)

Reps/e.v. is taken as 1.9×10^{-14}

(b) These observations were made in February 1953.

(c) The main decomposition product of the methyl-labeled choline is the volatile trimethylamine. Further details are recorded in the experimental part of this paper.

Table I (Continued)

- (d) These values were determined (private communication) by Dr. Sidney Weinhouse who found that radioactive formate was present in these percentages. The formic acid was steam distilled from the glycolic acid and its radioactivity measured as CO_2 after oxidation by mercuric ion.
- (e) Considerable radioactive decomposition of cholesterol-4- C^{14} has been observed by Prof. William G. Dauben (private communication). These observations will be described in detail in a separate publication.

Table II

Compounds Showing Less Than One Percent Radiation
Decomposition

Compound	Date Prepared	Sp. Act. ($\mu\text{c.}/\text{mg.}$)	Reps x 10^{-6}
Glycine-2-C ¹⁴ ·HCl	Dec. 1951	4.9	6
Guanine hydrochloride-4-C ¹⁴	March 1950	2.6	9
8-Azaguanine-4-C ¹⁴	March 1950	3.8	13
Adenine sulfate-4,6-C ¹⁴ ₂	March 1950	2.9	10
8-Azaadenine-4,6-C ¹⁴ ₂	March 1950	4.4	15
Stilbamidine diisethionate- amidine-C ¹⁴ ₂	Feb. 1951	13.7	31
Thyroxine- α -C ¹⁴	May 1950	1.0	3
Succinic acid-2-C ¹⁴	Dec. 1949	3.9	14

The data of Tables I and II show that there is a general relationship between the amounts of radiation (Reps) and the likelihood of finding some radiation decomposition; however, the choline and stilbamidine provide two major exceptions. Our data are too scanty to permit predictions as to the susceptibility of various organic groups to radiation damage — this we hope to do later with the accumulation of further information. For the present, however, it should be emphasized that users of C¹⁴-labeled compounds would be well advised to check the purity of their compounds frequently, particularly if the compound is of high specific activity and/or was prepared several years ago. For checking such purity, paper chromatography provides a very convenient tool, especially for finding non-volatile radioactive impurities. However, it may also be useful for volatile impurities if these compounds have either a basic or acidic function and can be kept in a non-volatile form on a paper chromatogram by using, respectively, an acidic or a basic solvent (for example, the chromatography of choline which is reported in this paper). There are several references in the literature^{2,3,4,5} to the use of paper chromatography for determining

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- (2) R. Ostwald, P. T. Adams and B. M. Tolbert, J. Am. Chem. Soc. 74, 2425 (1952).
 - (3) D. M. Hughes, R. Ostwald and B. M. Tolbert, J. Am. Chem. Soc. 74, 2434 (1952).
 - (4) A. A. Benson, J. A. Bassham, M. Calvin, T. C. Goodale, V. A. Haas and W. Stepka, J. Am. Chem. Soc. 72, 1710 (1950).
 - (5) W. Tarpey, H. Hauptmann, B. Tolbert and H. Rapoport, J. Am. Chem. Soc. 72, 5126 (1950).
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the purity of radioactive preparations.

Experimental

The valine-4,4'-C¹⁴₂, norvaline-3-C¹⁴, and norleucine-2-C¹⁴ which were found to show radiation decomposition were all prepared via an acetamidomalonate synthesis using as starting materials, respectively, isopropylmethyl-C¹⁴₂ iodide, propyl-1-C¹⁴ iodide, and butyl-1-C¹⁴ iodide. All of these preparations will be the subject of a forthcoming publication. The final purification of the labeled compounds was accomplished by elution from an ion-exchange column. Paper chromatography of the preparations (phenol-water in one direction, butanol-propionic acid-water in the other) showed that there was less than one percent of radioactive impurities present.

The preparation of the labeled choline has been reported earlier.⁶

(6) F. Christenson and B. M. Tolbert, Univ. of Calif. Radiation Lab. Report No. 962, unclassified, July 15, 1949.

The material described in the present paper was prepared at a later date (May 1952) using the same procedure. The later preparation included, in addition, the same paper chromatographic search for radiochemical impurities as was carried out on all the labeled compounds mentioned in this report. Again, these impurities were less than one percent of the preparation.

Since the radiation decomposition of the choline is so extensive (63%) we made some efforts to determine what the principal decomposition products were. Upon opening a sealed tube of the choline a pronounced odor of acetaldehyde was observed. The presence of acetaldehyde was established by the preparation directly, from the choline sample dissolved in water, of

the 2,4-dinitrophenylhydrazone. This derivative, after recrystallization once from glacial acetic acid and twice from ethanol, was found to have a melting point of 146° (literature value⁷: 147°). It was also found to be

(7) Shriner and Fuson, "Identification of Organic Compounds", 3rd Edition, John Wiley & Sons, Inc., New York, New York, p. 229.

completely non-radioactive.

About two-thirds of the radioactivity of the original choline preparation was volatile — it was lost in the operation of making a plate for counting the choline chloride. That this volatile product is trimethylamine was demonstrated as follows: (1) Alkali was added to a freshly opened tube of the choline and the volatile products were collected in a connected trap which contained a saturated solution of picric acid in ethanol and which was cooled in a liquid nitrogen bath. The trap was then removed, closed off from the atmosphere, and warmed, with shaking, to about 60° . After the trap was cooled back to 0° , crystalline picrate appeared. The solvent was removed by filtration and the precipitate was dried. Its melting point after two recrystallizations from ethanol was 211° (literature value⁸: 216°) and its activity was 1.67 mc./mmole (the activity of the

(8) Beilstein, "Handbuch der Organischen Chemie", Vol. VI, p. 280.

original choline chloride was 1.82 mc./mmole; both values determined by direct plating). (2) When the labeled choline chloride from a freshly opened tube was placed in acid solution and an aliquot portion paper chromatographed (one-dimensionally) using a n-butanol-conc. HCl-water (4:1:1 by vol.) solvent, two radioactive spots were obtained. The faster

moving spot (higher R_F value) contained 63% of the total activity and the slower moving spot (choline) contained 37% of the activity on the paper. The faster moving spot was cut out, eluted, and the material co-chromatographed with labeled trimethylamine hydrochloride. These materials proved to be chromatographically identical (one-dimensional paper chromatogram with the same acidic butanol solvent).

The syntheses of the labeled glycolic acids, and the criteria for their purity at the time of preparation, have been described in an earlier publication³. The same applies to the guanine, 8-azaguanine, adenine, and 8-azaadenine⁹, to the glycine¹⁰, to the succinic acid¹¹ and to the stilbamide¹². The thyroxine synthesis (R. M. Lemmon, unpublished work) was carried

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- (9) Edward L. Bennett, J. Am. Chem. Soc. 74, 2420 (1952).
 - (10) B. M. Tolbert and D. M. Hughes, Univ. of Calif. Radiation Lab. Report No. 705, Unclassified, May 16, 1950.
 - (11) E. C. Jorgensen, J. A. Bassham, M. Calvin and B. M. Tolbert, J. Am. Chem. Soc. 74, 2418 (1952).
 - (12) Edward L. Bennett, Cancer Research, in press.

out using glycine-2-C¹⁴ and following the original procedure of Harington¹³.

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- (13) G. R. Harington and G. Barger, Biochem. J. 21, 169 (1927).
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